Correlation and prediction of vapor-liquid equilibrium in electrolytic solutions

Anil Krishna Rastogi
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CORRELATION AND PREDICTION OF VAPOR-LIQUID EQUILIBRIUM IN ELECTROLYTIC SOLUTIONS

New Jersey Institute of Technology

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CORRELATION AND PREDICTION OF
VAPOR–LIQUID EQUILIBRIUM IN
ELECTROLYTIC SOLUTIONS

by

Anil Krishna Rastogi

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Dissertation submitted to the Faculty of the Graduate School of the New Jersey Institute of Technology in partial fulfillment of the requirements for the degree of Doctor of Engineering Science
1981
Title of Thesis: Correlation and Prediction of Vapor-Liquid Equilibrium in Electrolytic Solutions

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ABSTRACT

Title of Thesis: Correlation and Prediction of Vapor-Liquid Equilibrium in Electrolytic Solutions

Thesis directed by: Dimitrios Tassios, Professor of Chemical Engineering

Two expressions for the excess Gibbs free energy are presented which correlate and predict vapor-liquid equilibrium and the mean molal activity coefficient of an electrolyte in a ternary mixture containing water and either MeOH or EtOH. The proposed equations take into account coulombic forces between ions and the physical interaction forces between ion-solvent and solvent-solvent molecules in a solution.

Model #1 is a combination of an extended Debye and Hückel equation and the local composition of non-random two liquid (NRTL) model. A ternary mixture requires six adjustable binary parameters to predict activity coefficients. These six parameters are evaluated from three separate binary data reductions. Ternary data are predicted with an average error of $|\Delta \gamma|$ less than 0.03 up to $I = 2$. The parameters are considered temperature independent within a 30 to 40°C temperature range for aqueous electrolyte mixtures and within a 15 to 20°C temperature range for nonaqueous electrolyte mixtures. Both isothermal and isobaric ternary experimental data have been tested for ternary correlation. Iso-
thermal ternary data correlation for systems containing water-methanol solvents give results with an average error of $|\Delta Y|$ less than 0.01 up to $I = 6$. Whereas isobaric ternary data are correlated with an average error of $|\Delta Y|$ less than 0.02 up to $I = 4$. It is possible to extend this model to multi-component mixtures although this has not been investigated in this work.

Model #2 is a combination of the Bromley equation, the simplified NRTL equation and an additional ternary salting out expression. The behavior of each electrolyte-solvent binary is described by a one parameter form of the Bromley equation. The temperature dependency of the binary parameters has been established with a two constant equation. Prediction of aqueous electrolyte binary data ($\gamma^+$ and vapor pressure depression) is obtained with an average percent error less than 10.0 at intermediate temperatures. Correlation of ternary VLE and $\gamma^+$ data require four binary parameters and two additional ternary adjustable parameters. This model is limited to binary and ternary data correlation only. The maximum concentration range for the correlation of ternary systems containing water and methanol solvents is about $I = 3$. The correlation of vapor-liquid equilibrium data results in an average error of $|\Delta Y|$ less than 0.012, except for the LiCl-H$_2$O-MeOH system at 60°C where the average error in $|\Delta Y|$ is 0.02.
The thermodynamic study of electrolytic solutions can be categorized into three groups: one containing strong electrolytes in aqueous solvents; the second comprising volatile weak electrolytes in aqueous/nonaqueous solvents; and the third consisting of strong electrolytes in nonaqueous solvents or in mixed solvents.

The first type of system has been of interest in various chemical, metallurgical and geological problems. Debye and Hückel (1923) proposed the classic thermodynamic excess Gibbs free energy expression for strong electrolytes in a single solvent. Guggenheim (1935) extended the range of validity of the Debye-Hückel equation to 0.1 molal solutions. Recently, many workers have proposed semi-empirical correlations for concentrated electrolyte aqueous solutions viz, Bromley et al. (1972, 1973, 1974); Meissner et al. (1972); Pitzer et al. (1973, 1974, 1977, 1979); and Cruz and Renon (1978). Two important and different approaches among the above are by Bromley (1973) and Cruz and Renon. Bromley modified Guggenheim's equation to a one parameter form per binary whereas the Cruz and Renon expression is a combination of the Debye-Hückel equation, a salting out contribution given by the Born model, and the NRTL model.

The second type of systems recently became important due to the necessity of pollution control in the chemical and
petroleum industries. The recovery of weak volatile electrolytes such as ammonia, carbon dioxide, hydrogen sulfide, sulfur dioxide and hydrogen cyanide from effluent streams requires the thermodynamic representation of vapor-liquid equilibrium. The most interesting work reported in this area is by Van Krevelen (1949); Van Krevelen, Hof zer and Hunt ens (1949); Edwards et al. (1975, 1978); Beutier and Renon (1978); Chen et al. (1979); and Mason and Kao (1979).

Theoretical and correlation for the third type of system, electrolytes in nonaqueous solvents and in mixed solvents is sparse. A knowledge of the VLE and electrolyte activities in such solutions could be useful in different chemical and electrochemical applications. The correlation work in the literature for this category are by Rousseau et al. (1972, 1975, 1978); Bakerman and Tassios (1975); Hala (1969); Chen et al. (1979) and Tomasula and Tassios (1980). The most systematic approach is given by Hala and Chen et al. They have considered different ion-ion, ion-solvent and solvent-solvent interactions in an electrolytic solution. Both have used two types of terms for the Gibbs free energy. One of these terms is the Debye-Hückel equation to represent ion-ion interactions. For the other interactions, Hala used the two suffix Margules equation, whereas Chen et al. used an expression based on the two liquid nonrandom theory. Unfortunately, none of the above approaches presents a predictive scheme for a ternary electrolytic solution; also they are complex in nature.
It would be appropriate to categorize the present state of the art for such systems as in a developmental stage. Therefore we have taken an approach to develop thermodynamic analytical expressions to represent VLE of the third type of electrolytic solutions. Our models also combine the two types of terms, which are derived by the modification of Bromley (1973) and Cruz and Renon (1978) binary equations. These are different than the equations of Hala and Chen et al. The proposed models require a minimum amount of information. Model I has the possibility of extension to multicomponent mixtures containing more than two solvents and single or multi electrolytes.
DEDICATION

I dedicate this thesis to my parents whose encouragement and love made it possible.
ACKNOWLEDGEMENT

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INTRODUCTION

Fundamental knowledge of the VLE behavior of electrolytic solutions and electrolytes in mixed solvents is limited at present. The characterization of equilibrium properties of such systems has become important in the process design and process simulation of different processes.

In this study, two semi-empirical models have been developed based on molecular and ionic interactions in the solutions. Both the models combine modified forms of the NRTL equation and some form of the extended Debye-Hückel equation for physical and coulombic forces in a solution, respectively.

In Chapter 1 thermodynamic relationships for the VLE of electrolytic solutions are presented. In Chapters 2 and 3 the two models are developed. Also, their performance in correlating and predicting binary and ternary data is investigated. In Chapter 4 a comparative study of the two models is discussed. The detailed development of the two models is given in Appendices A, B and C.
CHAPTER 1

DEVELOPMENT OF FUNDAMENTAL RELATIONSHIPS FOR VAPOR-LIQUID EQUILIBRIUM IN ELECTROLYTIC SOLUTIONS

ABSTRACT

In vapor liquid equilibrium calculations it might be necessary to find $y-T$ data from known $x-P$ data or to find $y-P$ data from given $x-T$ data or it may be required to interpolate or extrapolate the limited $x-y-P-T$ data. Secondly, in chemical processes and electrochemical energy conversion, a knowledge of the activities of the solvents and the electrolyte may be useful to characterize the solution behavior. Therefore, in this chapter, some important thermodynamic relationships have been developed which are applied to a binary or a ternary electrolytic solution.
1.1 Criteria of Equilibria

Consider an electrolytic solution at equilibrium at a temperature 'T' and pressure 'P' as shown in (figure 1.1). In the liquid phase, the electrolyte will be in ionic equilibrium due to the dissociation of the electrolyte into ions. In a concentrated electrolytic solution, one expects the presence of ion-pairs and ions depending upon the degree of dissociation. However, in this work, the electrolyte is assumed to be completely dissociated into ions for the concentration range and solvents under consideration. Therefore the liquid-phase, specifically, will consist of solvent molecules and ions. The vapor phase will consist of solvent molecules and the electrolyte in molecular form (if the electrolyte is volatile), because, for the temperature range considered in this work, dissociation of the electrolyte in the vapor phase is negligible.

When two phases are in equilibrium, the chemical potential for component i in the two phases will be the same

\[ \mu_i^L = \mu_i^V \]  

(1-1)

The chemical potential in a single phase can be related to the fugacity of component i in a mixture by--

\[ d\bar{G}_i = d\mu_i = RT \ln f_i \]  

(1-2)

A combination of equations (1-1) and (1-2), results in--
Figure 1.1 Representation of the Vapor-Liquid Equilibrium in Electrolytic Solutions
The fugacity of a component can be further expressed in terms of measurable quantities, viz., \( X, Y, P \) and \( T \). A ternary mixture, in this study, is defined as a mixture of an electrolyte and two solvents and the ternary mixture, is considered a combination of three binary mixtures

1. Binary 1-2: electrolyte (1) - solvent (2)
2. Binary 1-3: electrolyte (1) - solvent (3)
3. Binary 2-3: solvent (2) - solvent (3)

An expression of the fugacity in terms of \( X-Y-P-T \) data, for an individual binary mixture and a ternary mixture, is considered in the following sections.
1.2 Mixture of Solvent (2) - Solvent (3) (Binary 2-3)

The liquid-phase fugacity for solvent 2 or 3 is given by

\[ f_L = x_i y_i (P.E.)_i \phi_i^O P_i^O \]  \hspace{1cm} (1-4)

where,

\[ (P.E.)_i = \exp \left[ \int_0^{\frac{P_i^O}{RT}} \frac{v_i^L}{RT} dP \right] \]  \hspace{1cm} (1-5)

\[ \phi_i^O = \exp \left[ - \frac{B_{ii} P_i^O}{RT} \right] \]  \hspace{1cm} (1-6)

The vapor phase fugacity is given by

\[ f_V = y_i P \phi_i^V \]  \hspace{1cm} (1-7)

where,

\[ \phi_i^V = \exp \left[ \frac{P}{RT} \left( B_{ii} + y_j^2 \delta_{ij} \right) \right] \]  \hspace{1cm} (1-8)

Combining equations (1-4) to (1-8) gives

\[ X_i y_i P_i^O = y_i P_i \]  \hspace{1cm} (1-9)

and

\[ F_i = \frac{\phi_i^V}{\phi_i^O (P.E.)_i} \]  \hspace{1cm} (1-10)

At low pressures, and \( T_{Ri} << 1.0 \), \( F_i \approx 1.0 \) (Appendix I)

This simplifies equation (1-9) to

\[ X_i y_i P_i^O = y_i P \]  \hspace{1cm} (1-11)

The total pressure is calculated by

\[ P = X_2 y_2 P_2^O + X_3 y_3 P_3^O \]  \hspace{1cm} (1-12)
Equations (1-11) and (1-12) are used to correlate or predict the VLE in a solvent-solvent binary.
1.3 Mixture of an Electrolyte and a Solvent (Binary 1-2 or 1-3)

Case I volatile electrolyte

The liquid-phase fugacity for the electrolyte is given by

$$\hat{f}^L_1 = m_Y h_1$$ \hspace{1cm} (1-13)

where,

$$H_1 = f(T) \text{ at low pressures}$$ \hspace{1cm} (1-14)

And the liquid-phase fugacity for the solvent is given by an equation similar to equation (1-11) in section (1-2)

$$f^L_i = x_i Y_i P_0^i$$ \hspace{1cm} (1-15)

The vapor-phase fugacity for both the electrolyte and the solvent is

$$\hat{f}^V_l = y_l P, \text{ assuming } \phi^V_l \approx 1.0$$ \hspace{1cm} (1-16)

($l = \text{electrolyte or solvent}$)

Combining equations (1-13), (1-14) and (1-16)

$$m_Y h_1 = y_1 P$$ \hspace{1cm} (1-17)

$$x_i Y_i P_0^i = y_1 P$$ \hspace{1cm} (1-18)

Case II non-volatile electrolyte

The vapor phase will have only solvent molecules. Equation (1-18) will be the only equilibrium relationship for the solvent $i$

$$x_i Y_i P_0^i = p$$ \hspace{1cm} (1-19)
Usually, binary electrolytic experimental data are expressed in terms of the osmotic coefficient $\phi$, which is defined as

$$
\phi = - \frac{1000}{\nu_m \nu_w} \ln n_a_i 
$$

(1-20)

for Case II, the activity and the activity coefficient of the solvent are interrelated by

$$
\hat{a}_i = \frac{P_i}{P_0} = \gamma_i X_i 
$$

(1-21)
1.4 Mixture of an Electrolyte (1) - Solvent (2) - Solvent (3) 
(Ternary 1-2-3)

Case I volatile electrolyte
The VLE relationship for the electrolyte and solvents (2) and (3) will be given by equations (1-17) and (1-18) respectively.

Case II non-volatile electrolyte
The vapor-phase will have only solvent (2) and (3) molecules. The VLE relationships will be the same as given in section (1.2), equations (1-11) and (1-12). In sections (1.2) to (1.4), the liquid mole-fraction of any component is defined, based on the complete dissociation of the electrolyte.

\[ X_i = \frac{N_i}{\nu m + N_2 + N_3} \quad (i = 2 \text{ or } 3) \quad (1-22) \]

and

\[ X_i = \frac{m}{\nu m + N_2 + N_3} \quad (1-23) \]

Equations (1-1) to (1-23) developed in sections (1.1) to (1.4) are used for relating X-Y-P-T and mean molal activity coefficient data in a binary or a ternary mixture. However, in this work all the systems used are non-volatile, except for the HCl-H₂O system where pressure is given as partial pressure of water.
1.5 Criteria for the Excess Gibbs Free Energy

In practical applications, where the liquid-phase composition and the temperature of the system are known, it is necessary to calculate the total pressure of the system and the vapor-phase composition or the mean molal activity coefficient of the electrolyte. In order to obtain this information the equations presented in sections (1.1) to (1.4) are applied, depending upon the type of system. For the above problem, the additional information needed are \( P_i^0 \), \( H_i \) (if electrolyte is volatile) and liquid-phase activity coefficients \( (\gamma_i^1 \text{ and } \gamma_i^2) \).

At a given temperature, the pure component vapor pressure can be estimated by equation (1-24) which expresses \( P_i^0 \) as a function of temperature:

\[
P_i^0 = \exp \left( C_1 + \frac{C_2}{C_3 + T} + C_4T + C_5T^2 + C_6 \ln T \right) 760 \quad (1-24)
\]

At low pressures, the Henry's constant for the electrolyte (volatile) is expressed as a quadratic function of temperature:

\[
H_i = a_1 + b_1T + c_1T^2 \quad (1-25)
\]

For the activity coefficients, an expression is required in terms of the known variables i.e. liquid-phase composition and temperature of the system. In the liquid-phase the total excess Gibbs free energy can be developed as a function of known variables considering the different interaction forces in solution. The excess Gibbs free energy is then used to obtain expressions for the activity coefficients, as shown
below—

\[
\frac{G^E}{RT} = f(T, N_1, N_2 \ldots) \quad (1-26)
\]

\[
\ln \gamma_i = \frac{\partial G^E/RT}{\partial N_i} \quad (1-27)
\]

\[
\ln \gamma^*_i = \frac{\partial G^E/RT}{\partial N_i} \quad (1-28)
\]

\[
\ln \gamma_\pm = \ln \gamma^*_\pm - \ln(0.001 \mu_m M_w + 1) \quad (1-29)
\]

where

\[
\gamma_\pm = \text{mean molal activity coefficient}
\]

\[
\gamma^*_\pm = \text{mean molar activity coefficient}
\]

In an electrolytic liquid solution, the total excess Gibbs free energy can be attributed to, mainly two type of molecular interactions. One interaction is due to the long-range electrostatic forces or ion-ion interactions. The other interaction takes into account the physical forces due to interactions of ion-solvent and solvent-solvent molecules. In Chapters 2 and 3, two different models have been considered for the excess Gibbs free energy. These models propose different forms of expressions to represent the non ideal behavior of a solution.

The equations developed in sections (1.1) to (1.5) are used first to fit experimental binary data with the activity coefficient expressions presented in Chapters 2 and 3. These are then extended to predict and/or correlate the ternary vapor-liquid equilibrium and mean molal activity coefficient data.
CHAPTER 2

CORRELATION AND PREDICTION OF VAPOR-LIQUID EQUILIBRIUM AND
THE MEAN MOLAL ACTIVITY COEFFICIENT BY MODEL I
IN ELECTROLYTIC SOLUTIONS

ABSTRACT

Mean activity coefficient data and vapor pressure depression data of aqueous and nonaqueous electrolytic solutions are correlated successfully. The maximum molality applicable to model for aqueous-electrolytic and MeOH-electrolytic solutions can be approximated up to I = 6. The binary parameters are considered temperature independent within a 30 to 40°C temperature range, presetting the ion-solvent nonrandom parameters $\alpha_A^i$ and $\alpha_B^i$ to 0.2 and 0.0 respectively. The prediction of ternary VLE and $\gamma_\pm$ is of acceptable quality with an average error of 0.028 in $\Delta Y$ and a 15% average error in $\gamma_\pm$ up to I = 2. Ternary data correlation of both isothermal and isobaric data are of good quality. In general, the prediction of binary data using the parameters obtained by ternary data regression is possible, with an average percent error in $\Delta P$ and $\gamma_\pm$ of 15%.
Model I: Combination of the Extended Debye-Hückel Equation and the NRTL Equation

2.1 Excess Gibbs Free Energy Function

A complete theoretical account of the thermodynamic properties of electrolyte solutions must deal with both long range interionic and short range interactions between ions and solvent molecules. In a dilute electrolytic solution, the magnitude of the long range electrostatic forces is dominant. Based on this fact, Debye and Huckel developed a limiting law to predict the properties of a dilute solution by considering the electrical potential at a point in the solution in terms of the concentrations and charges of the ions and the properties of the solvent. Gronwall, Lamer and Sandved (1928) modified the Debye-Huckel equation by extending the potential functions with additional higher order terms. However, in a solution of an electrolyte in mixed solvents, the solvent-solvent interactions are of as much importance as those of ion-ion and ion-solvent interactions. Therefore, in Model I, an empirically extended form of the Debye-Huckel equation is combined with the non-random two liquid model (NRTL). The NRTL part of the equations not only accounts for solvent-solvent interactions, but describes unaccounted ion-solvent interactions also.

The equations for the excess Gibbs free energy, $G^E$, are as below
\[
\frac{G^E_{\text{RT}}}{RT}\bigg|_{\text{Total}} = \frac{G^E_{\text{RT}}}{RT}\bigg|_{\text{Ext.D.H.}} + \frac{G^E_{\text{RT}}}{RT}\bigg|_{\text{NRTL}} 
\]

\[
\frac{G^E_{\text{RT}}}{RT}\bigg|_{\text{Extended D.H.}} = 2.303v \left( \frac{M_w N_T}{1000} \right) \frac{m}{t} |z_+ z_-| \left[ -A_y \frac{2}{\rho^3} \right] 
\]

\[
\{(1 + \rho l^{1/2})^2 - 2(1 + \rho l^{1/2}) + \ln(1 + \rho l^{1/2}) + \frac{3}{2}\} 
\]

\[
+ A_y^2 \left( \frac{2(a I - 2)}{3a^2} \right) (1 + a I)^{1/2} + \frac{4}{3} \frac{1}{l^{1/2}} \}
\]

\[
+ N_T \left[ (0.001v M_w + 1) \ln(0.001v M_w + 1) \right. 
\]

\[
\left. - 0.001v M_w \right] 
\]

where,

\[
\rho = 1.0 
\]

\[
a = 1.5/|z_+ z_-| 
\]

\[
A_y = \text{Debye-Hückel constant (see Appendix D)} 
\]

\[
\frac{G^E_{\text{RT}}}{RT}\bigg|_{\text{NRTL}} = \frac{1}{RT} \left\{ N_2 \left[ \frac{N_A Z_{\pm 2} + N_3 Z_{\pm 2}}{N_2 G_{32} + N_3 G_{32} + N_2} \right. 
\]

\[
+ N_3 \left[ \frac{N_A Z_{\pm 3} + N_2 Z_{23}}{N_3 G_{32} + N_2 G_{32} + N_3} \right. 
\]

\[
- N_A \left[ \frac{N_2 Z_{\pm 2}}{N_3 G_{32} + N_2} + \frac{N_3 Z_{\pm 3}}{N_2 G_{23} + N_3} \right. 
\]

\[
+ N_A N_2 N_3 \left[ \frac{Z_{32} G_{\pm 2}}{(N_3 G_{32} + N_2)^2} + \frac{Z_{23} G_{\pm 3}}{(N_2 G_{23} + N_3)^2} \right] \}
\]

where,
\[ Z_{ij} = \Delta g_{ij} G_{ij} \]
\[ G_{ij} = \exp\left[-\alpha_{ij} \frac{\Delta g_{ij}}{RT}\right] \quad (2-4) \]
\[ \alpha_{ij} = \alpha_{ji} \]
\[ \Delta g_{ij} \neq \Delta g_{ji} \]
\[ Z_{\pm i} = Z_{Ai} + \frac{\nu_B}{\nu_A} Z_{Bi} \]
\[ G_{\pm i} = G_{Ai} + \frac{\nu_B}{\nu_A} G_{Bi} \quad (2-5) \]

\[ Z_{Ai} = \Delta g_{Ai} G_{Ai} \quad \text{and} \quad Z_{Bi} = \Delta g_{Bi} G_{Bi} \]
\[ G_{Ai} = \exp\left[-\alpha_{Ai} \frac{\Delta g_{Ai}}{RT}\right] \]
\[ G_{Bi} = \exp\left[-\alpha_{Bi} \frac{\Delta g_{Bi}}{RT}\right] \quad (2-6) \]

A stepwise procedure for the development of these expressions is given in Appendix B, sections B.1 and B.2.
2.2 Procedure: Data Reduction and Prediction

In Model I, each of the three binaries have two adjustable parameters. In order to predict ternary behavior it is necessary to evaluate the binary parameters first. Also, the accuracy of the binary data correlation will justify the extension of Model I to ternary or multicomponent mixtures. A stepwise scheme for the data correlation and prediction is depicted in figure 2.1. A nonlinear subroutine LSQ2 is used in the binary or ternary data regression.

The detailed steps for binary and ternary correlation with different objective functions have been discussed separately, as below

A. Binary 2-3: Mixture of Solvent (2) - Solvent (3)

The activity coefficient expressions for these type of systems are obtained by substituting \( m = 0.0 \) and \( N_A = 0.0 \) in equations (2-2) and (2-3) and differentiating the resulting excess Gibbs free energy function. This results in the original NRTL equation of Renon and Prausnitz (1968)--

\[
\ln Y_i = \frac{x_i^2}{RT} \left[ \frac{Z_{ij} G_{ji}}{(X_j G_{ji} + X_i)^2} + \frac{Z_{ij}}{(X_i G_{ij} + X_j)^2} \right] \tag{2-7}
\]

\( G_{ij} \) and \( Z_{ij} \) are given by equation (2-4).

The experimental \( X-y-P-T \) data for these binaries are correlated for the two temperature independent parameters--\( \Delta g_{23} \) and \( \Delta g_{32} \) by presetting \( \alpha_{23} \) to -1.0 as recommended by Marina.
Desired: Correlation and Prediction of
\( Y_{+} \); P and/or \( Y_{\pm} \)
in a Ternary Electrolytic Solution

Figure 2.1 A Stepwise Scheme for Correlation and Prediction
of the VLE and Mean Molal Activity Coefficients with Model I
and Tassios (1972) or to 0.2, 0.3, 0.47 as recommended by
Renon and Prausnitz (1968). The objective function used in
the regression for this type of binary system is

$$o.f. = \sum_{s=1}^{NP} \sum_{i=2}^{3} \left( \frac{\gamma_{sca} - \gamma_{SE}^i}{\gamma_{SE}^i} \right)^2$$

(2-8)

$NP = \# \text{ of points in a system}$

$\gamma_{SE}^i$ is calculated using experimental X-y-P-T data with
equations (1-11) and (1-12).

B. Binary 1-2 or 1-3: Mixture of an Electrolyte and
a Solvent

The activity coefficient expressions are given by
equations (A-5), (A-10), (A-8), (A-12) and (A-9). Experimental
binary data can be correlated either through regression for
$G_{\pm i}$ and $Z_{\pm i}$ or $\Delta g_{Ai}$ and $\Delta g_{Bi}$ in equations (2-5) and (2-6).
However, the temperature independent form ($\Delta g_{Ai}$ and $\Delta g_{Bi}$)
would require values of $a_{Ai}$ and $a_{Bi}$. These two forms of the
parameters make Model I applicable to both isothermal and
isobaric data. Two objective functions are used to evaluate
the binary parameters.

$$o.f. \#1 = \sum_{s=1}^{NP} \left( \frac{\gamma_{sca} - \gamma_{SE}^i}{\gamma_{SE}^i} \right)^2 + \sum_{s=1}^{NP} \left( \frac{\gamma_{\pm ca} - \gamma_{\pm E}^i}{\gamma_{\pm E}^i} \right)^2$$

(2-9)

$$o.f. \#2 = \sum_{s=1}^{NP} \left( \frac{DP_{ca} - DP_{E}^i}{DP_{E}^i} \right)^2 + \sum_{s=1}^{NP} \left( \frac{\gamma_{\pm ca} - \gamma_{\pm E}^i}{\gamma_{\pm E}^i} \right)^2$$

(2-10)

where,

$$DP = P_{i}^{0} - P$$

(2-11)
\( \gamma_i,_{\text{EXP}} \) and DP are calculated using equations (1-17) to (1-21). If experimental binary data are available only as vapor pressure vs molality, the second term in equations (2-9) and (2-10) is zero. In the case where the data are \( \gamma_\pm \) vs molality only, the first term in equations (2-9) and (2-10) is zero.

C. Prediction of Ternary VLE (\( \gamma_2, \gamma_3, P \) and \( \gamma_\pm \))

The binary parameters obtained by individual binary data correlation with the best objective function [Equation (2-10)] are used to predict the activity coefficients \( \gamma_2, \gamma_3 \) and \( \gamma_\pm \) in a ternary mixture with equations (2-12) to (2-18) and (1-29).

\[
\ln \gamma_+^* = \ln \gamma_{\text{Ext.D.H.}}^* + \ln \gamma_{\text{NRTL}}^* \quad (2-12)
\]

\[
\ln \gamma_\pm^* = 2.303\left[-\frac{I^{1/2}}{1 + \rho I^{1/2}} + \frac{I}{(1 + a I)^2}\right] z_+^* z_-^* + \ln(0.001 \text{vM}_W + 1) \quad (2-13)
\]

\[
\ln \gamma_{\text{NRTL}}^* = \frac{\nu_A}{\nu} \frac{1}{RT} \left[ \frac{X_2\{X_3G_{32}Z_{\pm2} + X_2Z_{\pm2} - X_3Z_{32}G_{\pm2}\}}{(X_AG_{\pm2} + X_3G_{32} + X_2)^2} + \frac{X_3\{X_2G_{23}Z_{\pm3} + X_3Z_{\pm3} - X_2Z_{23}G_{\pm3}\}}{(X_AG_{\pm3} + X_2G_{23} + X_3)^2} - \frac{X_2Z_{\pm2}}{(X_3G_{32} + X_2)} - \frac{X_3Z_{\pm3}}{(X_2G_{23} + X_3)} + \frac{X_2X_3Z_{32}G_{\pm2}}{(X_3G_{32} + X_2)^2} + \frac{X_2X_3Z_{23}G_{\pm3}}{(X_2G_{23} + X_3)^2} \right] \quad (2-14)
\]
\[ \ln y_i(2 \text{ or } 3) = \ln y_i, \text{Ext.D.H.} + \ln y_i, \text{NRTL} \]  

\[ \ln y_i, \text{Ext.D.H.} = 2.303 \left( \frac{\nu_m}{1000} \right) [M_{wi}] \frac{A_y}{3} I^{1/2} \sigma_1(\rho I^{1/2}) \]

\[ - N_T M_w I^{1/2} \sigma_1(\rho I^{1/2}) \frac{\partial A_y}{\partial N_i} + M_{wi} A_y^{2} \]

\[ \psi_1(aI) + N_T M_w A_y \frac{\partial A_y}{\partial N_i} |\psi_1(aI)| Z_+ Z_- \]

\[ + \ln(0.001 \nu m M_w + 1) - 0.001 \nu m M_w \]

where,

\[ \sigma_1(\rho I^{1/2}), \psi_1(aI) \text{ are defined in equations (A-12) and (A-13)} \]

\[ \sigma_1(\rho I^{1/2}), \psi_1(aI) \text{ are given by equations (B-35) and (B-36)} \]

\[ \ln y_i, \text{NRTL} = \frac{1}{RT} \left[ \frac{X_A G_{ii} Z_{ii}}{X_A G_{ii} + X_j G_{ji} + X_i} \right]^2 + \frac{X_A X_j Z_{ij} G_{ij}}{(X_A G_{ij} + X_i G_{ij} + X_j)^2} + \frac{X_A X_j Z_{ij}^2 G_{ij}}{(X_A G_{ij} + X_i G_{ij} + X_j)^2} \]

\[ + \frac{X_A X_j Z_{ij} G_{ij}}{(X_A G_{ij} + X_i G_{ij} + X_j)^2} \]

\[ - 2X_A X_i X_j \left( \frac{Z_{ji} G_{ji}}{(X_j G_{ji} + X_i)^3} + \frac{Z_{ij} G_{ij} G_{ij}}{(X_i G_{ij} + X_j)^3} \right) \]

where,

\[ X_A = \nu_A X_1 \]
\[ i = 2 \text{ and } j = 3 \]

or

\[ i = 3 \text{ and } j = 2 \]

In equations (2-13) and (2-16) the Debye-Hückel constant and the slope of the Debye-Hückel constant are calculated by a procedure given in Appendix D. Once the solvent activity coefficients are estimated, they are then used to predict \( y_i \) and \( p \) with equations (1-11) and (1-12).

D. Correlation of Ternary Data

The activity coefficient equations (2-12) to (2-18) with equations (1-29), (1-11) and (1-12) are used in ternary data correlation. Three objective functions are attempted

\[
\text{o.f.} \#1 = \sum_{s=1}^{NP} \sum_{j=2}^{NP} \frac{\gamma_{jca} - \gamma_{jE}}{\gamma_{jE}}^2 + \sum_{s=1}^{NP} \frac{\gamma_{\pm ca} - \gamma_{\pm E}}{\gamma_{\pm E}}^2
\]

(2-19)

\[
\text{o.f.} \#2 = \sum_{s=1}^{NP} \frac{P_{ca} - P_{E}}{P_{E}}^2 \sum_{s=1}^{NP} \left( Y_3 - Y_3 \right) x 10 \]

\[
+ \sum_{s=1}^{NP} \frac{\gamma_{\pm ca} - \gamma_{\pm E}}{\gamma_{\pm E}}^2
\]

(2-20)

\[
\text{o.f.} \#3 = \sum_{s=1}^{NP} \sum_{j=2}^{NP} \frac{\gamma_{jca} - \gamma_{jE}}{\gamma_{jE}}^2 + \sum_{s=1}^{NP} \left( Y_3 - Y_3 \right) x 10 \]

\[
+ \sum_{s=1}^{NP} \frac{\gamma_{\pm ca} - \gamma_{\pm E}}{\gamma_{\pm E}}^2
\]

(2-21)
In equations (2-20) and (2-21), a weighing factor of ten is used for the deviation in the vapor-phase composition in order to make the magnitude of this term equal to that of the relative percent error in $\Delta P$ and $\gamma_\perp$. In the case where the ternary data are in the form of $m-X-y-P-T$ only, the second term in equation (2-19) is zero and the third term in equations (2-20) and (2-21) is zero. If the data are $m-X$ vs $\gamma_\perp$ only, the first term in equation (2-19), and the first and second terms in equations (2-20) and (2-21) are zero.

The activity coefficient expressions [(2-13) to (2-18)] have six adjustable parameters for a ternary mixture. However, in this work, data are regressed only for four parameters with preset values of $\Delta g_{23}$ and $\Delta g_{32}$ obtained by binary 2-3 data correlation, corresponding to $a_{23} = -1.0$. If the experimental data are isothermal, the parameters evaluated are $G_{+2}'$, $Z_{+2}'$, $G_{+3}$ and $Z_{+3}$. For isobaric data, the temperature independent parameters $\Delta g_A$, $\Delta g_{B2}$, $\Delta g_{A3}$ and $\Delta g_{B3}$ are evaluated.

A stepwise procedure was used for the rapid convergence of the regression program for the four parameters. In the case of isothermal data, first $G_{+2}$ and $Z_{+2}$ were set to the values obtained by binary aqueous electrolyte data correlation and the ternary data were regressed for $G_{+3}$ and $Z_{+3}$. The second time $G_{+3}$ and $Z_{+3}$ were fixed at the regressed values obtained in the first step and $G_{+2}$ and $Z_{+2}$ were evaluated by ternary data correlation. The third time, the values of $G_{+2}$ and $Z_{+3}$ obtained in the second trial were used and the step
one was repeated to obtain new values of $G_{\pm 3}$ and $Z_{\pm 3}$. This procedure is performed for four or five trials. Finally, the four parameters are evaluated together by ternary data reduction using the values of $G_{\pm 2}$, $Z_{\pm 2}$, $G_{\pm 3}$ and $Z_{\pm 3}$ obtained from the last step as starting values in the regression.

E. Binary Data Prediction Using the Parameters Evaluated by Ternary Data Reduction

The binary parameters [Section 2.2 D] obtained in ternary data correlation are used to predict the activity coefficients for binaries 1-2 and 1-3, with equations (A-5), (A-8), (A-10), (A-11) and (A-9). These are then used to calculate vapor pressure depressions.
2.3 Results

A list of binary and ternary systems used in this study is presented in Tables G.1 and G.2.

A. Binary Data Reduction

The solvent-solvent binary data correlation was obtained with two values of $\alpha_{23}$. Both the values result in the same order of $\Delta Y_{\text{AVG}}$ and $\Delta P_{\text{AVG}}$, Table G.3. The quantities $\Delta Y_{\text{AVG}}$ and $\Delta P_{\text{AVG}}$ are defined below

$$\Delta Y_{\text{AVG}} = \frac{\sum_{s=1}^{NP} |\Delta Y|_s}{NP}$$

$$\Delta P_{\text{AVG}} = \frac{\sum_{s=1}^{NP} |\Delta P|_s}{NP}$$

where

$$\Delta Y = Y_{3_{ca}} - Y_{3_{E}}$$

$$\Delta P = P_{ca} - P_{3}$$

The aqueous or nonaqueous electrolytic binary results are presented in Tables G.4 to G.7. The quantities used to define the accuracy of the correlation are as below

$$\text{Avg % error in } \gamma_{\pm} = \frac{\sum_{s=1}^{NP} \left| \frac{\gamma_{\pm_{ca}} - \gamma_{\pm_{E}}}{\gamma_{\pm_{E}}} \right|_s}{NP} \times 100$$

$$\text{Avg % error in } \Delta P = \frac{\sum_{s=1}^{NP} \left| \frac{\Delta P_{ca} - \Delta P_{E}}{\Delta P_{E}} \right|_s}{NP} \times 100$$
The objective function #2, given in equation (2-10) gives a better fit of the data, Tables G.4 and G.5, specifically for the nonaqueous electrolytic binaries. In further studies, only objective function #2 is used in the binary data correlation, except for the HCl-H₂O binary at 25°C. Since HCl is a volatile electrolyte, and the experimental data are in the form of partial pressure of water in the vapor phase, objective function #1 is used.

B. Temperature Dependency of the Binary Parameters

Next, the temperature dependency of the binary parameters was determined. The parameters \( \Delta g_{A_1} \) and \( \Delta g_{B_1} \) (equation 2-6) were evaluated by presetting \( \alpha_{A_1} \) and \( \alpha_{B_1} \). The meaningful values of \( \alpha_{A_1} \) and \( \alpha_{B_1} \) will be those which can interrelate two forms of the parameters: \( G_{\pm_i} \), \( Z_{\pm_i} \) and \( \Delta g_{A_1} \), \( \Delta g_{B_1} \). It was observed that one of the two \( \alpha \)'s (\( \alpha_{A_1} \) or \( \alpha_{B_1} \)) should be set to zero in order to represent one form of the parameters in terms of the other form. In this work, \( \alpha_{B_1} \) is set to zero and the value of \( \alpha_{A_1} \) was selected arbitrarily and is set equal to 0.2. The results of data reduction obtained for aqueous/nonaqueous electrolytic binaries in the temperature independent form are given in Tables G.6 and G.7. The accuracy of binary data correlation with \( \Delta g_{A_1} \) and \( \Delta g_{B_1} \) is of the same order as that obtained with \( G_{\pm_i} \) and \( Z_{\pm_i} \) (Tables G.4 and G.5). Typical results are given for the system NaCl-H₂O at 25°C in Figure 2.2 and the system LiBr-MeOH at 15°C, Figure 2.3. However, to test the validity of 0.2 and 0.0 for \( \alpha_{A_1} \) and \( \alpha_{B_1} \) respectively, different values of \( \alpha_{A_1} \) and \( \alpha_{B_1} \) were also tried, Tables G.8 and G.9. The
Figure 2.2 Comparison of Experimental Mean Molal Activity Coefficients and Vapor Pressure Depressions with those Predicted by Model I for the System NaCl-H_2O at 25°C.  ____ Experimental, Robinson & Stokes (1955); -authored Pair of Roots #1 (α_A2=0.2, α_B2=0.0);  △ Pair of Roots #2 (α_A2=0.2, α_B2=0.0);  □ Predicted Using Parameters of 60°C (see Table 2.1 for the roots).
Figure 2.3 Comparison of Experimental Vapor Pressure Depressions with Predicted and Correlated by Model I for the System LiBr-MeOH at 15°C

- Experimental, Skabichevskii (1969)
- Correlated (\(\alpha_A = 0.2; \alpha_B = 0.0\))
- Predicted Using Parameters of 30°C
<table>
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<tr>
<th>System</th>
<th>T  (°C)</th>
<th>$\Delta g_{A2}$</th>
<th>$\Delta g_{A3}$</th>
<th>% Error in $\gamma_+ \text{ Max}$</th>
<th>% Error in $\gamma_+ \text{ Avg}$</th>
<th>% Error in DP Max</th>
<th>% Error in DP Avg</th>
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<tr>
<td>NaCl-H$_2$O</td>
<td>25</td>
<td>-32.396</td>
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<td>1.4</td>
<td>2.2</td>
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<td>11.6</td>
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results are of poor quality. This again reinforces the use of \( a_{Ai} = 0.2 \) and \( a_{Bi} = 0.0 \).

The parameters obtained with \( a_{Ai}=0.2 \) and \( a_{Bi}=0.0 \) were used to predict binary data from one temperature to another temperature (Table G.10). In general, for an aqueous electrolyte binary, the data are predicted with a 15 percent average error in \( \gamma_+ \) and a seven percent average error in \( DP \) within a 30-40°C temperature range. The prediction of binary data for the system NaCl-H\(_2\)O at 25°C using the parameters obtained by the data correlation of the system NaCl-H\(_2\)O at 60°C is depicted in Figure 2.2. The availability of nonaqueous electrolytic binary data is limited, therefore it is not possible to establish a temperature range for such binary systems. However, for the two systems LiBr-MeOH and LiCl-MeOH (Table G.10), the results show that the data prediction is less reliable than aqueous electrolytic binaries. The typical result for the system LiBr-MeOH at 15°C is shown in Figure 2.3.

Aqueous electrolytic binary data reduction also indicated a multiplicity of roots for binary parameters (Table 2.1). A binary has two pair of roots. It is interesting to note that the pair of roots with positive values of \( \Delta g_{A2} \) should be used to predict data from one temperature to the other temperature, though these sets of parameters are less accurate in correlating binary data than the other pair of roots (Tables 2.1 and 2.2).
<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
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<th>T (°C)</th>
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<th>% Error in DP</th>
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<td>90</td>
<td>60</td>
<td>171.48</td>
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<td>NaCl-H_2O</td>
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<td>4.0</td>
<td>100</td>
<td>60</td>
<td>171.48</td>
<td>21.695</td>
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<td>$Z_{e}$</td>
<td>% Error in $\gamma_{e}$ Max</td>
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<td>8.0</td>
<td>12</td>
<td>60</td>
<td>4.8525</td>
<td>93.557</td>
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</table>
No multiplicity of roots was observed for the nonaqueous electrolyte binaries.

C. Maximum Molality Applicability

Before this model was extended to ternary systems, its maximum molality applicability was investigated. Three typical systems were used for this: CaCl\textsubscript{2}-H\textsubscript{2}O at 25°C, LiCl-H\textsubscript{2}O at 60°C and LiCl-MeOH at 60°C (Table 2.3). The data correlation for binaries CaCl\textsubscript{2}-H\textsubscript{2}O and LiCl-H\textsubscript{2}O indicate that for aqueous electrolytic mixtures the correlation yields a good fit up to \( I = 9 \), whereas for the system LiCl-MeOH, the data are fitted within an average fifteen percent error only up to \( I = 6 \). This molality limit will be different for different electrolytes, solvents and temperatures of the system. Based on the above study in water and MeOH solvents, the maximum concentration range of an electrolyte is considered \( I = 6 \).

D. Ternary Data Prediction

The binary parameters obtained by the individual binary data reduction were used to predict isothermal \( \gamma_+ \), vapor phase compositions and the total pressures in ternary solutions. For the solvent-solvent binary, both sets of parameters were used corresponding to two different values of \( \alpha_{23} \). The best results are tabulated in Tables 2.4 and 2.5.

Out of four ternary systems with \( m \) vs \( \gamma_+ \) data, only the two systems LiCl-H\textsubscript{2}O-MeOH at 25°C and HCl-H\textsubscript{2}O-MeOH at 25°C
<table>
<thead>
<tr>
<th>System</th>
<th># of Points 'm'</th>
<th>Max T (°C)</th>
<th>α₂₃</th>
<th>Δg₂₃</th>
<th>Δg₃₂</th>
<th>G±₂</th>
<th>Z±₂</th>
<th>G±₃</th>
<th>Z±₃</th>
<th>% Error in γ±</th>
<th>Max</th>
<th>Avg</th>
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<td>0.1</td>
<td>-0.9</td>
<td>336.47</td>
<td>0.1061</td>
<td>-33.295</td>
<td>10.9</td>
<td>39.70</td>
<td>50.7</td>
<td>13.5</td>
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<tr>
<td>LiCl-H₂O-MeOH</td>
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<td>1.0</td>
<td>0.1</td>
<td>-0.9</td>
<td>336.47</td>
<td>0.0557</td>
<td>-81.532</td>
<td>6.1745</td>
<td>56.382</td>
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<tr>
<td>HCl-H₂O-MeOH</td>
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<td>2.0</td>
<td>0.3</td>
<td>806.1</td>
<td>-321.4</td>
<td>0.1061</td>
<td>-33.295</td>
<td>10.9</td>
<td>39.70</td>
<td>28.2</td>
<td>4.1</td>
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<td>LiCl-H₂O-MeOH</td>
<td>45</td>
<td>1.0</td>
<td>0.3</td>
<td>806.1</td>
<td>-321.4</td>
<td>0.1061</td>
<td>-33.295</td>
<td>10.9</td>
<td>39.70</td>
<td>44.6</td>
<td>13.6</td>
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<tr>
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<td>0.1</td>
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<td>383.87</td>
<td>0.094</td>
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<td>12.35</td>
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*Results are shown excluding two points because error for this two points is >60%.
<table>
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<th>Points 'm' (°C)</th>
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<th>Δg&lt;sub&gt;A2&lt;/sub&gt; or G&lt;sub&gt;12&lt;/sub&gt;</th>
<th>Δg&lt;sub&gt;B2&lt;/sub&gt; or Z&lt;sub&gt;12&lt;/sub&gt;</th>
<th>Δg&lt;sub&gt;A3&lt;/sub&gt; or G&lt;sub&gt;13&lt;/sub&gt;</th>
<th>Δg&lt;sub&gt;B3&lt;/sub&gt; or Z&lt;sub&gt;13&lt;/sub&gt;</th>
<th>ΔY</th>
<th>ΔP (mmHg)</th>
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<td>383.8</td>
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<td>56.832</td>
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<td>-150.9</td>
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<td>0.0557</td>
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<tr>
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<td>1.9</td>
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<td>0.3</td>
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<tr>
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<td>1.9</td>
<td>25</td>
<td>-1.0</td>
<td>-150.9</td>
<td>336.47</td>
<td>36.949</td>
<td>1.261</td>
<td>6.876</td>
<td>54.598</td>
</tr>
<tr>
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<td>0.3</td>
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<td>91.21</td>
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<td>19.5</td>
<td>170.9</td>
<td>0.075</td>
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</table>

*Binary parameters are used corresponding to the values evaluated at 60°C, Tables 6.6 and 6.7

**Binary parameters are used corresponding to the values evaluated at 25°C, Tables 2.1 and 6.7

Note: 1. a) Parameters Δg<sub>A2</sub>, Δg<sub>B2</sub>, Δg<sub>A3</sub> and Δg<sub>B3</sub> are used for isobaric data with α<sub>AI</sub> = 0.2; α<sub>BI</sub> = 0.0
b) Parameters G<sub>12</sub>, Z<sub>12</sub>, G<sub>13</sub> and Z<sub>13</sub> are used for isothermal data

2. In general, prediction is better with parameters corresponding to α<sub>23</sub> = -1.0, therefore isobaric data are predicted using α<sub>23</sub> = -1.0 only.
Figure 2.4 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System HCl-H₂O-MeOH at 25°C. 

- Experimental, Akerlof (1930); 
- Correlated ($\alpha_{23} = -1.0$); 
- Predicted ($\alpha_{23} = 0.3$) 

Constant m = 0.02

Constant m = 0.05

Constant m = 0.5
and $\text{HCl-H}_2\text{O-MeOH}$ at $25^\circ\text{C}$ gave acceptable results. The average errors for the $\text{HCl-H}_2\text{O-MeOH}$ system at $25^\circ\text{C}$ are 13.5 and 4.1 for two values of $\alpha_{23}$, -1.0 and 0.3, respectively [Figures 2.4, G.4 and G.5]. The maximum error with $\alpha_{23} = 0.3$ is 28%, which is of good quality. For the system $\text{LiCl-H}_2\text{O-MeOH}$ at $25^\circ\text{C}$, again $\alpha_{23} = 0.3$ gives slightly better results with an average percent error of 14 and a maximum percent error a maximum of 45% (Figures G.6 and G.7). However, prediction of $\gamma_+$ for the system $\text{HCl-H}_2\text{O-EtOH}$ at $25^\circ\text{C}$ is possible only at low molalities and water concentration ($\text{HCl free} \geq 95\%$ [Figures G.1 to G.3]. The prediction of $\gamma_+$ data for the system $\text{NaCl-H}_2\text{O-MeOH}$ at $25^\circ\text{C}$ was not of acceptable quality.

Prediction of the vapor-phase composition and the total pressure was attempted for four systems, depending upon the availability of binary data (Table 2.5 and Figures G.1 to G.12). The average error in $\Delta Y$, in general, is about 0.028. Also, it is interesting to note that as the concentration of MeOH or EtOH increases, the prediction improves, Figures 2.4 and 2.5. However, prediction of the VLE for the system $\text{LiCl-H}_2\text{O-MeOH}$ at $60^\circ\text{C}$ is not of presentable quality.

E. Ternary Data Correlation

All four isothermal ternary data for $\gamma_+$ were correlated alone, Table G.11 and Figures G.1 to G.9. In general, the ternary $\gamma_+$ data are correlated with an average percent error less than eight, except for the $\text{NaCl-H}_2\text{O-MeOH}$ system at $25^\circ\text{C}$ where the maximum percent error in $\gamma_+$ is 27.0. Typical results are
Figure 2.5 Comparison of Experimental VLE with that Predicted Using Model I for the System \( \text{LiCl-H}_2\text{O-EtOH} \) at 25°C. 
- \( \bigcirc \) Experimental (LiCl free) Ciparis (1966);
- \( \bigcirc \) Experimental with LiCl, Ciparis (1966);
- \( \triangle \) Predicted with LiCl (\( \alpha_{23} = 0.3 \));
- \( \blacksquare \) Predicted with LiCl (\( \alpha_{23} = -1.0 \))
Figure 2.6 Comparison of Experimental VLE with that Predicted Using Model I for the System LiCl-H₂O-MeOH at 25°C. 
- Experimental (LiCl free), Ciparis(1966); 
- Experimental with LiCl, Ciparis(1966); 
- Predicted (α₂₃=0.3); 
- Predicted (α₂₃=-1.0)
given for the HCl-H₂O-MeOH system at 25°C in Figure 2.4.

Five isothermal and four isobaric ternary VLE data are correlated with three objective functions. Objective function #2, equation 2.20, yields the best results for both ΔY and ΔP (Tables G.13 and G.14 and Figures G.10 to G.19). In general, the average error in ΔY is about 0.015. The correlation of ternary data for the NaBr-H₂O-MeOH system at 25°C and 40°C (Figures G.13 and G.14) is good, even for molalities greater than six, although it was not possible to predict the data for m > 2. For the LiCl-H₂O-MeOH system at 60°C the prediction of ternary data was not possible, yet the data are correlated successfully up to I = 6 with an average error in ΔY of 0.015, Figure G.15. The data for the LiCl-H₂O-EtOH system at 25°C are fitted only up to m = 1.0, Figures G.11 and G.11. The data available at m = 4.0 for this system could not be fitted within acceptable quality, but this is expected since LiCl is incompletely dissociated in EtOH at this molality. The overall ternary data correlation is of good quality. The parameters obtained with three objective functions are listed in Tables G.13 and G.15.

F. Binary Data Prediction

Finally, aqueous/nonaqueous electrolyte binary data are predicted using the parameters obtained by ternary data reduction, Table 2.6. In general, aqueous electrolyte binary data are predicted with an average percent error of 15 in γ⁺
<table>
<thead>
<tr>
<th>System</th>
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<th>T (°C)</th>
<th>$G_{\pm i}$</th>
<th>$Z_{\pm i}$</th>
<th>% Error in $\gamma_\pm$</th>
<th>% Error in DP</th>
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</thead>
<tbody>
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<td>19.677</td>
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<td>5.126</td>
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<tr>
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<td>10</td>
<td>1.0</td>
<td>25</td>
<td>38.981</td>
<td>4.216</td>
<td>29.2</td>
<td>21.9</td>
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<td>HCl-EtOH</td>
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<td>84.647</td>
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<td>25</td>
<td>0.1382</td>
<td>601.16</td>
<td>-</td>
<td>-</td>
</tr>
</tbody>
</table>
Figure 2.7 Prediction of $\gamma_\mp$ and DP Data for the System LiCl-H$_2$O at 25°C Using the Parameters Obtained by Ternary Data Correlation with Model I
and DP. However, the prediction of DP for nonaqueous electrolyte is not possible. Typical results for the LiCl-H₂O system at 25°C are compared with the experimental data in Figure 2.7.
2.4 Discussion

The main objective of this work is to be able to predict or correlate $\gamma_i$ and the salting out or salting in effect in ternary mixtures. The thermodynamic representation of ternary systems would serve as a guideline in the extension of this model to multicomponent mixtures. An additional term was required with the Debye-Hückel term to represent ternary mixture behavior. Therefore it would be important to analyze the contribution of different terms to understand the behavior of ternary mixtures physically.

In a ternary mixture the impact of the NRTL term, though shifted a little, follows the same trend as it does in a solvent-solvent binary, Figure G.20. However, the contribution of the Debye-Hückel term in a ternary mixture decreases as the dielectric constant of the solvent decreases, i.e.

$$\ln\gamma_{D.H.EtOH} < \ln\gamma_{D.H.MeOH} < \ln\gamma_{D.H.H_2O}$$

So, in a ternary mixture of LiCl-H$_2$O-EtOH, the Debye-Hückel term will always result in salting in for EtOH, Figure 2.8. This salting in effect is due to the Debye Hückel term and is not counterbalanced by the NRTL term. In reality, EtOH is salted out, which is contrary to the effect of the Debye-Hückel term. Therefore it was necessary to include a higher order term ($\ln\gamma_{PHY}$) with the Debye-Hückel equation to cancel the salting-in effect, Figure 2.8. The extended term ($\ln\gamma_{PHY}$)
Figure 2.8 Contribution of Different Terms to $\ln \gamma_i$
for the System LiCl-H$_2$O-EtOH at 25°C in Model I
gives an opposite contribution to that of the Debye-Hückel term, i.e.

\[ \ln \gamma_{\text{PHY,EtOH}} > \ln \gamma_{\text{PHY,MeOH}} > \ln \gamma_{\text{PHY,\text{H}_2\text{O}}} \]

Next, it is important to justify the assumption of complete dissociation of the electrolyte for the applicability of this model. In general, dissociation of an electrolyte in a liquid solution depends upon the characteristics of the electrolyte, properties of the solvent and temperature of the system. As the dielectric constant of the solvent decreases, the ionization of electrolyte decreases also. If dissociation data are not available, it would be appropriate to accept Waddington's (1969) approximation as a guideline. According to Waddington, an electrolyte can be considered completely dissociated up to a moderate concentration range in a solvent with dielectric constant \( \geq 30 \). To determine the moderate range, the correlation of three typical binary data have been studied, Table 2.3. As mentioned in section 2.3, the maximum concentration range for water and MeOH solvents is accepted as \( I = 6.0 \). However, this is based on data at 25°C and 60°C, where the dielectric constant of MeOH is closed to 30. But, if the temperature of the system increases, the dielectric constant decreases and the molality range applicability should be expected to be less than \( I = 6.0 \). This is justified by the ternary VLE data correlation for the NaBr-H\textsubscript{2}O-MeOH system at one atm (temperature range 65-100°C) and isothermal data
Figure 2.9 Thermodynamic Consistency Test for the System H₂O-MeOH at 25°C
at 25°C and 40°C. The average errors in ΔY are 0.017 and 0.012 at 25°C and 40°C, respectively where the data are correlated up to I = 7.1 and I = 6.2 for the two temperatures respectively, Table G.12. However, for isobaric data at one atm, the average error goes up to 0.021, although data up to I = 4 only are used.

In general, ternary $\gamma_\pm$ and VLE data prediction, as shown in Tables 2.4 and 2.5, are of acceptable quality. These can be used as a guideline in preliminary design. Prediction of vapor phase composition and total pressure is better with $\alpha_{23} = -1.0$ than $\alpha_{23} = 0.3$. It is interesting to note that prediction of $\gamma$ and $P$ data for two isobaric systems using the temperature independent parameters (LiCl-H$_2$O-MeOH and NaBr-H$_2$O-MeOH at P = 1 atm, Table 2.5) is obtained with an average error in ΔY of 0.023 and 0.027, respectively. VLE data for four systems out of six systems presented in Table 2.5, are predicted with a $\Delta Y_{AVG}$ of less than 0.028. The average error in ΔY for the other two systems LiCl-H$_2$O-MeOH at 25°C and LiCl-H$_2$O-MeOH at 60°C is larger than 0.028. However, it was found that the maximum concentration limits for reasonable prediction is I = 2. The large errors are observed at higher molality which are due to incomplete dissociation of an electrolyte in the solution.

As shown in Tables 2.4 and 2.5, large errors are obtained for some systems, e.g. especially for the prediction of $\gamma_\pm$ for the HCl-H$_2$O-EtOH system at 25°C and the NaCl-H$_2$O-MeOH system at 25°C and $\gamma$ and $P$ data for the systems LiCl-H$_2$O-MeOH at 60°C and LiCl-H$_2$O-MeOH at 25°C. It is important as a first step to
question the accuracy of binary or ternary data, because the performance of the model in the correlation or prediction of the data is equally dependent upon the quality of the data itself. A plot of $\ln \gamma_2/\gamma_3$ vs $X_2$ (Figure 2.9) which enables the thermodynamic consistency of the data for the system H$_2$O-MeOH at 25°C shows a lot of scattering of the experimental data. This is the best set of data out of three sources which are definitely thermodynamically inconsistent. Hence, binary parameters evaluated for this binary will affect the prediction of $\gamma_\pm$ for the systems HCl-H$_2$O-MeOH at 25°C and NaCl-H$_2$O-MeOH at 25°C and $y$ and $P$ prediction for the systems LiCl-H$_2$O-MeOH at 25°C and NaBr-H$_2$O-MeOH at 25°C.

The solvent-solvent binary data can be tested for thermodynamic consistency, but there is no thermodynamic consistency test to check electrolyte-solvent binary or ternary data. In the literature, a great deal of aqueous electrolyte binary data are available and their quality and accuracy can be considered valid. However, the nonaqueous electrolyte binary or ternary data are not frequently available and those found in the literature can not be checked for accuracy. However, a comparison of isothermal and isobaric data prediction with the same binary parameters evaluated at 60°C indicate that isobaric data can be predicted with an average error in $\Delta Y$ of less than 0.028 whereas the error is large for isothermal ternary data prediction. This shows a possible inconsistency in the VLE data of the LiCl-H$_2$O-MeOH system at 60°C, though this can not be verified.
Another point to be examined in ternary prediction is the difference in the available molality range for binary and ternary systems. As in the case of the HCl-H$_2$O-EtOH system at 25°C, ternary data are available up to $m = 2.5$ for $x'_{\text{EtOH}} = 0.5$, whereas the corresponding HCl-EtOH binary data are available only up to 0.1 m. The data for $m > 0.1$ for the HCl-EtOH binary can not be used with this model because of the incomplete dissociation of HCl. The binary parameters obtained in the regression of the HCl-EtOH data up to 0.1 m can not be expected to perform well for higher molalities in a ternary mixture, especially when the concentration of EtOH increases, (Figures G.1 to G.3). This is also observed with the system NaBr-H$_2$O-MeOH at 25°C. The prediction of this ternary is possible only up to $m = 1.9$, because the binary NaBr-MeOH data are available only up to $m = 1.6$, Table 2.5.

The performance of this model in correlating ternary data as tested with four isothermal systems for $\gamma_\pm$ (Table G.11) and nine systems for VLE (5 isothermal, Table G.12 and 4 isobaric, Table G.14) is of good quality. The data correlated for the NaBr-H$_2$O-MeOH system at 25°C are compared with the results of Chen et al. (1979) (Figure G.14). The results obtained by this model are definitely superior to their model.

Finally, the prediction of the binary data with the parameters obtained by ternary data correlation is generally of acceptable quality for aqueous electrolyte binaries, Table 2.6 and Figure 2.7. The average error in DP and $\gamma_\pm$ is about 15%.
But the prediction of DP data for nonaqueous electrolytes is not possible, as shown for the systems LiCl-MeOH at 25°C and NaBr-MeOH at 25°C (Table 2.6).

2.5 Conclusions

The main objective of this work, which was to represent the thermodynamic behavior of strong electrolytic solutions, is achieved. Aqueous/nonaqueous binary data (DP and $\gamma_+$ vs m) are correlated up to $I = 6$ with an average percent error of 7.0. The prediction of $\gamma_+$ for ternary systems up to $I = 2$ is possible with an average percent error of 15.0. The prediction of ternary VLE data as shown with six systems is possible with an average error in $\Delta Y$ of 0.028 up to $I = 2$, except for the LiCl-H$_2$O-MeOH at 60°C, where error in $\Delta Y$ is large even at $I = 2$. The prediction of VLE data above $I = 2$ results in large errors. This sets the limitation of the model for ternary prediction. The correlation of ternary $\gamma_+$ and VLE data for systems containing water and MeOH is of good quality up to $I = 6$. Ternary system containing water and ethanol was correlated only up to $I = 1$. 
CHAPTER 3

CORRELATION OF VAPOR-LIQUID EQUILIBRIUM AND MEAN MOLAL ACTIVITY COEFFICIENTS WITH MODEL II IN ELECTROLYTIC SOLUTIONS

ABSTRACT

Vapor pressure depression data (DP) of 53 aqueous electrolytes at 100°C were correlated with the one parameter ($B_{12}$) Bromley equation. These $B_{12}$ constants can be used to calculate mean molal activity coefficients up to $I = 6$. However, maximum molality applicability for MeOH-electrolyte solutions is only up to $I = 3$. In addition, these $B_{12}$ values, along with those at 25°C reported by Bromley, can lead to reliable estimates of DP and $\gamma^+_\pm$ in the temperature range 25-100°C.

The extended form of the Bromley equation with the additional NRTL equation and the salting-out term give excellent correlation of the isothermal ternary VLE and $\gamma^+_\pm$ data in electrolytic solutions. This model requires two ternary adjustable parameters, therefore prediction of ternary data with binary data only is not possible.
Model II: Combination of the Bromley Equation, the Simplified NRTL Equation and the Salting Out Term

3.1 Gibbs Free Energy Expression

Bromley (1973) proposed a one parameter equation to correlate binary aqueous electrolytic mixtures. The same equation has been applied successfully to correlate nonaqueous binary data. This equation represents long-range electrostatic forces and ion-solvent interactions in a binary mixture. However, in a ternary mixture additional solvent-solvent molecular interactions must be considered. Therefore in Model II the Bromley equation has been extended to ternary mixtures and combined with a simplified form of the NRTL equation and an additional salting out term:

\[
\frac{G^E}{RT} \text{Total} = \frac{G^E}{RT} \text{Bromley} + \frac{G^E}{RT} \text{NRTL-S} + \frac{G^E}{RT} \text{Salting Out} \tag{3-1}
\]

A stepwise procedure to obtain the total Gibbs free energy expression is presented in Appendix-C (Sections C.1, C.2 and C.3). The final forms of the equations are given below

\[
\frac{G^E}{RT} \text{Bromley} = 2.303 \nu \frac{m}{\Gamma} \frac{N_T M_w}{1000} \left[ -A \gamma \frac{2}{\rho} \frac{1}{2} \left( 1 + \rho I^{1/2} \right) ^2 \right.
\]

\[
- 2 \left( 1 + \rho I^{1/2} \right) + \ln \left( 1 + \rho I^{1/2} \right) + \frac{3}{2}
\]

\[
\left. + \frac{0.06 + 0.6 B}{a^2} \left\{ \ln \left( 1 + aI \right) + \frac{1}{(1 + aI)^2} - 1 \right\} \right]
\]
\[ + \frac{B}{2} I^2 + N_T[(0.001 v m M_w + 1) \ln(0.001 v m M_w + 1) - 0.001 v m M_w] \]

where, \( B = B_{12}X_2' + B_{13}X_3' + B_{123} \frac{(X_2'X_3')^{1/2}e^{-\alpha X_3'}}{(1 + n_1^{1/2})^3} \) \( (3-3) \)

\( B_{123} \) is ternary adjustable parameter

\[ \frac{G^E}{RT} \text{NRTL-S} = \frac{N_T}{RT} \left\{ \frac{X_2X_3X_2'X_3'}{(X_AV_A + X_2 + X_3G_{32})} + \frac{X_2X_3X_2'X_3'}{(X_AV_A + X_3 + X_2G_{23})} \right\} + \frac{\nu V_A X_2X_3}{(X_AV_A + X_2 + X_3G_{32})^2} + \frac{\nu V_A X_2X_3}{(X_2G_{23} + X_3)^2} \] \( (3-4) \)

where, \( G_{ij} \) and \( z_{ij} \) are the binary solvent-solvent parameters, equation (2-4).

\[ \frac{G^E}{RT} \text{Salting Out} = \frac{\delta_{123}^{-1/2} \varepsilon^2}{\alpha n_1} \sum_k \frac{\nu_k^2}{b_k} \frac{n_2^2}{2} \delta'(N_2N_3)^{1/2} \] \( (3-5) \)

where,

\[ \delta' = e^{-\frac{\alpha X_3'}{2}(X_2'B_{13} - X_3'B_{12})} \] \( (3-6) \)

\( \alpha = 2.0 \)

\( \delta_{123} \) is a ternary adjustable salting out parameter. A combination of equations (3-2) to (3-5) with equations (1-27) to (1-29) is used to derive activity coefficient expressions for solvents and the electrolyte. The detailed procedure is given in Appendix-C.
Desired: Correlation of

\[ \gamma_1 \]; \( \gamma_1 \) and/or \( \gamma_2 \)

in a Ternary Electrolytic Solution

Figure 3.1 A Stepwise Scheme for Correlation for the VLE and Mean Molal Activity Coefficients with Model II
3.2 Procedure—Data Reduction

Each binary set is correlated with the appropriate equations for the activity coefficients and the VLE relationships developed in Chapter 1. Binary parameters are then used in the ternary expressions and the ternary isothermal data are correlated. A stepwise scheme for the data correlation is shown in Figure 3.1. A detailed description of the data reduction is given in the following sections.

A. Solvent-Solvent Binary

The activity coefficient equations for these types of systems are derived by setting $m = 0$ and $N_A = 0$ in equations (3-2) to (3-5). The expressions are the same as equation (2-7). The binary data correlation is the same as discussed in Section 2.2-A.

B. Electrolyte-Solvent Binary

An aqueous/non-aqueous electrolytic binary is correlated with the one parameter Bromley equation given by equations (C-1), (C-2) and (C-9). A Fibonacci single variable regression program is used to find the best value of the binary adjustable parameter $B_{12}$ or $B_{13}$. Two objective functions, equations (2-9) and (2-10)(Section 2.2-B), with equations (C-1), (C-2), (C-9), (1-19) to (1-21) are used to correlate the binary experimental data.
C. Electrolyte-Solvent-Solvent Ternary

A ternary mixture requires four binary parameters viz $\Delta g_{23}$, $\Delta g_{32}$ (preset $a_{23}$), $B_{12}$ and $B_{13}$ and two additional ternary parameters.

$$\ln \gamma^*_\pm = \ln \gamma^*_{\pm, \text{Bromley}} + \ln \gamma^*_{\pm, \text{NRTL-S}} + \ln \gamma^*_{\pm, \text{Salt Out}} \quad (3-7)$$

where,

$$\ln \gamma^*_{\pm, \text{Bromley}} = 2.303[-A_{\gamma} |Z^+_\pm Z^-_\pm| \frac{T}{T + \rho T_0 \gamma_{\pm} / k}]^{1/2}$$

$$+ \frac{(0.06 + 0.6B) |Z^+_\pm Z^-_\pm|}{(1 + aI)^2} + BI$$

$$+ 2.303 \frac{N_m}{1000} \sum \frac{m}{n} \frac{0.6 |Z^+_\pm Z^-_\pm|}{a^2}$$

$$\{ \ln(1 + aI) + \frac{1}{(1 + aI)} - 1 \} + \frac{I^2}{2} \frac{\partial B}{\partial N_1}$$

$$+ \ln(0.001\text{vMw} + 1) \quad (3-8)$$

where, $\frac{\partial B}{\partial N_1}$ is defined by equation (C-19)

$$\ln \gamma^*_{\pm, \text{NRTL-S}} = \frac{X_2X_3}{RT} \left[ - \frac{Z_{23}}{(X_{A^\gamma_A} + X_3G_{32} + X_2)^2} \right.$$

$$- \frac{Z_{23}}{(X_{A^\gamma_A} + X_2G_{23} + X_3)^2} + \frac{Z_{32}}{(X_3G_{32} + X_2)^2}$$

$$+ \frac{Z_{23}}{(X_2G_{23} + X_3)^2} \} \quad (3-9)$$
\[ \ln \gamma_{\pm, \text{Salt Out}} = \delta 123 \frac{e^2}{kT D} \sum_k \frac{v_k Z_k}{D_k} (N_2 N_3)^{1/2} \delta' \]

\[ \frac{N_1^{1/2}}{e^\alpha N_1^{1/2}} \left[ 1 - \frac{\alpha}{4} N_1^{1/2} \right] \]

(3-10)

For the solvents 2 and 3:

\[ \ln \gamma_i = \ln \gamma_{i, \text{Bromley}} + \ln \gamma_{i, \text{NRTL-S}} + \ln \gamma_{i, \text{Salt Out}} \]

(3-11)

where,

\[ \ln \gamma_{i, \text{Bromley}} = 2.303 \frac{\nu m}{1000} M_{w_i} \left[ A_\gamma \frac{1}{3} \sigma_2 (\rho 1/2) |Z^+_Z^-| \right] \]

\[- (0.06 + 0.6B) \frac{I}{Z} \psi_2 (aI) |Z^+_Z^-| + B \]

\[ + 2.303 \nu m \frac{N T m_w}{1000} - I - I_1/2 \sigma_2 (\rho 1/2) |Z^+_Z^-| \frac{\partial A}{\partial N_i} \]

\[ + 0.6 \frac{I}{2} \psi_2 (aI) |Z^+_Z^-| + I \frac{\partial B}{\partial N_i} \]

\[ + \ln(0.001 \nu m m_w + 1) - 0.001 \nu m m_w \]

(3-12)

\[ \sigma_2 (\rho 1/2), \psi_2 (aI), \sigma_1 (\rho 1/2), \psi_2 (aI) \text{ and } \frac{\partial B}{\partial N_i} \text{ are defined in equations (C-24) to (C-29). B is given by equation (3-3).} \]

\[ \ln \gamma_{i, \text{NRTL-S}} = \frac{1}{R T} \left[ \frac{X_A X_j v^2 Z_{ji} + X_i^2 Z_{ji} G_{ji}}{(X_A v^2_A + X_j G_{ji} + X_i)^2} \right] \]

\[ + \frac{X_A X_j v Z_{ij} + X_i^2 Z_{ij}}{(X_A v^2_A + X_i G_{ij} + X_j)^2} + \frac{X_A X_j v Z_{ji}}{(X_A v^2_A + X_j G_{ji} + X_i)^2} \]

\[ + \frac{Z_{ij}}{(X_i G_{ij} + G_j)^2} - 2X_A X_i X_j v \frac{Z_{ji}}{(X_j G_{ji} + X_i)^3} \]
\[ Z_{ij} G_{ij} \left[ \frac{1}{X_i G_{ij} + X_j} \right] \]  

(3-13)

Where,

\[ i = 2 \text{ and } j = 3 \]

or

\[ i = 3 \text{ and } j = 2 \]

\[
\ln \gamma_{i, \text{Salt Out}} = \frac{\delta_{123}}{e^{0.5 N_1}} - \frac{1}{2} \frac{(N_1)}{N_1^{1/2}} \delta' - \frac{N_2 N_3^{1/2}}{D} \frac{\partial \delta'}{\partial N_1} \]

(3-14)

Where, \( \delta' \) is given by equation (3-6) and \( \frac{\partial \delta'}{\partial N_1} \) is defined in equations (C-36) and (C-39).

Note: For the development of equations (3-7) to (3-14), see Appendix-C (Section C.4).

The four binary parameters are obtained by individual binary data correlation. A ternary mixture is correlated for the two ternary parameters, \( B_{123} \) and \( \delta_{123} \), using the LSQ2 non-linear regression subroutine. Again, as in Model I, three objective function equations, (2-19) to (2-21), have been tried to correlate the experimental data. In all the ternary data reductions, \( \Delta g_{23} \) and \( \Delta g_{32} \) are preset to their respective values obtained by binary data correlation with \( \alpha_{23} = -1.0 \).
Equations (3-7) to (3-14) with equations (1-11), (1-12), (1-22) and (1-24) are used to correlate the ternary VLE and \( \gamma_{\pm} \) data.
<table>
<thead>
<tr>
<th>System</th>
<th>T (°C)</th>
<th>m</th>
<th>Reference</th>
</tr>
</thead>
<tbody>
<tr>
<td>NaCl-H$_2$O</td>
<td>25, 50, 75, 100</td>
<td>0.1 - 6.0</td>
<td>Gibbard et al. (1974)</td>
</tr>
<tr>
<td></td>
<td>25, 60, 70, 80, 90, 100</td>
<td>0.05 - 4.0</td>
<td>Robinson and Stokes (1955)</td>
</tr>
<tr>
<td>KCl-H$_2$O</td>
<td>25, 40, 50, 60, 70, 80</td>
<td>0.1 - 4.0</td>
<td>Snipes et al. (1975)</td>
</tr>
<tr>
<td>KBr-H$_2$O</td>
<td>25, 60, 70, 80, 90, 100</td>
<td>0.1 - 4.0</td>
<td>Robinson and Stokes (1955)</td>
</tr>
<tr>
<td>MgSO$_4$-H$_2$O</td>
<td>25, 40, 50, 60, 70, 80</td>
<td>0.1 - 2.0</td>
<td>Snipes et al. (1975)</td>
</tr>
<tr>
<td>MgCl$_2$-H$_2$O</td>
<td>25, 40, 50, 60, 70, 80</td>
<td>0.1 - 2.0</td>
<td>Snipes et al. (1975)</td>
</tr>
<tr>
<td>Na$_2$SO$_4$-H$_2$O</td>
<td>25, 40, 50, 60, 70, 80</td>
<td>0.1 - 1.6</td>
<td>Snipes et al. (1975)</td>
</tr>
</tbody>
</table>

*In addition to the Weast compilation*
3.3 Results

A list of aqueous electrolyte binary systems used with this model, in addition to the systems presented in Table G.1, are given in Table 3.1. The results of binary and ternary data correlation are discussed below.

A. Aqueous Electrolyte Binary

Maximum Molality Applicability—Bromley recommended the applicability of his equation [Equations (C-1), (C-2) and (C-9)] up to \( I = 6 \) for strong electrolytes in water, i.e. nearly completely ionized. This has been demonstrated by combining equations (C-9), (1-20) and (1-21) in the form

\[
Y = B_{11} X \tag{3-15}
\]

where,

\[
Y = (1 - \phi) - 2.303 A \gamma |z_+z_-| \sigma_2 (\rho I^{1/2}) I^{1/2} + 2.303 [0.06 \frac{\psi_2(aI)}{2}] |z_+z_-| I \tag{3-16}
\]

\[
X = -2.303 [0.6I |z_+z_-| \frac{\psi_2(aI)}{2} + \frac{I}{2}] \tag{3-17}
\]

Figures H.1 and H.2 indicate that reasonably good results are obtained for strong electrolytes up to \( I = 6 \), i.e. \( m = 6 \) for 1-1 electrolytes (Figure H.1) and \( m = 2 \) for 2-1 electrolytes (Figure H.2). On the other hand, very poor results are observed for MgSO_4 (Figure H.3) which is incompletely ionized.
**Results at 100°C**—In the literature a good deal of data at 100°C are available as DP vs m, Weast (1969). Some typical systems are shown in Table H.1. Equation (C-9) with (l-21) was used to calculate the values of $B_{12}$ at 100°C and $\gamma_\pm$ values are obtained using equations (C-1) and (C-2). Values of m up to $I = 6$ were used as above, even though the range of applicability may be somewhat lower here because of the higher temperature. Hence, for 1-1 electrolyte data up to $m = 6$ (7 points) were used; for 1-2 and 2-1 electrolytes, up to $m = 2$ (3 points); for 2-2 electrolytes, up to $m = 1.5$, and since at $m = 1.5$ is not given, data to $m = 2$ ($I = 8$) were used. Finally, for higher electrolytes (3-1, 3-2, etc.) only two points ($m = 0.5$ and $m = 1.0$) could be used.

This was considered too limited a data base and these electrolytes were not included in this study. The obtained values of $B_{12}$, along with those at 25°C from Bromley, are presented in Table H.2.

In the case that data are correlated for the maximum m value ($m_{\text{max}}$) reported by Weast, the resulting error in DP ($DP'_{\text{max}}$) is also included in Table H.2. The larger values of $DP'_{\text{max}}$ as compared to those of $DP_{\text{max}}$ further support Bromley's suggestion that this equation is applicable only up to $I = 6$ for aqueous electrolytic mixtures.

The accuracy of the mean molal activity coefficients, calculated using these $B_{12}$ values with equations (C-1) and
Figure 3.2 Activity and Osmotic Coefficients for the System Water-
Sodium Chloride at 100°C

--- (Gibbard et al., 1974)
--- (Robinson and Stokes, 1955)

From Weast's Data

Calculated DP Values, eqn. (C-9), (1-21) and (1-22)
TABLE 3.2 Effect of the Number of Data Points Used in Evaluating $B_{12}$ on the Accuracy of the Calculated $\gamma_\pm$ Values (T = 100°C)

<table>
<thead>
<tr>
<th>Max Value</th>
<th>Number of Points</th>
<th>$B_{12}$</th>
<th>Error % in $\gamma_\pm$</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td>Max</td>
</tr>
<tr>
<td>NaCl-Water</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>2.0</td>
<td>3</td>
<td>0.0671</td>
<td>10.9</td>
</tr>
<tr>
<td>3.0</td>
<td>4</td>
<td>0.0595</td>
<td>5.3</td>
</tr>
<tr>
<td>4.0</td>
<td>5</td>
<td>0.0582</td>
<td>5.9</td>
</tr>
<tr>
<td>5.0</td>
<td>6</td>
<td>0.0572</td>
<td>6.4</td>
</tr>
<tr>
<td>6.0</td>
<td>7</td>
<td>0.0567</td>
<td>6.7</td>
</tr>
<tr>
<td>KBr-Water</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>2.0</td>
<td>3</td>
<td>0.0533</td>
<td>6.7</td>
</tr>
<tr>
<td>3.0</td>
<td>4</td>
<td>0.0479</td>
<td>4.9</td>
</tr>
<tr>
<td>4.0</td>
<td>5</td>
<td>0.0442</td>
<td>6.6</td>
</tr>
</tbody>
</table>
(C-2), is examined next. Figure 3.2 presents $\gamma_{\pm}$ values for aqueous NaCl calculated with this approach along with the data from Robinson and Stokes and those of Gibbard et al. (1974). The agreement can be considered reasonably good; maximum error is 6.7%; average error is 5.4%. The other system for which $\gamma_{\pm}$ data at 100°C are available is aqueous KBr (Robinson and Stokes, 1955). Since this system is not included in the Weast compilation, the $\phi$ values of Robinson and Stokes at $m = 0.5$, $1$, $2$, $3$ and $4$ were used. The results are presented in Figure H.4 and they are of the same quality as those in the NaCl case; maximum error is 6.6%; average error is 4.3%. Since this is not the case for all electrolytes, especially for electrolytes other than $1^{-1}$, the effect of the number of data points used on the accuracy of the calculated $\gamma_{\pm}$ values is examined in Table 3.2. The results obtained by using only three data points ($m$ up to 2) are comparable to those obtained by using all points available, up to $m = 6$ for NaCl and up to $m = 4$ for KBr.

**Estimation of DP and $\gamma_{\pm}$ Values in the Range 25-100°C**—In the typical case, values of DP and $\gamma_{\pm}$ are needed at temperatures other than 25 or 100°C. Hence, it would be desirable if $B_{12}$ values could be estimated within this temperature range. Bromley recommends two expressions for the temperature dependency of $B_{12}$:

$$B_{12} = B_0 \ln \left( \frac{T - 243}{T} \right) + \frac{B_1}{T} + B_2 + B_3 \ln T \quad (3-18)$$
<table>
<thead>
<tr>
<th>Electrolyte</th>
<th>$B^*$</th>
<th>$B_1^*$</th>
</tr>
</thead>
<tbody>
<tr>
<td>NaCl</td>
<td>0.05127</td>
<td>42.97</td>
</tr>
<tr>
<td>KBr</td>
<td>0.06867</td>
<td>43.26</td>
</tr>
<tr>
<td>KCl</td>
<td>0.04561</td>
<td>29.164</td>
</tr>
<tr>
<td>Na$_2$SO$_4$</td>
<td>0.040945</td>
<td>14.763</td>
</tr>
<tr>
<td>MgCl$_2$</td>
<td>0.01505</td>
<td>41.895</td>
</tr>
<tr>
<td>MgSO$_4$</td>
<td>0.0007</td>
<td>6.346</td>
</tr>
</tbody>
</table>

TABLE 3.3 Values of $B^*$ and $B_1^*$ in Equation (3-20) for the Systems in Figures H.5 and H.6.
and

\[ B_{12} = \frac{C^*}{T - 230} + \frac{C_{11}^{11}}{T} + C_{22}^{11} + C_{33}^{11} \ln T \] (3-19)

Since \( B_{12} \) values are available only at 25°C (Bromley) and 100°C (this study), equations (3-18) and (3-19) were tested in their two adjustable constants form by setting \( B_2^1, B_3^1, C_2^{11} \) and \( C_3^{11} \) equal to zero.

\[ B_{12} T = B_1^1 + B^* T \ln \left( \frac{T - 243}{T} \right) \] (3-20)

and

\[ B_{12} T = \frac{C^* T}{T - 230} + C_{11}^{11} \] (3-21)

The equations are written in this linear form so that they can be tested by plotting \( B_{12} T \) vs \( [T \ln ((T - 243)/T)] \) in equations (3-20) and vs \( [T/(T - 230)] \) for equation (3-21). Both expressions give reasonably good results as demonstrated in Figures H.5 and H.6 for equation (3-20). The straight lines were obtained by regressing all the points, excluding those from the Weast data. The values of the constants \( B^* \) and \( B_1^1 \) are reported in Table 3.3.

Use of equation (3-20) for interpolation purposes is demonstrated in Table H.3. The \( B_{12} \) values at 70°C were obtained from equation (3-20) with \( B^* \) and \( B_1^1 \) calculated using only the \( B_{12} \) (25°C) and the \( B_{12} \) (100°C) values. The large error for MgSO\(_4\) is due to incomplete dissociation figure H.3. Values of \( \gamma_+ \) at 70°C for the electrolytes of Table 3.3 are given in the references presented in Table 3.1. It should be
noted that the $γ_e$ data for KCl, MgCl₂, MgSO₄ and Na₂SO₄ (Snipes et al., 1975) were derived from heat of dilution data, and those for NaCl (Gibbard, 1974) are reported at rounded molalities and temperatures.

B. **Nonaqueous Electrolyte Binary**

**Maximum Molality Applicability**—As shown for the aqueous electrolyte binaries, the Bromley equation in the form of equation (3-15) is used to test the maximum molality range for MeOH electrolyte binaries. The results for the LiBr-MeOH system at 15°C and the LiCl-MeOH system at 60°C (Figures H.7 and H.8) show that the applicability of the Bromley equation is good only up to $I = 3$, for 1-1 electrolytes. For the higher order electrolytes (1-2, 2-1, 2-2, etc.) data are not available. The only data available for higher order electrolytes is for the CaCl₂-MeOH system at 25°C. Because of the scattering and unavailability of the experimental data at low molality (Figure H.9), it is not possible to conclude the maximum molality range for this system. Based on 1-1 electrolytes only the molality limit for MeOH system is set $I = 3$. Other nonaqueous binaries could not be tested since this type of data is not available in literature. It would be expected that the molality range would be even lower than $I = 3$ for solvents with dielectric constants less than that of MeOH.

**Accuracy of the Binary Data Correlation**—The results of some nonaqueous electrolyte binary data correlation are
given in Table H.4. The accuracy of data correlation of such systems is less than the accuracy of the corresponding aqueous electrolytic binaries (Table H.5). In a typical case of the system LiCl-H$_2$O at 60°C, the maximum percent error in DP is 2.0 and the average percent error is 1.0 (Table H.5) whereas for the system LiCl-MeOH at 60°C, the maximum percent error in DP is 15.6% and the average percent error is 7.3 (Table H.4). The correlation of the system CaCl$_2$-MeOH at 25°C, up to m = 2.6 is of poor quality (Table H.4). This is expected for 1-2, 2-1 or higher order electrolytes in MeOH because of the maximum molality limitation.

C. **Isothermal Ternary Data Correlation**

The mean molal activity coefficient data of three isothermal ternary systems are correlated with this model (Table H.6), Figure H.9 to H.17. The systems HCl-H$_2$O-MeOH at 25°C and NaCl-H$_2$O-MeOH at 25°C have average percent errors in $\gamma_\pm$ of 1.4 and 7.4, respectively (Figures H.13 to H.17). The correlation of the system HCl-H$_2$O-EtOH at 25°C is good up to m = 2.0 and EtOH concentration (HCl free) <9% (Figures H.10 and H.11). The results are of poor quality for the same system at $x'_{EtOH} = 0.5$, especially when m > 0.5.

The vapor-liquid equilibrium data of five isothermal ternary systems are correlated (Table H.7) with this model. Again, three objective functions [Equations (2-19), (2-20) and (2-21)] are applied for the data reduction. In general, objective
function #2 gives the best results. In Table H.7 results are given for the correlation up to \( m = 3 \) and also for the higher molality range. The overall performance of the model in correlating the ternary VLE data is of good quality within the limited range of molality (Figures H.17 to H.22). The molality applicability decreases to even less than \( I = 3 \) with an increase in temperature, as in the case of the LiCl-\( \text{H}_2\text{O-MeOH} \) system at 60°C, the fit is good only up to \( m = 2.0 \).

The ternary parameters \( \delta_{123} \) and \( B_{123} \) obtained by the ternary correlation, indicates that these cannot be considered temperature independent. Therefore this model is not applied to isobaric ternary systems.
3.4 Discussion

As already shown, only three DP-m points were used in evaluating $B_{12}$ for 1-2, 2-1 and 2-2 electrolytes. In addition, no experimental $\gamma^\pm$ data at 100°C are available for such electrolytes as in the case for 1-1 types, where good results are obtained from three points only (Table 3.2). Figures H.5 and H.6 and Table H.9 demonstrate, however, that the $B_{12}$ values obtained from the Weast data are very close to those obtained by extrapolation of the data in the range from 25 to 80°C. The closeness of the $\gamma^\pm$ values obtained from these two $B_{12}$ (100°C) values is depicted in Table H.10 for $\text{MgCl}_2$ with a maximum difference of 5.8%. For the $\text{Na}_2\text{SO}_4$ system, where the fractional difference between the two $B_{12}$ (100°C) values is the largest, the maximum difference in $\gamma^\pm$ is 5.7%. Therefore it is suggested that, in addition to the 1-1 electrolytes, reasonably accurate $\gamma^\pm$ values can be calculated for 1-2, 2-2, and 2-1 types from $B_{12}$ values obtained using three data points from the Weast compilation.

Table H.3 demonstrates that use of the $B_{12}$ values at 25 and 100°C, along with equation (3-20), can lead to reasonably accurate estimates of $\gamma^\pm$ and DP values at intermediate temperatures. However, when the same approach was used to evaluate the derivative ($dB_{12}/dT$), needed to calculate apparent relative molal enthalpies ($\phi L$) and relative partial molal enthalpies ($\overline{L}_2$) for four individual salts, $\text{NaCl}$, $\text{KCl}$, $\text{Na}_2\text{SO}_4$ and $\text{MgCl}_2$
Figure 3.3 Contribution of Different Terms to $\ln \gamma_i$ for the System LiCl-H$_2$O-EtOH at 25°C in Model II.
with water at 100°C the typical average error was about 50%.

This failure becomes apparent from Figures H.5 and H.6. While equation (3-20) is valid for interpolation purposes, it does not provide reliable values for the slope $\frac{d\beta}{dT}$. The calculated values for $\phi_L$ and $\overline{L}_2$, however, were in the right direction, but lower than the experimental ones. The expressions used to calculate these quantities are given by Bromley (1973).

When the Bromley equation was applied to nonaqueous electrolytic binaries, the maximum molality range is $\leq 3$, also the binary data reduction is less accurate for such binaries. This is expected since the empirical constants in the original Bromley equation (C-1) were obtained by applying aqueous electrolytic binary data only. However, considering the simplicity of this equation, the results for nonaqueous mixtures are of acceptable quality. The temperature dependency of such systems can also be established by equation (3-20), as shown by Tomasula and Tassios (1980) for the electrolyte-MeOH binaries.

The isothermal ternary data correlation of the VLE and $\gamma_\pm$ is of good quality for $m \leq 3.0$. An investigation of the contribution of different terms in a ternary system indicates that the salting in effect of the Debye-Hückel term in the Bromley equation (Figure 3.3) is compensated by the additional salting out term at low EtOH concentration (Figure 3.4) and
Figure 3.4 Contribution of the Salting-Out Term to $\ln \gamma_i$ for the System LiCl-H$_2$O-EtOH at 25°C in Model II
by the additional part of the Bromley equation at higher EtOH concentration (Figure 3.3). It is important to note that for ternary data correlation both binary and ternary experimental data are used.
3.5 Conclusions

A method for the correlation of the DP-m data for 53 aqueous electrolytes at 100°C (Weast, 1969), and the evaluation of $\gamma_\pm$ values for these electrolytes, is presented. A procedure for the estimation of $\gamma_\pm$, $\Delta P$, $\phi L$ and $L_2$ in the temperature range of 25-100°C for these electrolytes is also presented. While reasonably good results are obtained for $\gamma_\pm$ and $\Delta P$, $\phi L$ and $L_2$ values are smaller than the experimental ones by about 50%.

The binary Bromley equation is applied to nonaqueous electrolytic binaries and also has been extended to ternary systems of electrolyte in mixed solvents. The correlation of isothermal nonaqueous binary and ternary data is of good quality; however, the maximum concentration range for such systems is less than the range for aqueous electrolytic binaries.
CHAPTER 4

A COMPARATIVE STUDY OF TWO MODELS IN CORRELATING
AND PREDICTING BINARY/TERNARY $\gamma_\pm$ AND VLE
DATA IN ELECTROLYTIC SOLUTIONS

ABSTRACT

Model I is superior to Model II in correlating binary nonaqueous electrolytic and ternary electrolytic mixtures. However, Model II can be used to predict $\gamma_\pm$ using DP vs m data only in a binary mixture more accurately than Model I. Model II is limited only to isothermal ternary data correlation whereas Model I can be applied to predict and/or correlate isothermal or isobaric ternary data.
In principle, the two models presented in Chapters 2 and 3 are similar but consist of different forms of the expressions to represent various interactive forces in the liquid solution. Both the models have the Debye-Hückel equation and an intermediate term also called the transition term represents the change of magnitude of electrostatic forces from the dilute solution to the concentrated solution. The NRTL term has been included in both models, but it represents different molecular interactions in the two models. In Model I the NRTL term describes ion-solvent and solvent-solvent molecular interactions [Equation (2-3)], whereas in Model II ion-solvent interactions are represented by a term: \( B \frac{1}{2} \) [Equation (3-2)] and the solvent-solvent molecular interactions are by the NRTL-S [Equation (3-4)]. Also, in Model II an additional salting out term is used [Equation (3-5)], which is not needed with Model I.

A. Binary Data Correlation

Both models simplify to the original NRTL expression [Equation (2-7)] for a solvent-solvent binary. Electrolyte-solvent binary data reduction require two parameters in Model I, i.e. \( G_{\pm i}; Z_{\pm i} \) or \( \Delta g_{Ai}; \Delta g_{Bi} \) and one parameter '\( B_{\pm i} \)' with Model II. In general, the fit of aqueous electrolyte binaries is better with Model II than Model I, Tables G.4 and H.5. But the correlation of nonaqueous electrolyte binaries shows the reverse trend, Tables G.5 and H.4. However, on the overall
analysis of binary data evaluation, it is concluded that both models can be applied successfully. Model I can be used up to \( I = 6 \) for electrolyte MeOH binaries and up to even higher molality ranges for aqueous electrolyte binaries, Table 2.3, whereas Model II is limited to \( I = 3 \) for MeOH-electrolyte binaries and \( I = 6 \) for aqueous electrolyte mixtures. The binary parameters in Model I are considered temperature independent within a 30 to 40°C temperature range, but in Model II, the temperature dependency of the binary parameter \( B_{1i} \) is represented by a two parameter expression, Equation (3-20).

This indicates the applicability of Model I to isothermal and isobaric systems without any alterations in the Model itself. The biggest advantage of Model II is that it requires only three data points (DP vs m) to find the optimum value of a single parameter \( B_{1i} \) in a binary mixture which can lead to reliable prediction of \( \gamma_{\pm} \) data for the whole concentration range. The use of three typical data points (DP vs m) with Model I is too small for the evaluation of two parameters in a binary mixture and also the parameters obtained with three points only, cannot be expected to predict \( \gamma_{\pm} \) with reasonable accuracy.

B. Ternary Data Prediction and Correlation

Model I requires only binary parameters for the ternary VLE and \( \gamma_{\pm} \) data prediction. The binary parameters are obtained by three respective binary data correlation, Tables
2.4 and 2.5. Model II is good only for binary/ternary data correlation. Prediction of a ternary mixture is not possible, because of two ternary parameters $B_{123}$ and $\delta_{123}$ which should be obtained by ternary data reduction. Also, in Model II binary and ternary parameters ($B_{1i}$, $B_{123}$, $\delta_{123}$) are temperature dependent, hence its applicability is limited to only isothermal data. Model I can be used to correlate binary or ternary data individually whereas with Model II both binary and ternary data are used for ternary data correlation.

Finally, Model I has the possibility of extension to multicomponent systems containing more than two solvents and one electrolyte. In a multicomponent mixture, only binary parameters are required with Model I. The extension of Model II to multicomponent mixtures will be a tedious task.
APPENDIX A

EXPRESSIONS FOR THE ACTIVITY COEFFICIENT OF THE SOLVENT AND THE MEAN ACTIVITY COEFFICIENT OF AN ELECTROLYTE IN A BINARY MIXTURE FOR MODEL I
In a binary mixture, the activity coefficients are a combination of an extended form of the Debye-Hückel equation and the modified NRTL equation proposed by Cruz and Renon (1978).

\[
\ln \gamma_\pm = \ln \gamma_\pm, \text{Ext. D.H.} + \ln \gamma_\pm^*, \text{NRTL} \tag{A-1}
\]

\[
\ln \gamma_i = \ln \gamma_i, \text{Ext. D.H.} + \ln \gamma_i^*, \text{NRTL} \tag{A-2}
\]

Gronwall, LaMer and Sandved (1928) extended the Debye-Hückel equation to higher order terms for symmetrical valence type electrolytes

\[
\ln \gamma_\pm^* = -\frac{(\varepsilon \bar{Z})^2}{2DkT} \frac{\kappa}{1 + \kappa a} + \sum_{m=1}^{\infty} \left( \frac{\varepsilon \bar{Z}^2}{DkT} \right)^{2m+1} \left( \frac{1}{2} X_{2m+1}(\kappa a) - 2m Y_{2m}(\kappa a) \right) \tag{A-3}
\]

where \( X \) and \( Y \) are functions of \( \kappa a \) and

\[
\kappa = \sqrt{\frac{8\pi Ne^2 \bar{Z}^2 c}{1000 DkT}} \tag{A-4}
\]

The additional higher order terms in equation (A-3) take into account long-range electrostatic forces in the concentrated solution. Further, Gronwall, LaMer and Grief (1931) extended the above theory to unsymmetrical electrolytes. A semi-empirical extended form of the Debye-Hückel equation is proposed in this work which is analogous to those proposed by Gronwall et al. An additional term with the original D.H. term represents electrostatic forces in concentrated electrolytic solutions.

\[
\ln \gamma_\pm \text{Ext.D.H.} = 2.303\left[ -A_{\gamma} \frac{I^{1/2}}{1 + \rho I^{1/2}} + A_{\gamma}^2 \frac{I}{(1 + a I)^n} \right] |Z_+Z_-| \tag{A-5}
\]

where \( \rho, a \) and \( n \) are adjustable parameters.
The expression for the activity coefficient of the solvent is obtained through the excess Gibbs free energy function, as shown below

\[
\frac{G^E}{RT} = \sum N_i \ln \gamma_i^* \, dN_i \tag{A-6}
\]

\[
\ln \gamma_i^* = \ln \gamma_i^* + \ln (0.001 \, \text{vmm} + 1) \tag{A-7}
\]

\[
\ln \gamma_i \, \text{Ext. D.H.} = \frac{\partial G^E / RT}{\partial N_i} \bigg|_{T_i, P_i, N_i \neq i} \tag{A-7A}
\]

The NRTL part in equations (A-1) and (A-2) for the activity coefficients are the same as given by Cruz-Renon (1978)

\[
\ln \gamma_i^*_{\text{NRTL}} = \frac{1}{RT} \frac{V_A}{V} \left[ \frac{X_i^2 Z_i^+}{(X_A G_i^+ + X_i)^2} - Z_i^+ \right] \tag{A-8}
\]

and

\[
\ln \gamma_i_{\text{NRTL}} = \frac{1}{RT} X_A^2 \frac{G_i^+ Z_i^+}{(X_A G_i^+ + X_i)^2} \tag{A-9}
\]

Equations (A-5) and (A-7A) have three known adjustable parameters, \(\rho\), \(a\) and \(n\). Equations (A-5) and (A-7A), when combined with the NRTL equations (A-8) and (A-9), have five parameters, \(\rho\), \(a\), \(n\), \(G_i^+\) and \(Z_i^+\) for a binary mixture. Also, it should be noted that the final form of the \(\ln \gamma_i^* \, \text{Ext. D.H.}\) equation will depend upon the integration of the \(\ln \gamma_i^*\) term. The integration is accomplished by fixing a value of \(n\), which can be an integer or a noninteger. So the first five parameters were reduced to the two NRTL parameters, \(G_i^+\) and \(Z_i^+\), by presetting
the values of ρ, α and n. Secondly, the \( \ln \gamma_1^, \) Ext. D.H. expression is derived by equations (A-6) to (A-7A). Equation (A-5) and the final form of equation (A-7A), when combined with equations (A-8) and (A-9), were used to correlate both binary aqueous electrolyte and nonaqueous electrolyte data. It was found that the best results are obtained (Tables G.4 and G.5) by setting the three adjustable parameters in the extended Debye-Hückel equation to

\[
\begin{align*}
\rho &= 1.0 \\
\alpha &= 1.5/|Z_+Z_-| \\
n &= 1/2
\end{align*}
\]  

(A-10)

When the parameters of equation (A-10) are substituted in equation (A-5) the following form of the expression for the solvent activity coefficient is obtained

\[
\ln \gamma_1^, \ Ext. \ D.H. = \frac{2.303 \ \text{vM}_W}{1000} \left[ \frac{A}{3} I^{1/2} \sigma_1(\rho I^{1/2}) + A \frac{I}{2} \psi_1(aI) \right] \\
|Z_+Z_-| + \ln(0.001 \ \text{vM}_W + 1) - 0.001 \ \text{vM}_W
\]  

(A-11)

where

\[
\sigma_1(\rho I^{1/2}) = \frac{3}{(\rho I^{1/2})^3} [(1 + \rho I^{1/2}) - 2\ln(1 + \rho I^{1/2}) - \frac{1}{(1 + \rho I^{1/2})}]
\]  

(A-12)

and

\[
\psi_1(aI) = \frac{2}{3aI} \left[ \frac{2(aI - 2)}{aI} (1 + aI)^{1/2} + \frac{4}{aI} \frac{(aI - 2)}{(1 + aI)^{1/2}} - 2(1 + aI)^{1/2} \right]
\]  

(A-13)
APPENDIX B

A STEPWISE PROCEDURE FOR THE DEVELOPMENT OF
TERNARY ACTIVITY COEFFICIENTS
FOR MODEL I
MODEL I:

Combination of the Extended Debye-Hückel Equation and the Modified NRTL Equation

B.1—Development of the \( \frac{G}{RT} \) Ext. D. H. expression

The extended Debye-Hückel part of the mean molal activity coefficient developed in Appendix A for a binary mixture has been extended to a ternary mixture containing one electrolyte and two solvents. This is obtained by modifying the Debye-Hückel constant for the solvent mixture. For a ternary mixture, in equation (A-5), the Debye-Hückel constant is

\[
A_y = 1.3246 \times 10^6 \, d^{1/2} \left[ \frac{1}{D} \right]^{3/2}
\]  

where

D and d = Dielectric constant and density of a solvent mixture (electrolyte-free) (Appendix D)

The excess Gibbs free energy function for a ternary mixture can be derived by integrating the expression for \( \ln \gamma^* \) for a ternary system. Combination of equations (A-6), (A-7), (A-5) and (B-1) yields

\[
\frac{G}{RT} \text{ Ext. D. H.} = \sqrt{N_1} \left[ 2.303 \left\{ -A_y \left( \frac{1}{1 + \rho I^{1/2}} \right) + A_y^2 \left( \frac{1}{1 + a I^{1/2}} \right) \right\} \right]
\]

\[
|Z_+ Z_-|dN_1 + \sqrt{N_1} \ln (0.001 \, \nu \, M_w \, \nu + 1)dN_1
\]  

Equation (B-2) can be integrated, term by term, with the following additional equations
\[
\begin{align*}
\bar{m} &= \frac{1000 \, N_1}{N_T \, M_w} \quad \text{(B-3)} \\
N_T &= N_2 + N_3 \quad \text{(B-4)} \\
I &= \frac{1}{2} \, m \sum_k \nu_k z_k^2 = \frac{1}{2} \, m \xi \quad \text{(B-5)} \\
\xi &= \text{a constant} \\
M_w &= \frac{N_2}{N_T} \, M_{w2} + \frac{N_3}{N_T} \, M_{w3} \quad \text{(B-6)} \\
\frac{\partial m}{\partial N_1} &= \frac{1000}{N_T M_w} \quad \text{(B-7)} \\
\frac{\partial I}{\partial m} &= \frac{1}{2} \, \xi \quad \text{(B-8)}
\end{align*}
\]

Integration of different terms is as below

\[
\begin{align*}
\int_0^{N_1} \ln(0.001 \, \nu M_w \, + 1) \, \partial N_1 &= \int_0^m \ln(0.001 \, \nu M_w \, + 1) \frac{\partial N_1}{\partial m} \, \partial m \\
&= \frac{\nu N_T M_w}{1000} \left[ \frac{1000}{\nu M_w} \{ (0.001 \, \nu M_w \, + 1) \ln(0.001 \, \nu M_w \, + 1) - 0.001 \, \nu M_w \} \right] \\
&= \frac{\nu N_T}{N_1} \left[ (0.001 \, \nu M_w \, + 1) \ln(0.001 \, \nu M_w \, + 1) - 0.001 \, \nu M_w \right] \quad \text{(B-9)}
\end{align*}
\]
\[ \nu \int_{0}^{N_1} A_\gamma \frac{|z^+ z^-|^{1/2}}{1 + \rho I^{1/2}} \, dN_1 = \gamma_\gamma^{1/2} \int_{0}^{I} \frac{I^{1/2}}{1 + \rho I^{1/2}} \, dI \]

\[ = \nu A_\gamma \frac{N_{TW}}{1000} |z^+ z^-|^{1/2} \, dI \]

\[ = \nu A_\gamma \frac{N_{TW}}{1000} |z^+ z^-|^{1/2} \, dI \]

\[ = \nu A_\gamma^2 |z^+ z^-| \frac{I}{(1 + aI)^{1/2}} \, dN_1 \]

\[ = \nu A_\gamma^2 |z^+ z^-| \frac{N_{TW}}{1000} \frac{2 \xi}{(1 + aI)^{1/2}} \, dI \]

\[ = \nu A_\gamma^2 |z^+ z^-| \frac{N_{TW}}{1000} \frac{2 \xi}{(1 + aI)^{1/2}} \, dI \]

\[ + \frac{4}{3a^2} \]

Utilizing the change of variables \( m/I = 2/\xi \), and combining equations (B-9 to (B-11), results in the excess Gibbs free energy expression, equation (2-2).

B.2--Development of the \( \frac{G^E}{RT} \) expression

Note: Equations are derived taking into account that \( \gamma^*_z \to 1.0 \) as \( x_1 \to 0.0 \) (Asymmetric Convention).
Renon and Prausnitz (1968) proposed an expression for the excess Gibbs free energy in a multicomponent mixture based on the Non-Random Two Liquid Theory. Since the original NRTL equation applies to mixtures following the symmetric convention, it is converted for ternary mixtures utilizing the asymmetric convention as indicated below

\[
\boxed{g_{\text{NRTL}}^E' = \sum_l \frac{\Sigma_m Z_{ml} x_m}{n G_{nl}}} \quad (B-12)
\]

where,

\[
Z_{ml} = \Delta g_{ml} \frac{G_{ml}}{G_{nl}}
\]

\[
G_{ml} = \text{Exp}[-\alpha_{ml} \frac{\Delta g_{ml}}{RT}] \quad (B-12a)
\]

\[
\Delta g_{ml} = g_{ml} - g_{ll}
\]

\[
\alpha_{ml} = \alpha_{lm} \text{ and } \Delta g_{ml} \neq \Delta g_{lm}
\]

Equation (B-12) can be expanded for a mixture containing electrolyte molecules - 1, cation - A, anion - B, and solvents 2 and 3

\[
g_{\text{NRTL}}^E' = x_A \left[ \frac{x_A Z_{AA} + x_B Z_{BA} + x_2 Z_{1A} + x_2^2 Z_{2A} + x_2^3 Z_{3A}}{x_A G_{AA} + x_B G_{BA} + x_1 G_{1A} + x_2 G_{2A} + x_3 G_{3A}} \right]
\]

\[
+ x_B \left[ \frac{x_A Z_{AB} + x_B Z_{BB} + x_1 Z_{1B} + x_2 Z_{2B} + x_2^3 Z_{3B}}{x_A G_{AB} + x_B G_{BB} + x_1 G_{1B} + x_2 G_{2B} + x_3 G_{3B}} \right]
\]

\[
+ x_1 \left[ \frac{x_A Z_{11} + x_B Z_{11} + x_1 Z_{11} + x_2 Z_{21} + x_3 Z_{31}}{x_A G_{11} + x_B G_{11} + x_1 G_{11} + x_2 G_{21} + x_3 G_{31}} \right]
\]
\[ + x_2 \left( \frac{x_A^2 Z_{A2} + x_B^2 Z_{B2} + X_1^2 Z_{12} + x_2^2 Z_{22} + X_3^2 Z_{32}}{x_A^2 G_{A2} + x_B^2 G_{B2} + X_1^2 G_{12} + x_2^2 G_{22} + X_3^2 G_{32}} \right) \]
\[ + x_3 \left( \frac{x_A^3 Z_{A3} + x_B^3 Z_{B3} + X_1^3 Z_{13} + x_2^3 Z_{23} + X_3^3 Z_{33}}{x_A^3 G_{A3} + x_B^3 G_{B3} + X_1^3 G_{13} + x_2^3 G_{23} + X_3^3 G_{33}} \right) \]  

Equation (B-13) is simplified by setting \( Z_{ll} \) = 0.0 and \( G_{ll} \) = 1.0 based on the original development of the equation (B-12). Cruz and Renon (1972) proposed the following additional assumptions for an electrolytic mixture considering that the energy parameter \( g_{ml} \) increases from low to large numerical values in the following order:

(solvent-ion) < (solvent or electrolyte)-(solvent or electrolyte)<< (electrolyte-ion) or (ion-ion of opposite signs) < (ion-ion of the same signs). On the right sign <<, very large values of \( g_{ml} \) are found, and true local mole fractions are taken equal to zero. Thus

\[ Z_{1A} = Z_{1B} = Z_{A1} = Z_{B1} = Z_{AB} = Z_{BA} = 0.0 \]
\[ G_{1A} = G_{1B} = G_{A1} = G_{B1} = G_{AB} = G_{BA} = 0.0 \]  

(B-14)

For specific ion interaction limitation--

\[ G_{AA} = G_{BB} = Z_{AA} = Z_{BB} = 0.0 \]

And, also

\[ X_{2A} = X_{2B} = X_{3A} = X_{3B} = 1.0 \]
\[ G_{2A} = G_{2B} = G_{2A} = G_{3B} = 1.0 \]  

(B-15)
where

\[ x_{\ell m} = \frac{x_{\ell} G_{\ell m}}{\sum_n x_n G_{nm}} \]  \hspace{1cm} (B-16)

\[ Z_{\ell m} = \Delta g_{\ell m} G_{\ell m} \]  \hspace{1cm} (B-17)

Substitution of equations (B-15) to (B-17) into equation (B-13) yields

\[ g'_{\text{NRTL}} = x_A \left[ \frac{x_2 \Delta g_{2A} + x_3 \Delta g_{3A}}{x_2 + x_3} \right] + x_B \left[ \frac{x_2 \Delta g_{2B} + x_3 \Delta g_{3B}}{x_2 + x_3} \right] \]

\[ + x_1 \left[ \frac{x_2 Z_{21} + x_3 Z_{31}}{x_1 + x_2 G_{21} + x_3 G_{31}} \right] \]

\[ + x_2 \left[ \frac{x_A Z_{A2} + x_B Z_{B2} + x_1 Z_{12} + x_3 Z_{32}}{x_A Z_{A3} + x_B Z_{B2} + x_1 Z_{12} + x_3 Z_{32} + x_2} \right] \]

\[ + x_3 \left[ \frac{x_A Z_{A3} + x_B Z_{B3} + x_1 Z_{13} + x_2 Z_{23}}{x_A Z_{A3} + x_B Z_{B3} + x_1 Z_{13} + x_2 Z_{23} + x_3} \right] \]  \hspace{1cm} (B-18)

\[ E'_{\text{NRTL}} = (N_1 g'_{\text{NRTL}}) = N_A \left[ \frac{N_2 \Delta g_{2A} + N_3 \Delta g_{3A}}{N_2 + N_3} \right] \]

\[ + N_B \left[ \frac{N_2 \Delta g_{2B} + N_3 \Delta g_{3B}}{N_2 + N_3} \right] \]

\[ + N_1 \left[ \frac{N_2 Z_{21} + N_3 Z_{31}}{N_1 + N_2 G_{21} + N_3 G_{31}} \right] \]

\[ + N_2 \left[ \frac{N_A Z_{A2} + N_B Z_{B2} + N_1 Z_{12} + N_3 Z_{32}}{N_A Z_{A2} + N_B Z_{B2} + N_1 Z_{12} + N_3 Z_{32} + N_2} \right] \]

\[ + N_3 \left[ \frac{N_A Z_{A3} + N_B Z_{B3} + N_1 Z_{13} + N_2 Z_{23}}{N_A Z_{A3} + N_B Z_{B3} + N_1 Z_{13} + N_2 Z_{23} + N_3} \right] \]  \hspace{1cm} (B-19)
\[ RT \ln \gamma_A' = \frac{\partial \tilde{G}^E}{\partial \tilde{G}_N} \bigg|_{T,P,N}^{\neq A} \]
\[ = \frac{N_A^{Z_{A2}}}{N_A + N_3} + N_2 \left[ \frac{Z_{A2}}{A_A + B_B + C_C + D_D + E_E} \right] \]
\[ + \frac{(N_A Z_{A2} + N_B Z_{B2} + N_1 Z_{12} + N_3 Z_{32}) G_{A2}}{(N_A G_{A2} + N_B G_{B2} + N_1 G_{12} + N_3 G_{32} + N_2)^2} \]
\[ + \frac{Z_{A13}}{A_A + B_B + C_C + D_D + E_E} \]
\[ - \frac{(N_A Z_{A3} + N_B Z_{B3} + N_1 Z_{13} + N_2 Z_{23}) G_{A3}}{(N_A G_{A3} + N_B G_{B3} + N_1 G_{13} + N_2 G_{23} + N_3)^2} \]
\[ \text{(B-20)} \]

\[ \lim_{N_A \to 0} RT \ln \gamma_A' = \frac{N_A^{Z_{A2}}}{N_A + N_3} + \frac{N_A Z_{A2}}{N_A + N_3} \]
\[ - \frac{N_A N_2 N_3 Z_{32} G_{A2}}{(N_3 G_{32} + N_2)^2} + \frac{N_A N_3 Z_{A3}}{(N_2 G_{23} + N_3)^2} \]
\[ - \frac{N_A N_2 N_3 Z_{A3} Z_{23}}{(N_2 G_{23} + N_3)^2} \]
\[ \text{(B-21)} \]

Similarly
\[ \lim_{N_B \to 0} RT \ln \gamma_B' = \frac{N_B^{Z_{B2}}}{N_B + N_3} + \frac{N_B Z_{B2}}{N_B + N_3} \]
\[ - \frac{N_B N_2 N_3 Z_{32} G_{B2}}{(N_3 G_{32} + N_2)^2} + \frac{N_B N_3 Z_{B3}}{(N_2 G_{23} + N_3)^2} \]
\[ - \frac{N_B N_2 N_3 Z_{B3} Z_{23}}{(N_2 G_{23} + N_3)^2} \]
\[ \text{(B-22)} \]
\[ E(\text{ternary}) = G_{\text{NRTL}} - N_A \lim_{N_A \to 0} RT \ln \gamma_A' - N_B \lim_{N_B \to 0} RT \ln \gamma_B' \]  

Substituting equations (B-19), (B-21) and (B-22) in equation (B-23) results in the following expression for \[ E(\text{ternary}) = N_1 \left[ \frac{N_2 Z_{21} + N_3 Z_{31}}{N_1 + N_2 G_{21} + N_3 G_{31}} \right] \]

\[ + N_2 \left[ \frac{N_A Z_{A2} + N_B Z_{B2} + N_1 Z_{12} + N_3 Z_{32}}{N_A G_{A2} + N_B G_{B2} + N_1 G_{12} + N_3 G_{32} + N_2} \right] \]

\[ + N_3 \left[ \frac{N_A Z_{A3} + N_B Z_{B3} + N_1 Z_{13} + N_2 Z_{23}}{N_A G_{A3} + N_B G_{B3} + N_1 G_{13} + N_2 G_{23} + N_3} \right] \]

\[ - N_2 \left[ \frac{N_A Z_{A2} + N_B Z_{B2}}{N_3 G_{32} + N_2} \right] - N_3 \left[ \frac{N_A Z_{A3} + N_B Z_{B3}}{N_2 G_{23} + N_3} \right] \]

\[ + N_2 N_3 Z_{32} \left[ \frac{N_A G_{A2} + N_B G_{B2}}{(N_3 G_{32} + N_2)^2} \right] \]

\[ + N_2 N_3 Z_{23} \left[ \frac{N_A G_{A3} + N_B G_{B3}}{(N_2 G_{23} + N_3)^2} \right] \]  

(B-24)

Considering macroscopic electrical neutrality

\[ N_A \nu_B = N_B \nu_A \]  

(B-25)

Substituting equation (B-25) into equation (B-24) gives

\[ E(\text{ternary}) = N_1 \left[ \frac{N_2 Z_{21} + N_3 Z_{31}}{N_1 + N_2 G_{21} + N_3 G_{31}} \right] \]

\[ + N_2 \left[ \frac{N_A (Z_{A2} + \nu_B Z_{B2}) + N_1 Z_{12} + N_3 Z_{32}}{N_A (G_{A2} + \nu_A G_{B2}) + N_1 G_{12} + N_3 G_{32} + N_2} \right] \]

\[ + N_3 \left[ \frac{N_A (Z_{A3} + \nu_B Z_{B3}) + N_1 Z_{13} + N_2 Z_{23}}{N_A (G_{A3} + \nu_A G_{B3}) + N_1 G_{13} + N_2 G_{23} + N_3} \right] \]
Let

\[ G_{\pm 2} = G_{A2} + \frac{v_B}{v_A} G_{B2} \]

\[ G_{\pm 3} = G_{A3} + \frac{v_B}{v_A} G_{B3} \]

\[ Z_{\pm 2} = Z_{A2} + \frac{v_B}{v_A} Z_{B2} \]

\[ Z_{\pm 3} = Z_{A3} + \frac{v_B}{v_A} Z_{B3} \]  \hspace{1cm} (B-27)

Combining equations (B-25) and (B-27) and setting \( N_\perp = 0.0 \) for the case of complete dissociation, leads to equation (2-3) which is the final expression for \( \frac{G^E}{RT} \) used in this study.

B.3--Development of the ternary \( \ln \gamma_\perp, \ln \gamma_2 \) and \( \ln \gamma_3 \) expressions

The total excess Gibbs free energy function in this model
is obtained by combining equations (2-2) and (2-3). The activity coefficient expressions are obtained by the appropriate differentiation of the total excess Gibbs free energy expression, equations (1-27) to (1-29). The differentiation of the Debye-Hückel and NRTL terms of the expression have been performed separately as shown below.

B.3-I--Debye-Hückel equation--

\[ \nu \ln \gamma^*_{\text{Ext.D.H.}} = \frac{3}{3N_1} \frac{G^E_{\text{ternary}}}{RT} \text{Ext.D.H.} \quad T, P, N_2, N_3 \tag{B-28} \]

Since, \( \frac{G^E_{\text{ternary}}}{RT} \text{Ext.D.H.} \) was obtained by the integration of \( \ln \gamma^*_{\text{Ext.D.H.}} \), differentiation of this excess Gibbs free energy function gives the same expression for \( \ln \gamma^*_{\text{Ext.D.H.}} \), equation (2-13).

For solvents (2) and (3)

\[ \ln \gamma^*_2 \text{Ext.D.H.} = \frac{3}{3N_2} \frac{G^E}{RT} \text{Ext.D.H.} \quad T, P, N_1, N_3 \]

Equation (2-2) is differentiated term by term by utilizing the change of variables, i.e. \( m = \frac{2}{z} \)

I term = \( \frac{3}{3N_2} \left[ -2.303\nu |Z_+Z_-| \frac{2}{\xi} \frac{N T M}{1000} A_y \frac{2}{3} (z + z_1^1/2)^2 
- 2(z_1 + z_1^1/2) + \ln(z_1 + z_1^1/2) + \frac{3}{2} \right] \)

I term = \( -2303 \frac{\nu |Z_+Z_-|}{1000} \frac{2}{\xi} \frac{N T M}{\rho ^3} \frac{2}{3} \left(z + z_1^1/2 \right)^2 - 2 \left(z + z_1^1/2 \right) 
+ \ln(1 + z_1^1/2) + \frac{3}{2} \frac{3A_y}{3N_2} + A_y \frac{2}{\rho ^3} \frac{1}{2} \left(z + z_1^1/2 \right)^2 \)
\[-2(1 + \rho I^{1/2}) + \ln(1 + \rho I^{1/2}) + \frac{3}{2} \frac{\partial}{\partial N_2}(N_T M_w) \]

\[+ N_T M_w A \gamma \rho^2 \frac{\partial}{\partial I}\left\{ \frac{1}{2}(1 + \rho I^{1/2})^2 - 2(1 + \rho I^{1/2}) \right\} \]

\[+ \ln(1 + \rho I^{1/2}) + \frac{3}{2} \frac{\partial I}{\partial m} \frac{\partial m}{\partial N_2} \]  \hspace{1cm} (B-29)

from equations (B-3) and (B-6)

\[\frac{\partial m}{\partial N_2} = \frac{m M_w}{N_T M_w} \]  \hspace{1cm} (B-30)

\[\frac{\partial (N_T M_w)}{\partial N_2} = M_w \]  \hspace{1cm} (B-31)

Simplifying equation (B-29) and substituting equations (B-30) and (B-31) results in the following expression for term I

\[I \text{ term } = 2.303 \frac{\nu m}{1000} \left|Z_+Z_-\right|[A_\gamma I^{1/2}] \frac{1}{(\rho I^{1/2})^3}(1 + \rho I^{1/2}) \]

\[\quad - 2 \ln(1 + \rho I^{1/2}) - \frac{1}{(1 + \rho I^{1/2})} - N_T M_w I^{1/2} \]

\[\frac{1}{(\rho I^{1/2})^3}\left\{ \frac{1}{2}(1 + \rho I^{1/2})^2 - 2(1 + \rho I^{1/2}) + \ln(1 + \rho I^{1/2}) \right\} \]

\[+ \frac{3}{2} \frac{\partial A_\gamma}{\partial N_2} \]  \hspace{1cm} (B-32)

II term = \[\frac{\partial}{\partial N_2}[2.303 \frac{\nu |Z_+Z_-|}{1000} \frac{2}{\xi} N_T M_w A_\gamma \left\{ 2(a I - 2) \right\} 3a^2 + (1 + a I)^{1/2} + \frac{4}{3a^2}] \]
\[ = 2.303 \sqrt{\frac{Z_+ Z_-}{1000}} \frac{2}{\zeta} \left[ \frac{2(aI - 2)}{3a^2} + (1 + aI)^{1/2} + \frac{4}{3a^2} N_T M_w \right] \]

\[ \frac{\partial A_x^2}{\partial N_2} + A_x^2 \left( \frac{2(aI - 2)}{3a^2} + (1 + aI)^{1/2} + \frac{4}{3a^2} \right) \frac{\partial (N_T M_w)}{\partial N_2} \]

\[ + A_y^2 N_T M_w \left( \frac{2(aI - 2)}{3a^2} + (1 + aI)^{1/2} + \frac{4}{3a^2} \right) \frac{\partial I}{\partial m} \frac{\partial m}{\partial N_2} \]

**II term**

\[ = 2.303 \left( \frac{vm}{1000} \right) \left| \frac{Z_+ Z_-}{M_w} \right| \frac{\partial I}{\partial I} \frac{2(aI - 2)}{aI} + (1 + aI)^{1/2} \]

\[ + \frac{4}{aI} - \left( \frac{aI - 2}{aI} - 2(1 + aI)^{1/2} \right) \]

\[ + N_T M_w \left( \frac{2(aI - 2)(1 + aI)^{1/2}}{aI} + \frac{4}{aI} \right) \frac{\partial I}{\partial I} \frac{\partial A_y}{\partial N_2} \]

(B-33)

**III term**

\[ = \frac{\partial}{\partial N_2} \left[ N_T \left( 0.001 \frac{vmM_w}{vmM_w + 1} \right) \ln(0.001 \frac{vmM_w + 1}{vmM_w + 1}) \right. \]

\[ - 0.001 \frac{vmM_w}{vmM_w} \]}

\[ = 0.001 \frac{vmM_w + 1}{vmM_w + 1} - 0.001 \frac{vmM_w}{vmM_w} \frac{\partial N_T}{\partial N_2} \]

\[ + N_T \left( \ln(0.001 \frac{vmM_w + 1}{vmM_w + 1}) \frac{\partial}{\partial N_2} (0.001 \frac{vmM_w + 1}{vmM_w + 1}) \right) \]

\[ + (0.001 \frac{vmM_w + 1}{vmM_w + 1}) \frac{\partial}{\partial N_2} \ln(0.001 \frac{vmM_w + 1}{vmM_w + 1}) - 0.001 \frac{vmM_w}{vmM_w} \]

(B-34)

The combination of equations (B-32), (B-33) and (B-34) yields equation (2-16). A similar approach leads to the expression for \( \ln y^3_{\text{Ext.D.H.}} \).

In equation (2-16) \( \sigma^1_1(\rho I^{1/2}) \) and \( \psi^1_1(aI) \) are given by
\[ \sigma_1 \left( \rho \right)^{1/2} = \frac{2}{\left( \rho \right)^{1/2}} \left[ \frac{1}{2} \left( 1 + \rho \right)^2 - 2(1 + \rho) + \ln(1 + \rho) \right] + \frac{3}{2} \]  

\[ \psi_1 \left( a \right) = \frac{2}{3a} \left[ \frac{2(a - 2)}{a} \right] \left( 1 + a \right)^{1/2} + \frac{4}{a} \]  

B.3.II--NRTL equation

\[ \ln \gamma^*_\pm = \frac{\partial}{\partial N_A} \left[ \frac{G^E(\text{ternary})}{RT} \right]_{NRTL} \left| \begin{array}{c} T, \rho, N_2, N_3 \\ \end{array} \right. \]  

Assuming complete dissociation of the electrolyte

\[ N_A = \nu_A N_1 \]  

\[ \frac{\partial N_A}{\partial N_1} = \nu_A \]  

Combining equations (B-37) and (B-39) yields

\[ \nu \ln \gamma^*_\pm = \nu_A \ln \gamma_A, \text{NRTL} \]  

\[ \ln \gamma^*_\pm \] can be obtained by differentiating equation (B-28) with respect to \( N_A \)

\[ \ln \gamma_A = \frac{1}{RT} \left[ \frac{N_2 \left( N_{A2} Z_{22}^0 + N_{32} G_{32} Z_{22}^0 + N_{23} Z_{23}^0 - N_{A2} Z_{22}^0 Z_{22}^0 - N_{32} Z_{23}^0 \right)}{(N_{A2} + N_{32} + N_2)^2} \right. \]

\[ + \frac{N_3 \left( N_{A3} Z_{33}^0 + N_{23} G_{23} Z_{33}^0 + N_{33} Z_{33}^0 - N_{A3} Z_{33}^0 Z_{33}^0 - N_{23} Z_{23} G_{33} \right)}{(N_{A3} + N_{23} + N_3)^2} \]

\[ - \left\{ \frac{N_{A2} Z_{22}^0}{(N_{A3} G_{32} + N_2)^2} + \frac{N_{A3} Z_{33}^0}{(N_{A2} G_{23} + N_3)^2} \right\} \]
Combining equations (B-37) and (B-41) and converting moles to the mole fraction leads to equation (2-14).

The activity coefficient of the solvent is obtained by differentiating \( \frac{G^E}{RT} \) with respect to \( N_2 \)

\[
\ln \gamma_{2,\text{NRTL}} = \frac{\partial}{\partial N_2} \left[ \frac{G^E(\text{ternary})}{RT} \right]_{T,P,N_A,N_B,N_3}
\]

\[
= \frac{1}{RT} \left[ \frac{NAZ_{\pm 2} + N_3Z_{32}}{(NA^G_{\pm 2} + N_3G_{32} + N_2)} - \frac{N_2(N_AZ_{\pm 2} + N_3Z_{32})}{(NA^G_{\pm 2} + N_3G_{32} + N_2)^2} \right]
\]

\[
+ \frac{N_3Z_{23}}{(NA^G_{\pm 3} + N_2G_{23} + N_3)} - \frac{N_A(N_AZ_{\pm 3} + N_2Z_{23})G_{23}}{(NA^G_{\pm 3} + N_2G_{23} + N_3)^2}

- N_A \left\{ \frac{Z_{\pm 2}}{(N_3G_{32} + N_2)^2} - \frac{N_2Z_{\pm 2}}{(N_3G_{32} + N_2)^2} - \frac{N_3Z_{\pm 3}G_{23}}{(N_2G_{23} + N_3)^2} \right\}
\]

\[
+ N_3 \left\{ \frac{Z_{32}G_{\pm 2}}{(N_3G_{32} + N_2)^2} + \frac{Z_{23}G_{\pm 3}}{(N_2G_{23} + N_3)^2} \right\}
\]

\[
-N_3 \left\{ \frac{2Z_{32}G_{\pm 2}}{(N_3G_{32} + N_2)^3} + \frac{2Z_{23}G_{\pm 3}G_{23}}{(N_2G_{23} + N_3)^3} \right\}
\]  (B-42)

Equation (B-42) can be simplified in the form of equation (2-17). Utilizing a similar approach an expression for \( \ln \gamma_{3,\text{NRTL}} \) is obtained.
APPENDIX C

A STEPWISE PROCEDURE FOR THE DEVELOPMENT
OF TERNARY ACTIVITY COEFFICIENT
EXPRESSIONS FOR MODEL II
MODEL II: Combination of the Bromley Equation; The Simplified NRTL Equation and the Salting Out Term

Bromley (1973) presented a generalized analytic correlation for mean activity coefficients of electrolytes in binary aqueous electrolytic solutions.

\[
\ln \gamma_x = 2.303 \left[ -A_\gamma |Z_+Z_-| \frac{I^{1/2}}{1 + \rho I^{1/2}} + \frac{(B_O - B_{1i}) I}{(1 + a I)^\eta} + B_{1i} I + C I^2 \right] \quad (C-1)
\]

where

\[
A_\gamma = \text{Debye-Hückel constant} \quad \text{(Appendix-L)}
\]

\[
I = \frac{1}{2} \sum m_k Z_k^2 \quad \text{(C-1-A)}
\]

\[
\eta, \rho, a, B_O, B_{1i} \text{ and } c: \text{ adjustable parameters}
\]

On the basis of a comprehensive study using data for \(\gamma_x\), but also osmotic coefficients and other related thermodynamic properties at 25°C, as well as at other temperatures up to 200°C, Bromley concluded that reasonable agreement is obtained by using

\[
\eta = 2
\]

\[
a = 1.5/|Z_+Z_-|
\]

\[
c = 0.0
\]

\[
\rho = 1.0
\]

\[
(B_O - B_{1i}) = (0.06 + 0.6B_{1i})|Z_+Z_-| \quad \text{(C-2)}
\]

in equation (C-1).

Substitution of (C-2) into (C-1) simplifies this equation in terms of only one adjustable parameter 'B_{1i}' per binary. When the above equation was applied to correlate non-aqueous
electrolytic binary with the same constants, the fit was good. Considering the simplicity and success of the above equation in correlating binary systems and the success of the NRTL equation in correlating solvent-solvent binaries, Model II proposes to combine them with an additional ternary salting out term for the correlation of electrolyte (1) - solvent (2) - solvent (3) ternary systems. The following procedure is followed in developing Model II:

C.1: Development of the $\frac{G^E_{(ternary)}}{RT}$ expression

Equation (C-1) with (C-2) is first integrated to derive

$$\frac{G^E_{(binary)}}{RT} = 2.303v \left[ \frac{N_1}{I} \left\{ -A_y \left| Z^+Z^- \right| \frac{1^{1/2}}{1 + \rho I^{1/2}} \right\} dN_1 \right.$$

$$+ \left\{ \frac{N_1}{I} (0.06 + 0.6B_{1i}) \right\} dN_1$$

$$+ \int \left\{ \frac{B_{1i}}{I} dN_1 \right\}$$

$$+ \left\{ \frac{N_1}{I} \ln(0.001vM_w + 1) \right\} dN_1$$

(C-3)

Equation (C-3) can be integrated term by term with equations (B-3) to (B-8)

$$I \text{ term} = 2.303v \int \left[ -A_y \left| Z^+Z^- \right| \frac{1^{1/2}}{1 + \rho I^{1/2}} \frac{\delta N_1}{\delta m} \frac{\delta m}{\delta I} \delta I \right.$$
I term = $-2.303 v A \frac{N_i M_w}{1000} \frac{m}{I} \left[ \frac{1}{2} \left( 1 + \rho I^{1/2} \right)^2 \right.$

$- 2(1 + \rho I^{1/2}) + \ln(1 + \rho I^{1/2}) + \frac{3}{2}] \left| Z_+ Z_- \right| \tag{C-4}$

II term = $2.303 v \int \frac{I (0.06 + 0.06 B_{li})^I}{(1 + aI)^2} \frac{\delta N_i}{\delta m} \frac{\delta m}{\delta I} \delta I$

II term = $2.303 v \frac{N_i M_w}{1000} \left| Z_+ Z_- \right| \frac{0.06 + 0.6 B_{li}}{a^2} \left[ \ln(l + aI) + \frac{1}{(l + aI) - 1} \right]

III term = $2.303 v \int_{B_{li}}^{I} \frac{I}{\rho I^{1/2}} \frac{\delta N_i}{\delta m} \frac{\delta m}{\delta I} \delta I$

III term = $2.303 v \frac{N_i M_w}{1000} \frac{m}{I} B_{li} \frac{I_2^{2}}{2} \tag{C-6}$

IV term = $v \int_{0}^{1} \ln(0.001 v M_w + 1) \frac{dN_i}{\delta m} \delta m$

$= N_i \left[ (0.001 v M_w + 1) \ln(0.001 v M_w + 1) \right.

$- 0.001 v M_w \right] \tag{C-7}$

Combining equations (C-4) to (C-7) leads to the following expression for

$$\frac{G_E(binary)}{RT}_{Bromley} = 2.303 v \frac{N_i M_w}{1000} \frac{m}{I} \left[ -A_y \left| Z_+ Z_- \right| \right.$$

$$\left. \frac{1}{\rho^3} \left\{ \frac{1}{2} \left( 1 + \rho I^{1/2} \right)^2 - 2(1 + \rho I^{1/2}) + \ln(1 + \rho I^{1/2}) \right. \right.$$}

$$\left. + \frac{3}{2} \right\} + \frac{0.06 + 0.6 B_{li}}{a^2} \left| Z_+ Z_- \right|$$

$$\left. \left\{ \ln(1 + aI) + \frac{1}{(1 + aI) - 1} \right\} + \frac{B_{li} I^2}{2} \right]$$
In a binary aqueous/nonaqueous electrolyte solution, equation (C-8) can be used to derive the expression for the activity coefficient of the solvent, by equation (1-27)

\[
\ln \gamma_{i,\text{Bromley}}^\text{(binary)} = 2.303 \frac{\mu M}{1000} M_w^{1/2} \gamma^{1/2} A_\gamma \frac{1}{3} (\sigma I^{1/2}) |Z_+ Z_-| \\
- (0.06 + 0.6 B_{li}) \frac{I}{2} \psi (aI) |Z_+ Z_-| - B \frac{I}{2} \\
+ \ln (0.001 \mu M_w + 1) - 0.001 \mu M_w
\] (C-9)

Equation (C-8) is modified for a ternary system by the appropriate substitution of \( B \) for \( B_{li} \) and the Debye-Hückel constant for mixture. The most important conditions to be satisfied are

\[
\begin{align*}
\frac{C}{RT} \text{Bromley} \bigg|_{\text{ternary}} & = \frac{C}{RT} \text{Bromley} \bigg|_{\text{binary}} \\
\text{Lim}_{N_2 \to 0} & \text{ or } \text{Lim}_{N_3 \to 0}
\end{align*}
\] (C-10)

This requires--

\[
\begin{align*}
\text{Lim}_{N_2 \to 0} B & = B_{12} \\
\text{Lim}_{N_2 \to 0} A_\gamma (N_2, N_3, d_2, d_3, D_2, D_3, T) & = A_\gamma (d_2, D_2, T)
\end{align*}
\] (C-11)

or

\[
\begin{align*}
\text{Lim}_{N_3 \to 0} B & = B_{13} \\
\text{Lim}_{N_3 \to 0} A_\gamma (N_2, N_3, d_2, d_3, D_2, D_3, T) & = A_\gamma (d_3, D_3, T)
\end{align*}
\] (C-11)
The above constraints lead equation (C-8) to
\[ \frac{G^E_{\text{ternary}}}{RT} \text{ Bromley} \]

Note: equation (3-2) involves no solvent-solvent interaction parameters.

C.2—Development of the \[ \frac{G^E_{\text{ternary}}}{RT} \] Expression

The NRTL equation developed in model I, equation (2-3) is simplified further by considering that the NRTL equation in Model II accounts only for solvent-solvent interactions. Hence, assuming that

\[ X_{A2} = X_{B2} = X_{A3} = X_{B3} = 1.0 \] \hspace{1cm} (C-12)

in equation (2-3) then,

\[ \Delta g_{A2} = \Delta g_{A3} = \Delta g_{B2} = \Delta g_{B} = 0.0 \]

\[ G_{A2} = G_{B2} = G_{A3} = G_{B3} = 1.0 \]

\[ G_{\pm 2} = G_{\pm 3} = \frac{v}{v_A} \] \hspace{1cm} (C-13)

\[ Z_{A2} = Z_{B2} = Z_{A3} = Z_{B3} = 0.0 \]

\[ Z_{\pm 2} = Z_{\pm 3} = 0.0 \] \hspace{1cm} (C-14)

\[ \frac{G^E_{\text{ternary}}}{RT}_{\text{NRTL-S}} = \frac{1}{RT} \left[ \frac{N_A N_2 N_3 Z_{32}}{N_A v_A + N_2 + N_3 G_{32}} + \frac{N_A N_2 N_3 Z_{23}}{N_A v_A + N_3 + N_2 G_{23}} \right] \]

\[ + \frac{v}{v_A} N_A N_2 N_3 \left[ \frac{Z_{32}}{(N_3 G_{32} + N_2)^2} + \frac{Z_{23}}{(N_2 G_{23} + N_3)^2} \right] \] \hspace{1cm} (C-15)
Equation (C-15) can be converted in terms of mole fractions, equation (3-4).

C.3--Development of the Salting-Out Term for a Ternary Mixture

It is well known that the addition of an electrolyte in mixed solvents, causes salting-out of one of the solvents. Combination of the Bromley equation and the simplified NRTL equation alone is not enough to account for the salting-out effect, therefore an additional salting-out term was sought. Different theories have been proposed specifically by Debye-McAulay (1925), Butler (1929) and Born (1932), to account for salting-out. In this work, an expression based on the above theories, has been proposed, which requires one ternary salting-out parameter 'δ_{123}' and is shown below

\[
\frac{g_{E}^{E(\text{ternary})}}{RT}_{\text{salt-out}} = \frac{D'}{D} - \frac{D}{2} \sum_{k} \frac{v_{k}z_{k}^{2}}{Dk} \frac{1}{2} \tag{C-16}
\]

where,

\[
D' = D\left[1 - \frac{\delta_{123}}{e^{\alpha N_{1}}} \left(X_{2}'X_{3}'\right)^{1/2}N_{1}(X_{2}'B_{13} - X_{3}'B_{12}) e^{\alpha X_{2}'}\right] \tag{C-17}
\]

\[\alpha = 2.0\]

Combining equations (C-16) and (C-17) leads to the form--

\[
\frac{g_{E}^{E(\text{ternary})}}{RT}_{\text{salt-out}} = N_{T} \frac{g_{E}^{E(\text{ternary})}}{RT}_{\text{salt-out}} \tag{C-18}
\]

equation (3-5).
C.4—Development of the Ternary $\ln \gamma_1^+, \ln \gamma_2$ and $\ln \gamma_3$ Expressions

The total excess Gibbs free energy function is obtained by combining equations (3-2) to (3-5). The activity coefficient expressions are obtained by the appropriate differentiation of the total Gibbs free energy function. Since the activity coefficients are a combination of three different terms, the differentiation of each term is performed separately, as below

C.4-I—The Bromley Equation

Combining equations (1-26) (3-2) and (3-3) results in--

\[ \ln \gamma^*_{\text{Bromley}} = \frac{3}{3N_1} \{ \ln(1 + aI) + \frac{1}{1 + aI} - 1 \} + \frac{B}{2} I^2 \]

\[ - 2(1 + \rho I^{1/2}) + \ln(1 + \rho I^{1/2}) \cdot \frac{3}{2} + \frac{(0.06 + 0.06B)}{a^2} \]

\[ \{ \ln(1 + aI) + \frac{1}{1 + aI} - 1 \} + \frac{B}{2} I^2 \]

\[ \frac{3}{3N_1} \{ (0.001vMw + 1) \ln(0.001vMw + 1) \]

\[ - 0.001vMw \} \] (C-18)

Equation (C-18) simplifies to the original $\ln \gamma^*_{\text{Bromley}}$ equation with an additional term, equation (3-8), where

\[ \frac{\partial B}{\partial N_1} = -\frac{3}{2} \alpha B_{123} \frac{1}{N_1^{1/2}} \frac{1}{(1 + aN_1^{1/2})^4} e^{-\alpha X_1^*} \frac{(X_2 X_3)^{1/4}}{4} \] (C-19)
For solvent 2, equation (3-2) with (3-3) is differentiated term by term

I term = \( \frac{3}{N_2} \left[ -2.303 \sqrt[3]{\frac{m}{1000}} A \gamma I^{1/2} \frac{1}{(N_{T_i} M_w)^2} \right] \)

\( - 2(1 + \rho I^{1/2}) + \ln(1 + \rho I^{1/2}) + \frac{3}{2} \}

I term = 2.303 \( \frac{v_m}{1000} \) \( M_{w_2} \) \( A \gamma I^{1/2} \frac{1}{(N_{T_i} M_w)^2} \{ (1 + \rho I^{1/2}) \)

\( - 2 \ln(1 + \rho I^{1/2}) - \frac{1}{(1 + \rho I^{1/2})} - N_{T_i} M_w \)

\( I^{1/2} \frac{2}{(N_{T_i} M_w)^3} \{ \frac{1}{2}(1 + \rho I^{1/2})^2 - 2(1 + \rho I^{1/2}) \)

\( + \ln(1 + \rho I^{1/2}) + \frac{3}{2} \} \frac{\partial A \gamma}{\partial N_2} \]

II term = \( \frac{3}{N_2} \left[ 2.303 \sqrt[3]{\frac{m}{1000}} A \gamma (0.06 + 0.06 B) \right] \)

\[ \{ \ln(1 + aI) + \frac{1}{(1 + aI) - 1} \} \]

II term = + 2.303 \( \frac{v_m}{1000} \) \( -M_{w_2} \) \( (0.06 + 0.6 B) \frac{I}{2} \frac{2}{aI} \)

\[ \{ \frac{(1 + 2aI)}{(1 + aI)^2} - \frac{\ln(1 + aI)}{aI} \} + 0.6 N_{T_i} M_w \frac{I}{2} \frac{2}{aI} \]

\[ \{ \frac{\ln(1 + aI)}{aI} - \frac{1}{(1 + aI)^2} \} \frac{\partial B}{\partial N_2} \]

III term = 3 \( \frac{3}{N_2} \left[ 2.303 \sqrt[3]{\frac{m}{1000}} B I \right] \)

III term = + 2.303 \( \frac{v_m}{1000} \) \( -M_{w_2} \) \( B \frac{I}{2} + N_{T_i} M_w \) \( I \frac{\partial B}{\partial N_2} \)

IV term = \( \frac{3}{N_2} \left[ (0.001 \sqrt{v_m M_w} + 1) \ln(0.001 \sqrt{v_m M_w} + 1) - 0.001 \sqrt{v_m M_w} \right] \)
IV term = \ln(0.001v_mM_w + 1) - 0.001v_mM_w \quad (C-23)

Combining equations (C-20) to (C-23) results in
\ln \gamma_2^{(ternary)} expression, equation (3-12). Similarly an ex-
pression for \ln \gamma_3^{(ternary)} can be obtained.

The different terms of equation (3-12) are defined below

\[
\sigma_2(\rho^{1/2}) = \frac{3}{(\rho^{1/2})^3}[(1 + \rho^{1/2}) - 2 \ln(1 + \rho^{1/2}) - \frac{1}{1 + \rho^{1/2}}]
\]

\[
(C-24)
\]

\[
\sigma_2^{1/2}(\rho^{1/2}) = \frac{2}{(\rho^{1/2})^3}[\frac{1}{2}(1 + \rho^{1/2}) - 2(1 + \rho^{1/2}) + \ln(1 + \rho^{1/2}) + \frac{3}{2}]
\]

\[
(C-25)
\]

\[
\psi_2(aI) = \frac{2}{aI}[\frac{1 + 2aI}{1 + aI} - \frac{\ln(1 + aI)}{aI}]
\]

\[
(C-26)
\]

\[
\psi_2^{1/2}(aI) = \frac{2}{aI}[\frac{\ln(1 + aI)}{aI} - \frac{1}{1 + aI}]
\]

\[
(C-27)
\]

\[
\frac{\partial B}{\partial N_2} = (B_{12} - B_{13}) \frac{X_2'}{N_T} + \frac{B_{123}}{(1 + aN_1^{1/2})^3} \left[ \frac{1}{2} \frac{X_3'}{X_2} \frac{1}{X_2^{1/2}} - (X_2X_3')^{1/4} \right]
\]

\[
(C-28)
\]

\[
\frac{\partial B}{\partial N_3} = (B_{13} - B_{12}) \frac{X_3'}{N_T} + \frac{B_{123}}{(1 + aN_1^{1/2})^3} \left[ \frac{1}{2} \frac{X_2'}{X_3} \frac{1}{X_3^{1/2}} - (X_2X_3')^{1/4} \right]
\]

\[
(C-29)
\]
C.4-II—The NRTL-S equation

Equation (C-15) can be differentiated appropriately to obtain the activity coefficient expression for electrolyte and solvents. Utilizing equations (B-35) to (B-38) with equation (C-15) leads to

\[
\ln \gamma^{*}_{\text{ternary}, \text{NRTL-S}} = \frac{N_2N_3}{RT} \left[ -\frac{Z_{32}}{(N_{A\text{v}_A} + N_2 + N_3G_{32})^2} \right.
\]

\[ \quad - \frac{Z_{23}}{(N_{\text{v}_A} + N_3 + N_2G_{23})^2} + \frac{Z_{32}}{(N_3G_{32} + N_2)^2} \]

\[ \quad + \frac{Z_{23}}{(N_2G_{23} + N_3)^2} \]  \hspace{1cm} (C-30)

Equation (C-30) can be converted in terms of mole fraction, equation (3-8).

\[
\ln \gamma^{(\text{ternary})}_{2, \text{NRTL-S}} = \frac{\partial}{\partial N_2} \left[ \frac{G^E(\text{ternary})}{RT} \right]_{\text{NRTL-S}}
\]

\[
\ln \gamma^{(\text{ternary})}_{2, \text{NRTL-S}} = \frac{1}{RT} \left[ \frac{N_3Z_{32} (N_{A\text{v}_A} + N_2 + N_3G_{32}) - N_2N_3Z_{32}}{(N_{A\text{v}_A} + N_2 + N_3G_{32})^2} \right.
\]

\[ \quad + \frac{N_3Z_{23} (N_{A\text{v}_A} + N_3 + N_2G_{23}) - N_2N_3G_{23}Z_{23}}{N_{A\text{v}_A} + N_3 + N_2G_{23})^2} \]

\[ \quad + \frac{\nu}{\text{v}_A} N_A N_3 \left( \frac{Z_{32}}{(N_3G_{32} + N_2)^2} + \frac{Z_{23}}{(N_2G_{23} + N_3)^2} \right) \]
\[
\ln \gamma_{2, \text{NRTL-S}}(\text{ternary}) = \frac{1}{RT} \left( \frac{\nu}{\nu_A} \right) N_A N_3 \left\{ \frac{Z_{32}}{(N_A \frac{\nu}{\nu_A} + N_2 N_3 G_{32})^2} \right. \\
\left. + \frac{Z_{23}}{(N_A \frac{\nu}{\nu_A} + N_3 + N_2 G_{23})^2} + N_3^2 \frac{G_{32} G_{32}}{(N_A \frac{\nu}{\nu_A} + N_2 + N_3 G_{32})^2} \right. \\
\left. + \frac{Z_{23}}{(N_A \frac{\nu}{\nu_A} + N_3 + N_2 G_{23})^2} \right. \\
\left. \frac{\nu}{\nu_A} N_A N_3 \left\{ \frac{Z_{32}}{(N_3 G_{32} + N_2)^2} + \frac{Z_{23}}{(N_2 G_{23} + N_3)^2} \right. \right.
\right. \\
\right. - 2 \left. \frac{\nu}{\nu_A} N_A N_2 N_3 \left\{ \frac{Z_{32}}{(N_3 G_{32} + N_2)^3} + \frac{Z_{23} G_{23}}{(N_2 G_{23} + N_3)^3} \right. \right\} \right) \right) \right) \right) \right) (C-31)
\]

Using the above approach, an expression for \( \ln \gamma_{3, \text{NRTL-S}}(\text{ternary}) \) can be obtained. These expressions can be rearranged in terms of mole fractions, equation (3-13).

C.3-III—The Salting-Out Term

Equation (3-5) with equation (3-6) is differentiated to obtain the salting-out contribution for the activity coefficients.
\[
\ln \gamma^*_{\text{salt-out}} (\text{ternary}) = \delta_{123} \frac{e^2}{kT} \sum_k \frac{v_k Z_k^2}{B_k} (N_2 N_3)^{1/2} \frac{1}{2} \frac{N_1^2}{e^{\alpha N_1}} \frac{1}{2}
\]
\[
\frac{\partial}{\partial N_1} \left[ \frac{1}{2} \frac{1}{e^{\alpha N_1}} \right] = \frac{1}{\alpha N_1^{1/2}} [N_1 - \frac{1}{2} \alpha] \frac{1}{N_1^{1/2}} \frac{1}{2}
\]

(C-32)

(C-33)

Combining equations (C-32) and (C-33) leads to equation (3-1). And

\[
\ln \gamma_{2, \text{salt-out}} (\text{ternary}) = \delta_{123} \frac{e^2}{kT} \sum_k \frac{v_k Z_k^2}{B_k} \left( N_2 N_3 \right)^{1/2} \frac{1}{2} \frac{1}{e^{\alpha N_1}} \frac{1}{2}
\]

\[
+ \left( N_2 N_3 \right)^{1/2} \frac{\partial}{\partial N_2} \delta' + \delta' \left( N_2 N_3 \right)^{1/2} \frac{\partial}{\partial N_2} \left( \frac{1}{D} \right)
\]

(C-34)

where,

\[
\frac{\partial}{\partial N_2} \left( N_2 N_3 \right)^{1/2} = \frac{1}{2} \left( \frac{N_3}{N_2} \right)
\]

(C-35)

\[
\frac{\partial \delta'}{\partial N_2} = \frac{\partial}{\partial N_2} \left[ e^{\alpha X_2^2 (X_2 B_{13} - X_3 B_{12})} \right]
\]

\[
= e^{\alpha X_2^2 (X_2 B_{13} - X_3 B_{12})} \frac{\partial}{\partial N_2} \left[ e^{\alpha X_2^2 (X_2 B_{13} - X_3 B_{12})} \right] + e^{\alpha X_2^2 (X_2 B_{13} - X_3 B_{12})} \frac{\partial}{\partial N_2} \left( X_2 B_{13} - X_3 B_{12} \right)
\]

\[
\frac{\partial \delta'}{\partial N_2} = \alpha e^{\alpha X_2^2 (X_2 B_{13} - X_3 B_{12})} \left( X_3 B_{12} + (B_{12} + B_{13}) e^{\alpha X_2 X_3 N_T} \right)
\]

\[
\frac{\partial}{\partial N_2} \left[ \frac{1}{D} \right] = - \frac{1}{2} \frac{\partial D}{\partial N_2}
\]

(C-36)

(C-37)

A combination of equations (C-34) to (C-37) results in equation (3-14).

Similarly an expression for \( \ln \gamma_{3, \text{salt-out}} (\text{ternary}) \) can be
obtained with the following additional relationships

\[
\frac{\partial}{\partial N_3} \left[ \left( N_2 N_3 \right)^{1/2} \right] = \frac{1}{2} \frac{N_2}{N_3}^{1/2} \quad (C-38)
\]

\[
\frac{\partial \delta^*}{\partial N_3} = -e^{\frac{\alpha X^*}{N_T}} \left[ \alpha \left( \frac{X^*_2 - X^*_3 B_{12} + X^*_3 B_{13}}{N_T} \right) \right] \quad (C-39)
\]

\[
\frac{\partial}{\partial N_3} \left[ \frac{1}{D} \right] = - \frac{1}{D^2} \frac{\partial D}{\partial N_3} \quad (C-40)
\]
APPENDIX D

DEBYE–HÜCKEL CONSTANTS, DIELECTRIC CONSTANTS AND DENSITIES OF PURE SOLVENTS AND MIXED SOLVENTS; VAPOR PRESSURE CONSTANTS OF PURE SOLVENTS
The Debye-Hückel constant at the system temperature and pressure is given by

\[ A_\gamma = \left( \frac{2\pi Nd}{1000} \right)^{1/2} \left( \frac{\varepsilon}{\varepsilon_0 DT} \right)^{3/2} \]  

(D-1)

or

\[ A_\gamma = 1.8246 \times 10^6 \, d^{1/2} \left( \frac{1}{D^2 T} \right)^{3/2} \]  

(D-2)

where,

- \( d \) - density of the solvent
- \( D \) - dielectric constant of the solvent

The values of the above two properties for the pure solvent or the mixed solvent mixture are presented below.

A. Electrolyte-solvent binary (binary 1-2 or 1-3)

- \( d = d_i \) - pure solvent (2 or 3) liquid density at the system temperature and pressure.

The density data are estimated if experimental data are not available by the following relationship

\[ d = \frac{M_w}{V_T^L} \]  

(D-3)

- \( M_w = M_{w_i} \) - molecular weight of the solvent \( i \)
- \( V_T^L = v_{o_i}^L \) - pure solvent liquid molar volume.

A quadratic equation is used to calculate the pure solvent liquid molar volume, as given in the monograph by Prausnitz et al. (1967)

\[ v_{o_i}^L = a' + b'T + c'T^2 \]  

(D-4)
The constants $a'$, $b'$ and $c'$ are obtained, using experimental liquid volume data at three temperatures, by a method used in the monograph. In Table D-1, liquid volume data are listed for the solvents used in this work.

- $D - D_1$ - pure component dielectric constant at the system temperature and pressure.

B. Electrolyte (1) - solvent (2) - solvent (3) ternary

- $d$ - solvent mixture (electrolyte free) density at the system temperature and pressure.

- $I$ - if the experimental solvent mixture data are available, a six-constant polynomial is fitted to get the concentration dependency of the density. The polynomial expression is then used to calculate the density at different temperatures.

$$d = a_1 + a_2x_3' + a_3x_3'^2 + a_4x_3'^3 + a_5x_3'^4 + a_6x_3'^5$$ (D-5)

In tables D-2 and D-3, liquid density data for $H_2O$-MeOH at $25^\circ C$ and $H_2O$-EtOH at $25^\circ C$, used in this work are tabulated.

- $II$ - if the experimental mixture density data are not available, the solution density is approximated using a linear relationship for the volume equation (D-3). Where,

$$V_T^L = x_2'V_2^oL + x_3'V_3^oL$$ (D-6)

$$M_w = x_2'M_2w_2 + x_3'M_3w_3$$ (D-7)
Figure D.1 Comparison of Experimental and Estimated Densities for the System H₂O-MeOH at 25°C
Figure D.2 Comparison of Experimental and Estimated Densities for the System H$_2$O-EtOH at 25°C
\( v_2^L \) and \( v_3^L \) are calculated, using equation (D-4) for the pure solvents. Alternatively,

\[
d = x_2 v_2^L + x_3 v_3^L
\]  

(D-8)

The validity of equations (D-3) and (D-8) for the mixture is shown in the figures D.1 and D.2 for the H\(_2\)O-MeOH system at 25°C and the H\(_2\)O-EtOH system at 25°C respectively.

\( D = \) Dielectric constant of the liquid solvent mixture (electrolyte-free)

III - if the experimental solvent mixture dielectric constant data are available, a six-constant polynomial is fitted to get the concentration dependency of the dielectric constant. The polynomial expression is then used to calculate the dielectric constant at a different concentrations.

\[
D = A_{1} + A_{2}x_{3}^{2} + A_{3}x_{3}^{3} + A_{4}x_{3}^{4} + A_{5}x_{3}^{5} + A_{6}x_{3}^{6}
\]  

(D-9)

The data given by Akerlöf (1932) for the mixture dielectric constant have been used in this work. The data at a constant composition has been represented as a function of temperature by

\[
\ln D = 2.303 AD_1 + AD_2 \ln(T - 293.15)
\]  

(D-10)

The constants \( AD_1 \) and \( AD_2 \) are listed in the Tables D.4 and D.5 for the mixtures considered in this study.

IV - if the experimental data are not available, the dielectric constant of the solution is approximated using either equation (D-11) or (D-12)
Figure D.3 Comparison of Experimental and Estimated Dielectric Constants of the Mixture H₂O-MeOH at 25°C

Akerlòf (1932)

Eqn. (D-11)

Eqn. (D-12)
Figure D.4 Comparison of Experimental and Estimated Dielectric Constants of the Mixture H$_2$O-EtOH at 25°C
\[ D = D_2 \exp[AX_3'] \quad \text{(D-11)} \]

where,
\[ A = \ln\left[\frac{D_3}{D_2}\right] \quad \text{[Note: } D_3 < D_2\text{]} \]

or
\[ D = D_2X_2' + D_3X_3' \quad \text{(D-12)} \]

The experimental dielectric constant data are compared with the expressions (D-11) and (D-12) in figures (D.3) and (D.4) for the \text{H}_2\text{O-MeOH} system at 25\degree\text{C} and the \text{H}_2\text{O-EtOH} system at 25\degree\text{C} respectively.

Effect of Temperature on \( A_\gamma \)

The Debye-Hückel constant '\( A_\gamma \)' is \( \alpha \frac{1}{T^{3/2}} \) (equation D-2). Also, the dielectric constant and the density of the solvent are functions of temperature. Therefore \( A_\gamma \) is a strong function of temperature.

Effect of Pressure on \( A_\gamma \)

This work is limited only to low pressures. At low pressures, the liquid density and the dielectric constant can be considered to be pressure independent. Therefore \( A_\gamma \) has no effect of pressure.

Vapor-Pressure of the Pure Solvent

A six-parameter equation given in the monograph of Prausnitz et al. (1967) has been used to estimate the vapor pressures of the pure components, equation (1-24). Constants \( C_1, C_2, C_3, C_4, C_5 \) and \( C_6 \) are listed in Table D.6.
<table>
<thead>
<tr>
<th>Solvent</th>
<th>$T(°K)$</th>
<th>$v^L_{°L}$, cc/gmole</th>
</tr>
</thead>
<tbody>
<tr>
<td>EtOH</td>
<td>273.15</td>
<td>57.141</td>
</tr>
<tr>
<td></td>
<td>323.15</td>
<td>60.356</td>
</tr>
<tr>
<td></td>
<td>373.15</td>
<td>64.361</td>
</tr>
<tr>
<td>MeOH</td>
<td>273.15</td>
<td>39.556</td>
</tr>
<tr>
<td></td>
<td>373.15</td>
<td>44.874</td>
</tr>
<tr>
<td></td>
<td>473.15</td>
<td>57.939</td>
</tr>
<tr>
<td>Water</td>
<td>277.15</td>
<td>18.06</td>
</tr>
<tr>
<td></td>
<td>323.15</td>
<td>18.278</td>
</tr>
<tr>
<td></td>
<td>373.15</td>
<td>18.844</td>
</tr>
<tr>
<td>$X_{\text{MeOH}}$</td>
<td>$d$, cc/gmole</td>
<td></td>
</tr>
<tr>
<td>------------------</td>
<td>--------------</td>
<td></td>
</tr>
<tr>
<td>0.0</td>
<td>0.99707</td>
<td></td>
</tr>
<tr>
<td>0.04085</td>
<td>0.98472</td>
<td></td>
</tr>
<tr>
<td>0.06168</td>
<td>0.97919</td>
<td></td>
</tr>
<tr>
<td>0.11445</td>
<td>0.96649</td>
<td></td>
</tr>
<tr>
<td>0.19739</td>
<td>0.94796</td>
<td></td>
</tr>
<tr>
<td>0.24867</td>
<td>0.93658</td>
<td></td>
</tr>
<tr>
<td>0.34382</td>
<td>0.91534</td>
<td></td>
</tr>
<tr>
<td>0.49446</td>
<td>0.88242</td>
<td></td>
</tr>
<tr>
<td>0.61267</td>
<td>0.8579</td>
<td></td>
</tr>
<tr>
<td>0.69241</td>
<td>0.8421</td>
<td></td>
</tr>
<tr>
<td>0.78454</td>
<td>0.82458</td>
<td></td>
</tr>
<tr>
<td>0.89229</td>
<td>0.8051</td>
<td></td>
</tr>
<tr>
<td>1.0</td>
<td>0.78663</td>
<td></td>
</tr>
</tbody>
</table>
TABLE D.3

Liquid Density Data for the H₂O-EtOH System @ 25°C

Ref: Perry and Chilton (1973)

<table>
<thead>
<tr>
<th>X'_{EtOH}</th>
<th>d, cc/gmole</th>
</tr>
</thead>
<tbody>
<tr>
<td>0.0</td>
<td>0.99708</td>
</tr>
<tr>
<td>0.0417</td>
<td>0.98043</td>
</tr>
<tr>
<td>0.0891</td>
<td>0.96639</td>
</tr>
<tr>
<td>0.1436</td>
<td>0.95067</td>
</tr>
<tr>
<td>0.2069</td>
<td>0.93148</td>
</tr>
<tr>
<td>0.2813</td>
<td>0.90985</td>
</tr>
<tr>
<td>0.3699</td>
<td>0.88699</td>
</tr>
<tr>
<td>0.4773</td>
<td>0.86340</td>
</tr>
<tr>
<td>0.6102</td>
<td>0.83911</td>
</tr>
<tr>
<td>0.7788</td>
<td>0.81362</td>
</tr>
<tr>
<td>1.00</td>
<td>0.78506</td>
</tr>
<tr>
<td>$x'_{\text{EtOH}}$</td>
<td>$a_{D1}$</td>
</tr>
<tr>
<td>-----------------</td>
<td>---------</td>
</tr>
<tr>
<td>0.0</td>
<td>1.9051</td>
</tr>
<tr>
<td>0.0417</td>
<td>1.8727</td>
</tr>
<tr>
<td>0.0891</td>
<td>1.8367</td>
</tr>
<tr>
<td>0.1436</td>
<td>1.7968</td>
</tr>
<tr>
<td>0.2069</td>
<td>1.752</td>
</tr>
<tr>
<td>0.2813</td>
<td>1.7024</td>
</tr>
<tr>
<td>0.3699</td>
<td>1.6500</td>
</tr>
<tr>
<td>0.4773</td>
<td>1.5926</td>
</tr>
<tr>
<td>0.6102</td>
<td>1.530</td>
</tr>
<tr>
<td>0.7788</td>
<td>1.4625</td>
</tr>
<tr>
<td>1.00</td>
<td>1.3979</td>
</tr>
</tbody>
</table>

Ref: Akerlöff (1932)
TABLE D.5

Constants for Calculating the Dielectric Constants of Water-MeOH Mixtures at Various Temperatures, Equation (D-10)

Ref: Akerlöf (1932)

<table>
<thead>
<tr>
<th>X'MeOH</th>
<th>AD1</th>
<th>AD2</th>
</tr>
</thead>
<tbody>
<tr>
<td>0.0</td>
<td>0.19051</td>
<td>-0.00205</td>
</tr>
<tr>
<td>0.0588</td>
<td>1.8799</td>
<td>-0.00208</td>
</tr>
<tr>
<td>0.1233</td>
<td>1.8505</td>
<td>-0.00212</td>
</tr>
<tr>
<td>0.1942</td>
<td>1.8190</td>
<td>-0.00218</td>
</tr>
<tr>
<td>0.2727</td>
<td>1.7865</td>
<td>-0.00225</td>
</tr>
<tr>
<td>0.3600</td>
<td>1.7513</td>
<td>-0.00234</td>
</tr>
<tr>
<td>0.4576</td>
<td>1.7120</td>
<td>-0.00244</td>
</tr>
<tr>
<td>0.5676</td>
<td>1.6658</td>
<td>-0.00252</td>
</tr>
<tr>
<td>0.6923</td>
<td>1.616</td>
<td>-0.00248</td>
</tr>
<tr>
<td>0.8351</td>
<td>1.5648</td>
<td>-0.00242</td>
</tr>
<tr>
<td>1.00</td>
<td>1.5099</td>
<td>-0.00234</td>
</tr>
</tbody>
</table>
### TABLE D.6

Constants for the Vapor Pressure, Equation (D-12)

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$C_1$</th>
<th>$C_2$</th>
<th>$C_3$</th>
<th>$C_4$</th>
<th>$C_5$</th>
<th>$C_6$</th>
<th>Ref</th>
</tr>
</thead>
<tbody>
<tr>
<td>EtOH</td>
<td>123.9120350</td>
<td>-8754.0896</td>
<td>0.0</td>
<td>0.020198435</td>
<td>0.0</td>
<td>-18.1</td>
<td>Prausnitz et al. (1966)</td>
</tr>
<tr>
<td>MeOH</td>
<td>12.3858228</td>
<td>-3880.50203</td>
<td>0.0</td>
<td>-24.355</td>
<td>0.0</td>
<td>0.0</td>
<td>Hala (1969)</td>
</tr>
<tr>
<td>Water</td>
<td>70.4346943</td>
<td>-7362.6981</td>
<td>0.0</td>
<td>0.006952085</td>
<td>-</td>
<td>-9.0</td>
<td>Prausnitz et al. (1966)</td>
</tr>
</tbody>
</table>
APPENDIX E

CALCULATION OF FUGACITY COEFFICIENTS AND POYNTING EFFECT
The following two data points are used for the calculation of $\phi^O_i$, $\phi^V_i$, P.E. and $F_i$—

[Ref: Hala (1968)]

<table>
<thead>
<tr>
<th>Point #</th>
<th>System</th>
<th>T(°C)</th>
<th>$X_3$</th>
<th>$Y_3$</th>
<th>$P$(mmHg)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>$H_2O(2)$-$MeOH(3)$</td>
<td>60</td>
<td>0.0343</td>
<td>0.2106</td>
<td>183.64</td>
</tr>
<tr>
<td>2</td>
<td>$H_2O(2)$-$MeOH(3)$</td>
<td>60</td>
<td>0.7582</td>
<td>0.901</td>
<td>538.64</td>
</tr>
</tbody>
</table>

Pure component second virial coefficients for $H_2O$ and MeOH and the cross-virial coefficient for $H_2O$-MeOH mixture are estimated by Pitzer's correlations, [Smith and VanNess (1973)]

\[
B^O = 0.083 - \frac{0.422}{T_r^{1.6}} \quad (E-1)
\]

\[
B^1 = 0.139 - \frac{0.172}{T_r^{4.2}} \quad (E-2)
\]

\[
B_{ij} = (B^O + \omega B^1) \frac{RT_C}{P_C} \quad (E-3)
\]

Pure component liquid molar volumes and pure component vapor pressures are calculated as shown in Appendix D.

The mixture properties necessary to estimate the cross-virial coefficient of the mixtures are calculated by

\[
T_{ci} = \sum_{i=2}^{3} Y_i T_{ci} \quad (E-4)
\]

\[
P_{ci} = \sum_{i=2}^{3} Y_i P_{ci} \quad (E-5)
\]
\[ \omega_{ij} = \sum_{l=2}^{3} y_i \omega_i \]  \hspace{1cm} (E-6)

\[ \delta_{23} = 2B_{23} - B_{22} - B_{33} \]  \hspace{1cm} (E-7)

Applying equations (E-4) to (E-7) with equations (E-1) to (E-3) the cross-properties for the two data points are listed in Table E.2.

The pure component properties listed in Table E.1 and the mixture properties presented in Table E.2 are used in equations (1-5), (1-6), (1-8) and (1-10) to calculate \( \phi_i^0 \), \( \text{(P.E.)}_i \), \( \hat{\phi}_i^v \) and \( F_i \) (see Tables E.3 and E.4 for the two data points).

The values of \( F_i \) in Tables E.3 and E.4 for \( \text{H}_2\text{O} \) and \( \text{MeOH} \) justifies the assumption that, at low pressures, \( F_i = 1.0 \).
\[
\begin{array}{cccccccc}
\text{Component} & \nu_i^\text{OL} & p_i^\text{L} & B_i & b_i & \omega_i & B_{ii} \\
& (\text{cc/gmole}) & (\text{mmHg}) & & & & (\text{cc/gmole}) \\
\text{H}_2\text{O} & 18.364 & 149.383 & -1.13793 & -2.6577 & 0.348 & -503.4 \\
\text{MeOH} & 41.8187 & 634.315 & -0.75793 & -0.91993 & 0.556 & -665.9 \\
\end{array}
\]

\[
\begin{array}{cccc}
\text{Point #1} & \omega_{23} & B_{23} & \delta_{23} \\
& & (\text{cc/gmole}) & (\text{cc/gmole}) \\
1 & 0.3912 & -513.0 & 143.3 \\
2 & 0.5354 & -622.3 & -75.3 \\
\end{array}
\]

\[
\begin{array}{cccc}
\phi_i^\text{O}, \text{ P.E., } \hat{\phi}_i^\text{V} \text{ and } F_i \text{ for Data Point #1} \\
\text{Component} & \phi_i^\text{O} & (\text{P.E.})_i & \hat{\phi}_i^\text{V} & F_i \\
\text{H}_2\text{O} & 0.99556 & 1.00003 & 0.99562 & 1.0009 \\
\text{MeOH} & 0.99413 & 0.99909 & 0.99492 & 1.0075 \\
\end{array}
\]

\[
\begin{array}{cccc}
\phi_i^\text{O}, \text{ P.E., } \hat{\phi}_i^\text{V} \text{ and } F_i \text{ for Data Point #2} \\
\text{Component} & \phi_i^\text{O} & (\text{P.E.})_i & \hat{\phi}_i^\text{V} & F_i \\
\text{H}_2\text{O} & 0.98703 & 1.00034 & 0.98548 & 0.9981 \\
\text{MeOH} & 0.98288 & 0.9998 & 0.98288 & 1.0002 \\
\end{array}
\]
APPENDIX F

COMPUTER PROGRAMS
CORRELATION AND PREDICTION OF ACTIVITY COEFFICIENTS: TOTAL PRESSURE AND VAPOR PHASE COMPOSITIONS

This appendix contains the following programs:

F.1 Main Program

Calls subroutines INPDAT, LSQ2, FIBN, TITLE. The main program reads the different indicating markers to perform correlation or prediction of binary or ternary VLE. The comment cards included in this section explain the different options used in this program.

F.2 INPDAT

Subroutine reads the input data.

F.3 FITIT [calls POLIFI]

Subroutine fits a polynomial of degree 5.

F.4 POLIFI

Subroutine makes a least-square fit for FITIT [calls subroutine DETERM].

F.5 DETERM

Subroutine performs the error analysis for POLIFI.

F.6 VAPPRE

Subroutine calculates the pure component vapor pressures at the system temperature [Equation (1-24)].
F.7 TEMPD

Subroutine calculates pure component liquid molar volumes at the system temperature. This also calculates the dielectric-constant of the mixtures. This calls subroutine FITIT to obtain the polynomial constants for the concentration dependence of the dielectric constants at the system temperature.

The following equations are used for the temperature dependence.

For the liquid molar volume

\[ V = A + BT + CT**2 \]

Where A, B, C - constants. T - temperature of the system.

For the dielectric constant

\[ DS = \exp[2.303(ADT1 + ADT2(T - 293.15))] \]

Where ADT1, ADT2 - Constants for the temperature dependency of a solvent mixture. T - temperature of the system. DS - dielectric constant of the mixture (see Appendix-D).

F.8 LSQ2

Subroutine applies a search technique to find the best unknown variables that will result in the minimum value of an objective function [number of variables should be greater or equal to 2]. Calls subroutine FN.
F.9 FN

Subroutine calls different subroutines to calculate the different contributions for the activity coefficients in two models. Finally it calls subroutine MINFUN to set up the objective function. Calls subroutines NRTL1, BROML, ADITON, NRTL2, DEBHUC, VAPPRE, MINFUN.

F.10 FIBN

Subroutine applies the Fibonacci method to find the one unknown variable that will minimize the non-linear objective function.

F.11 FUNCT

Subroutine calls different subroutines as in subroutine FN.

F.12 NRTL1

Subroutine calculates solvent-solvent interaction contributions of the activity coefficients in a ternary mixture using the simplified and the modified form of the NRTL equation in Model #2 [see Chapter 3; Equation (2-4)].

F.13 BROML

Subroutine calculates ion-ion and ion-solvent interaction contributions of the activity coefficients in a binary or ternary mixture using the Bromley equation in Model #2 [see Chapter 3; Equations (3-8), (3-12) and (C-9)].
F.14 ADITON

Subroutine calculates salting-out contribution for the activity coefficients in a ternary mixture in Model #2 [see Chapter 3; Equations (3-10) and (3-14)].

F.15 FUNCB

Subroutine for mixing rule for the Bromley parameter in a ternary mixture [see Chapter 3; Equation (3-3)].

F.16 FUNCT

Subroutine calculates the Debye-Huckel constant of mixtures and derivatives of the Debye-Huckel constant with respect to the number of moles of solvents [see Appendix-D].

F.17 NRTL2

Subroutine calculates ion-solvent and solvent-solvent interaction contributions of the activity coefficient in a binary or ternary mixture using the modified NRTL equation in Model #1 [see Chapter 2; Equations (2-7), (2-14) and (2-17)].

F.18 DEBHUC

Subroutine calculates ion-ion interaction contributions for the activity coefficients in a binary or ternary mixture in Model #1 [see Chapter 2; Equations (2-13) and (2-16)].

F.19 MINFUN

Subroutine sets up the objective function.
F.20 TITLE

Subroutine makes tables for output results.

The comment cards included in the programs are assumed to make this program self-explanatory.
MAIN PROGRAM -

This program is to correlate or predict either binary or ternary VLE data or binary & ternary VLE data together. The program can be used either for isobaric or isothermal data.

A binary in this work is defined as a mixture of either two solvents or one electrolyte and one solvent. A ternary mixture is defined as a mixture of one electrolyte and two solvents.

This program is written by Anil K. Rastogi at the New Jersey Institute of Technology as a part of doctoral dissertation, year 1981.

NSET - # of system data to be used
LL - total # of trial to be used in LSQ2
EE - tolerance error for the LSQ2

NAME(J) - information about the system; J = 10; 20 cards
XLIM - (maximum molality - 1) range of the electrolyte
XLLIM - minimum molality range of the electrolyte
NAME1, ETC - name of the system to be used

INDF - # of increments to be given for the maximum molality range

KP - = 1 EQUATIONS USED ARE BROOMEY; SIMPLIFIED NRTL1;
    AND SALTINGOUT.(MODEL # 2 IN THE THESIS)
    = 2 EQUATIONS USED ARE EXTENDED DEBYE-HUCKEL &
    NRTL2.(MODEL # 1 IN THE THESIS)

NBIN - (when KP = 1)
    = 1 PARAMETER B012 REGRESSED
PARAMETER DO13 REGRESSED
PARAMETER B123 OR DELTA REGRESSED (WHEN M=1)
PARAMETERS B123 & DELTA REGRESSED (WHEN M=2)
(P=[2]

GP2 & ZP2 REGRESSED (IF NNRTL=2)
DG2 & DGB2 REGRESSED (IF NNRTL=1)
GP3 & ZP3 REGRESSED (IF NNRTL=2)
DG3 & DGB3 REGRESSED (IF NNRTL=1)
DG2 & DGB2 REGRESSED
GP2, ZP2, GP3 & ZP3 REGRESSED (IF NNRTL=2)
DG2, DGB2, DG3 & DGB3 REGRESSED (IF NNRTL=1)

NPION - TYPE OF POSITIVE ION FOR THE CRYSTALLOGRAPHIC RADI
= 1 H ION
= 2 Li ION
= 3 Na ION
= 4 Ca ION

NNION - TYPE OF NEGATIVE ION FOR THE CRYSTALLOGRAPHIC RADI
= 1 Br ION
= 2 Cl ION

NREG - = 1 VLE DATA ARE PREDICTED
       = 2 VLE DATA ARE CORRELATED

NDEN - = 1 EXPERIMENTAL DATA ARE USED FOR THE DENSITY. A
       SIX PARAMETER POLYNOMIAL IS USED FOR THE CONCENTRATION
       DEPENDENCY OF THE DENSITY.
       = 2 AN APPROXIMATE RELATIONSHIP FOR THE DENSITY IS
       USED ( = MW/V)

NBROM - = 1 DELTA IS FIXED AND B123 IS REGRESSED (M=1)
        = 2 B123 IS FIXED AND DELTA IS REGRESSED (M=1)

NTYPE - = 1 DATA OF BINARY 1-2 ARE REGRESSED
        = 2 DATA OF BINARY 1-3 ARE REGRESSED
        = 3 DATA OF BINARY 2-3 ARE REGRESSED
        = 4 DATA OF TERNARY 1-2-3 ARE REGRESSED
        = 5 DATA OF TERNARY 1-2-3 & BINARY 1-3 REGRESSED
        = 6 DATA OF TERNARY 1-2-3 & BINARY 1-2 REGRESSED
        = 7 DATA OF TERNARY 1-2-3 & BINARY 2-3 REGRESSED
        = 8 DATA OF TERNARY 1-2-3 & BINARY 1-2 & BINARY 1-3
           REGRESSED
        = 9 DATA OF TERNARY 1-2-3 & BINARY 1-2 & BINARY 1-3
           & BINARY 2-3 REGRESSED

NMIN - INTEGER FOR THE DIFFERENT OBJECTIVE FUNCTION
= 1 SUM OF ((GCAL-GEXP)/GEXP)**2. IS MINIMIZED
= 2 SUM OF (DP/PEXP)**2. + (DY*10.)**2. IS MINIMIZED
= 3 SUM OF ((GCAL-GEXP)/GEXP)**2. + (DY*10.)**2. IS
   MINIMIZED
= 4 SUM OF ((DPCAL-DPEXP)/DPEXP)**2. IS MINIMIZED FOR
   THE BINARY

ALFA - ALFA FOR THE NON-ELECTROLYTE BINARY IN THE NRTL
EQUATION

DG23, DG32 - NON-ELECTROLYTE BINARY PARAMETERS IN THE NRTL EQUATION

ALFA2, ALFB2 - IONIC ALFAS FOR 1-2 & 1-3 BINARY (EQ. 2-6) BINARIES ALFA3, ALFB3

DGA2, DGB2 - TEMPERATURE INDEPENDENT PARAMETERS IN THE MODIFIED NRTL EQUATION FOR THE IONS (EQ. 2-6) MODEL #1

DGA3, DGB3 - NRTL EQUATION FOR THE IONS (EQ. 2-6) MODEL #1

GPN2, ZPN2 - IONIC PARAMETERS IN THE MODIFIED NRTL EQUATION

GPN3, ZPN3 - MODEL #1 (EQ. 2-5)

B012, B112 - BROMLEY BINARY TEMPERATURE INDEPENDENT PARAMETERS

B013, B113 - BROMLEY BINARY TEMPERATURE INDEPENDENT PARAMETERS (CHAPTER 3)

B123 - TERNARY ADJUSTABLE PARAMETER FOR THE MIXING RULE IN THE BROMLEY EQUATION

DELTA - TERNARY SALTING OUT PARAMETER FOR THE MODEL #2

ALPHA1 - TOLERANCE LIMIT FOR THE FIBBNAUCHI SUBROUTINE

AXT1, BXT1 - PARAMETER LIMITS FOR THE FIBBNAUCHI SUBROUTINE

M - INTEGER TO SPECIFY $H$ OF PARAMETER TO BE REGRESSED

= 1 ONE PARAMETER (B012 OR B013 OR B123 OR DELTA)

= 2 MORE THAN 1 PARAMETER

XTX - INITIAL VALUES OF THE PARAMETERS FOR THE LSQ2

DXX - INCREMENT FOR THE PARAMETERS IN THE LSQ2

****************************************************************************************************************************

COMMON XOMOL(99), XX(3,99), XXF(3,99), G6(3,99), PP(99), YY(3,99),

GNT(3,99), GEL(3,99), GCAL(3,99), AMW(3), ERROR(3,99)

G6H(3,99), GPHY(3,99), TTT(99), BM(99), GADD(3,99)

YCAL(3,99), ADD(6), PTC(99)

COMMON NP, INDF, FNP, FMN, FZP, FZN, FK, ALFA, DG23, DG32, GPN2,

GPN3, ZPN2, ZPN3, KF, NBIN, NPION, NNION, DELTA

COMMON NDEN, NNRTL, ALFA2, ALFB2, ALFA3, ALFB3, DGA2, DGB2,

DG23, DG32, Z23, Z32 - CAL/sMOLE-K

****************************************************************************************************************************

% DGA3, DGB3, NBROM, NREG, NTYPE, NMIN

COMMON B012, B112, B013, B113, CP2(6), CP3(6), CV2(3), CV3(3),

B123, ADT(2,20), NXD, XD(20)
DIMENSION YV(3,99),XL(3,99),P(99),XMOL(99),XF(3,99),
G(3,99),T(99),PSH(3),BY(99),DP(99)
REAL *8 NAME(10)
REAL *8 NAME1,NAME2,NAME3

DIMENSION XT(6),DX(6),Y(7),X(7,10),XTX(6),DXX(6)
READ(5,9001)NSET,LL,EE
9001 FORMAT(2I3,F10.1)
WRITE(6,9002)NSET,LL,EE
9002 FORMAT(’1’,10X,’% OF DATA SET TO BE USED =’,I3,’TRIAL’)
DO 9110 JAN=1,NSET
DO 4100 I=1,20
READ(5,2100)(NAME(J),J=1,10)
WRITE(6,2100)(NAME(J),J=1,10)
4100 CONTINUE
2100 FORMAT(10A8)
READ(5,2200)XLIM,XLLIM
WRITE(6,2200)XLIM,XLLIM
2200 FORMAT(2F10.5)
READ(5,6080)NAME1,NAME2,NAME3
FORMAT(3A8)
READ(5,2300)INDF,KP,NBIN,NNION,NDEN,NNRTL,NTYPE,NMIN
WRITE(6,2300)INDF,KP,NBIN,NNION,NDEN,NNRTL,NTYPE,NMIN
2300 FORMAT(1112)
CALL INPDAT(XL,XF,XMOL,T,AMW,FK,FNP,FNM,FZP,FZN,
G,NPT,YV,P,NCOMP,ADD,CV2,CV3,CP2,CP3,ADT,NXD,NID)
FORMAT(BF10.5)
READ(5,2400)ALFA,DG23,DG32,ALFA2,ALFB2,ALFA3,ALFB3
WRITE(6,2400)ALFA,DG23,DG32,ALFA2,ALFB2,ALFA3,ALFB3
READ(5,2400)GPN2,ZPN2,GPN3,ZPN3,DA2,DGB2,DGA3,DGB3
WRITE(6,2400)GPN2,ZPN2,GPN3,ZPN3,DA2,DGB2,DGA3,DGB3
READ(5,2400)B012,B112,B013,B113,B123,DELTA
WRITE(6,2400)B012,B112,B013,B113,B123,DELTA
READ(5,9141)ALPHA1,XT1,BXT1
WRITE(6,9141)ALPHA1,XT1,BXT1
9141 FORMAT(F10.7,2F10.4)
READ(5,2300)M,MM
WRITE(6,2300)M,MM
DO 4400 I=1,MM
READ(5,2600)XTX(I),DXX(I)
WRITE(6,2600)XTX(I),DXX(I)
4400 CONTINUE
2600 FORMAT(2F10.4)
C
C INITIALIZATION FOR THE LSQ2
DO 4300 KJ=1,INDF
DO 4200 JJ=1,7
DO 4200 JI=1,10
X(JJ,JI)=0.0
4200 CONTINUE

DO 4401 IK=1,MM
XT(IK)=XTX(IK)
DX(IK)=DXX(IK)
4401 CONTINUE
ALPHA=ALPHA1
AXT=AXT1
BXT=BXT1

THIS PART OF THE MAIN PROGRAM SPECIFIES THE TYPE OF
DATA TO BE USED. ALSO DATA ARE CONVERTED INTO DIFFERENT
SYMBOLS FOR MINIMUM AND MAXIMUM MOLALITY LIMITS.

XLIM=XLIM+1.0
J=0
WRITE(6,2410)XLIM
2410 FORMAT(1,'5X','**************MAXIMUM MOLALITY USED=',F10.5)
DO 4310 I=1,NPT
GO TO (9901,9902,9003,9004,9005,9006,9007,9008,9000),NTYPE
9901 IF(XL(3,I).EQ.0.0)GO TO 9000
       GO TO 4310
9902 IF(XL(2,I).EQ.0.0)GO TO 9000
       GO TO 4310
9003 IF(XHOL(I).EQ.0.0)GO TO 9000
       GO TO 4310
9004 IF(XHOL(I).EQ.0.0)GO TO 4310
       IF(XL(2,I).EQ.0.0)GO TO 4310
       IF(XL(3,I).EQ.0.0)GO TO 4310
       GO TO 9000
9005 IF(XHOL(I).EQ.0.0)GO TO 4310
       IF(XL(3,I).EQ.0.0)GO TO 4310
       GO TO 9000
9006 IF(XHOL(I).EQ.0.0)GO TO 4310
       IF(XL(2,I).EQ.0.0)GO TO 4310
       GO TO 9000
9007 IF(XL(2,I).EQ.0.0)GO TO 4310
       IF(XL(3,I).EQ.0.0)GO TO 4310
       GO TO 9000
IF(XMOL(I).EQ.0.0)GO TO 4310
IF(XMOL(I).LT.XLLIM)GO TO 4310
IF(XMOL(I).LE.XLIM)GO TO 3100

J=J+1
NP=J
YY(2,J)=YV(2,I)
YY(3,J)=YV(3,I)
PP(J)=P(I)
TTT(J)=T(I)
XXMOL(J)=XMOL(I)
XX(1,J)=XL(1,I)
XX(2,J)=XL(2,I)
XX(3,J)=XL(3,I)
XXF(2,J)=XF(2,I)
XXF(3,J)=XF(3,I)
GG(1,J)=G(1,I)
GG(2,J)=G(2,I)
GG(3,J)=G(3,I)

CONTINUE
IF(KP.LE.1)MO=2
IF(KP.GT.1)MO=1

PROGRAM FOR THE PREDICTION

IF(NREG.GT.1)GO TO 110
WRITE(6,160)MO
160 FORMAT(///,5X,'VLE DATA ARE PREDICTED USING MODEL #',I3)
   LIC=1
   CALL FN(YYD,XT,LIC)
   GO TO 3300

PROGRAM FOR THE CORRELATION

WRITE(6,170)KP
170 FORMAT(///,5X,'VLE DATA ARE CORRELATED USING MODEL #',I3)
   IF(M.GT.1)GO TO 3200
   CALL FIBN(ALPHA,AXT, EXT)
   GO TO(240,250,260),NBIN
   WRITE(6,270)B012
240  FORMAT(///,10X,'B012 = ',G12.5)
270  GO TO 3300
250  WRITE(6,280)B013
280  FORMAT(///,10X,'B013 = ',G12.5)
260  GO TO(210,220),NBROM
FROM HERE MAIN PROGRAM ARRANGES THE OUTPUT

GCAL(J,I)  - CALCULATED ACTIVITY COEFFICIENT
GG(J,I)  - EXPERIMENTAL ACTIVITY COEFFICIENT
ERROR(J,I)  - % ERROR IN ACTIVITY COEFFICIENT

ALL THE SEPARATE PARTS LISTED BELOW ARE IN THE LN FORM.

GDH(J,I)  - DEBYE-HUCKEL PART OF THE ACTIVITY COEFFICIENT
GPHY(J,I)  - EXTENDED D.H. PART OF THE COULOMBIC TERMS
BESIDES D.H. PART IN THE ACTIVITY COEFFICIENT
GADD(J,I)  - SALTING OUT CONTRIBUTION TO THE ACTIVITY COEFFICIENT
GNRT(J,I)  - NRTL CONTRIBUTION TO THE ACTIVITY COEFFICIENT

YSUM = 0.0
PSUM = 0.0
NV = 0
NVF = 0
DO 4530 I = 1, NP
IF(XXMOL(I) .EQ. 0.0) GO TO 3400
IF(XX(3,I) .EQ. 0.0) GO TO 3401
IF(XX(2,I) .EQ. 0.0) GO TO 3402
IF(GG(1,I) .EQ. 1.0) GO TO 3400
IF(GG(2,I) .EQ. 1.0) GO TO 3405
GO TO 3416
3405
GG(2,I) = 1.0 E 58
GG(3,I) = 1.0 E 58
ERROR(2,I) = 1.0 E 58
ERROR(3,I) = 1.0 E 58
GO TO 3415
3401 GG(3,I)=1.0 E 58
ERROR(3,I)=1.0 E 58
IF(GG(2,I).EQ.1.0)GO TO 3403
IF(GG(1,I).EQ.1.0)GO TO 3400
GO TO 3415
3403 GG(2,I)=1.0 E 58
ERROR(2,I)=1.0 E 58
GO TO 3415
3402 GG(2,I)=1.0 E 58
ERROR(2,I)=1.0 E 58
IF(GG(3,I).EQ.1.0)GO TO 3404
IF(GG(1,I).EQ.1.0)GO TO 3400
GO TO 3415
3404 GG(3,I)=1.0 E 58
ERROR(3,I)=1.0 E 58
GO TO 3415
3400 GG(1,I)=1.0 E 58
ERROR(1,I)=1.0 E 58
IF(XX(2,I).EQ.0.0)GO TO 3415
IF(XX(3,I).EQ.0.0)GO TO 3415
3416 DP(I)=PTC(I)-PP(I)
PSUM=PSUM+ABS(DP(I))
NVP=NVP+1
3410 NV=NV+1
DY(I)=YCAL(3,I)-YY(3,I)
YSUM=YSUM+ABS(DY(I))
WRITE(6,2110)
2110 FORMAT(/,'COMPONENT #',2X,'MOLE-FRACTION',8X,'YEXP ',8X,'YCAL','10X','YCAL-YEXP',7X,'PCAL-PEXP')
2115 FORMAT(8X,I2,5(5X,F12.5))
DO 4520 J=2,NCOMP
WRITE(6,2115)J,XX(J,I),YY(J,I),YCAL(J,I),DY(I),DP(I)
4520 CONTINUE
3415 WRITE(6,2120)
2120 FORMAT(1X,'COMPONENT #',3X,'MOLALITY ',2X,'LOG NRTL ',
2X,'LOG COULOMBIC ',1X,'LOG PHYSICAL ',1X,'LOG ADDITION ',
3X,'GAM CAL ',2X,'GAM EXP ',2X,'% ERROR IN GAMMA ')
DO 4530 J=1,NCOMP
WRITE(6,2125)J,XXMOL(I),GNRT(J,I),GDH(J,I),GPHY(J,I),GADD(J,I),
G3CAL(J,I),GG(J,I),ERROR(J,I)
4530 CONTINUE
2125 FORMAT(8X,I2,8X,G10.3,3X,G10.3,5X,G10.3,4X,G10.3,5X,G10.3,4X,
3X,G10.3,3X,G10.3,4X,G10.3)

TABLE FOR THE MEAN MOHAL ACTIVITY COEFFICIENT

CALL TITLE(NAME1,NAME2,NAME3,ALFA,DG23,DG32,GPN2,GPN3,ZPN2,
% ZPN3,B012,B013,B112,B113,B123,DELTA,KP,NTYPE,XX(2,1),
% XXMOL(1),NNRTL,DGA2,DGA3,DGB2,DGB3,ALFA2,ALFB2,ALFA3,
% ALFB3
WRITE(6,7165)
7165 FORMAT(/' MOLALITY', 'X2', '10X', 'X3', '6X', 'MEANMOLALGAM', '1X',
% 'MEANMOLALGAM', '2X', 'ERROR')
WRITE(6,7167)
7167 FORMAT(37X,'EXPERIMENTAL', 'CALCULATED')

GSUM=0.0
DO B110 I=1,NP
  IF(ERROR(I,I),GE.1.0E58)GO TO B110
  NS=NS+1
  GSUM=GSUM+ABS(ERROR(I,I))
WRITE(6,7170)XXMOL(I),XX(2,I),XX(3,I),GG(I,I),GCAL(I,I),
% ERROR(I,I)
B110 CONTINUE

IF(NS*.LE.0)NS=1
GMEAN=GSUM/NS
WRITE(6,7180)GMEAN
7180 FORMAT(/' AVG ERROR IN MEAN MOLAL ACTIVITY COEFFICIENTS'
% ',G12.5)
WRITE(6,7181)NS
7181 FORMAT(7X,' BASED ON #OF POINTS FOR GAMA SALT=',I3)

C C C

CALL TITLE(NAME1,NAMN2,NAMN3,ALFA,DGA2,DGA3,DGB2,DGB3,ALFA2,ALFB2,ALFA3,
% ZPN3,B012,B013,B112,B113,B123,DELTA,KP,NTYPE,XX(2,1),
% XXMOL(1),NNRTL,DGA2,DGA3,DGB2,DGB3,ALFA2,ALFB2,ALFA3,
% ALFB3)
WRITE(6,7190)
7190 FORMAT(/' MOLALITY', 'X2', '9X', 'DY', '9X', 'DP')
DO B220 I =1,NP
  IF(XX(2,I),EQ.0.0) GO TO B220
  IF(XX(3,I),EQ.0.0) GO TO B220
WRITE(6,7195)XXMOL(I),XX(2,I),XX(3,I),YY(3,I),YCAL(3,I),
% DY(I),DP(I)
B220 CONTINUE

IF(NVLE.0)NV=1
IF(NVPLE.0)NVP=1
DYAVG=YSUM/NV
DPAVG=PSUM/NVP
WRITE(6,7200) DYAVG
7200 FORMAT(/' AVERAGE DY(YCAL-YEXP)=',G12.5)
WRITE(6,7201)NV
7201 FORMAT(9X,'BASED ON # OF POINTS FOR Y = ',I3)
WRITE(6,7210)DPAVG
7210 FORMAT(///,10X,'AVERAGE DP(PCAL-PEXP)=',&9Xf'G12.5')
WRITE(6,7221)NVP
7221 FORMAT(9X,'BASED ON # OF POINTS FOR DP = ',I3)

TABLE FOR THE BINARY DELTA P

CALL TITLE(NAME1,NAME2,NAME3,ALFA,DG23,DG32,GPN2,GPN3,ZPN2, 
ZPN3,B012,B013,B112,B113,B123,DELTA,KP,NTYPE,XX(2,1), 
XXMOL(1),NNRTL,DGA2,DGA3,DGB2,DGB3,ALFA2,ALFB2,ALFA3, 
ALFB3)
NDP=0
XDPSUM=0.0
WRITE(6,8290)
8290 FORMAT(///,2X,'MOLALITY',3X,'XSOLVENT',6X,'DPEXP',6X,'DPCAL', 
% 2X,'% ERROR IN DP')
DO 8230 I=1,NP
CALL VAPPRE(CP2,CP3,PSM,TTT(I))
IF(XX(2,I).EQ.0.0)GO TO 8235
IF(XX(3,I).EQ.0.0)GO TO 8240
GO TO 8245
8235 IF(GG(3,I).EQ.1.0)GO TO 8245
DPEXP=PSM(3)-XX(3,I)*GG(3,I)*PSM(3)
DPCAL=PSM(3)-XX(3,I)*GICAL(3,I)*PSM(3)
PDP=(DPCAL-DPEXP)/DPEXP*100.0
XD=XX(3,I)
GO TO 8250
8240 IF(GG(2,I).EQ.1.0)GO TO 8245
DPEXP=PSM(2)-XX(2,I)*GG(2,I)*PSM(2)
DPCAL=PSM(2)-XX(2,I)*GICAL(2,I)*PSM(2)
PDP=(DPCAL-DPEXP)/DPEXP*100.0
XD=XX(2,I)
8250 WRITE(6,8255)XXMOL(I),XD,DPEXP,DPCAL,PDP
8255 FORMAT(2X,G10.3,G12.5,G10.3,G12.5,G12.5)
NDP=NDP+1
XDPSUM=XDPSUM+ABS(PDP)
GO TO 8230
8245 NDP=1
823Q CONTINUE
PDPAVG=XDPSUM/NDP
WRITE(6,8265)PDPAVG
8265 FORMAT(///,6X,'AVG % ERROR IN DP = ',G12.5)
WRITE(6,8270)NDP
8270 FORMAT(7X,'BASED ON # OF POINTS = ',I3)
IF(KP.GT.1)GO TO 4300
WRITE(6,9220)
C 9220 FORMAT('1',4X,'MOLALITY',5X,'X3',10X,'BMIXTURE')
WRITE(6,9230)(XXMOL(I),XXF(3,I),BM(I),I=1,NP)
9230 FORMAT('1',4X,610.3,2612.5)
4300 CONTINUE
WRITE(6,7202)
7202 FORMAT('1')
9110 CONTINUE
STOP
END
SUBROUTINE INPDAT

THIS SUBROUTINE READS ALL INPUT VLE DATA AND PURE COMPONENT DATA

******************************************************************************

NCOMP - # OF COMPONENTS (= 3 FOR THIS PROGRAM)
NP - # OF DATA POINTS IN A SYSTEM
NXD - # OF SOLUTION DIELECTRIC CONSTANT DATA POINTS
NXDD - # OF SOLUTION DENSITY DATA POINTS
CP2,CP3 - PURE COMPONENT VAPOR PRESSURE CONSTANTS FOR THE COMPONENTS 2 & 3 RESPECTIVELY.
AMW(I) - MOLECULAR WEIGHT OF THE COMPONENT I
V21,V22,V23 - PURE COMPONENT LIQUID VOLUMES AT DIFFERENT TEMPERATURES
V31,V32,V33 - TEMPERATURE DEPENDENT LIQUID VOLUME CONSTANTS
CV2(I), CV3(I) - TEMPERATURE DEPENDENT LIQUID VOLUME CONSTANTS FOR THE COMPONENT 2 & 3 RESPECTIVELY.
ADT(1,I) - DIELECTRIC CONSTANTS FOR A TEMPERATURE OF THE SOLUTION.
ADT(2,I) - DIELECTRIC CONSTANTS FOR A TEMPERATURE OF THE SOLUTION.
XD(I) - MOLE-FRACTION OF THE SOLVENT(3) IN THE SOLUTION FOR THE DIELECTRIC CONSTANTS(ELECTROLYTE FREE)
FK - TOTAL # OF IONS OF THE ELECTROLYTE
FNP - # OF POSITIVE IONS IN THE ELECTROLYTE
FNM - # OF NEGATIVE IONS IN THE ELECTROLYTE
FZP - POSITIVE ION VALENCY(Absolute Units)
FZN - NEGATIVE ION VALENCY(Absolute Units)
DENS(I) - DENSITY OF THE SOLUTION(ELECTROLYTE FREE)
XDD(I) - MOLE-FRACTION OF THE SOLVENT(3) IN A SOLUTION FOR THE DENSITY DENS(I).

XMOL(I) - MOLALITY OF THE I TH POINT
X(J,I) - LIQUID MOLE FRACTION; J DENOTES COMPONENT I DENOTES I TH POINT
Y(J,I) - VAPOR PHASE MOLE FRACTION
G(J,I) - EXPERIMENTAL VALUES OF THE ACTIVITY COEFFICIENT
P(I) - TOTAL PRESSURE OF THE SYSTEM
T(I) - TEMPERATURE OF THE SYSTEM
XF(J,I) - SOLVENT MOLE FRACTIONS(ELECTROLYTE FREE)
FSM(J) - PURE COMPONENT VAPOR PRESSURE OF THE COMPONENT J AT THE SYSTEM TEMPERATURE.
DIMENSION X(3,99),XF(3,99),AMW(3),G(3,99),Y(3,99),P(99),
% ADD(6),T(99),PSM(3),XMOL(99),DENS(20),XDD(20)
DIMENSION CV2(3),CV3(3),CP2(6),CP3(6),ADT(2,20),XD(20)

WRITE(6,1000)
1000 FORMAT(//'INPUT DATA','
READ(5,2000)NCOMP,NF,NXD,NXDD
WRITE(6,2000)NCOMP,NF,NXD,NXDD
READ(5,2100)(CP2(I),I=1,6)
READ(5,2100)(CP3(I),I=1,6)
WRITE(6,2100)(CP2(I),I=1,6)
WRITE(6,2100)(CP3(I),I=1,6)
2100 FORMAT(12.7,F11.5,F8.3,F13.9,F3.1,F10.5)
READ(5,2200)(AMW(I),I=1,NCOMP)
WRITE(6,2200)(AMW(I),I=1,NCOMP)
READ(5,2200)V21,T21,V22,T22,V23,T23
WRITE(6,2200)V21,T21,V22,T22,V23,T23
READ(5,2200)V31,T31,V32,T32,V33,T33
WRITE(6,2200)V31,T31,V32,T32,V33,T33

CALCULATION OF LIQUID MOLAR VOLUME CONSTANTS USING
THREE VALUES OF LIQUID MOLAR VOLUME AT THREE
DIFFERENT TEMPERATURES.

CV2(3)=((T23-T21)*(V22-V21)-(V23-V21)*(T22-T21))/((T23**2-
% T21**2,)*(T23-T21)-(T33**2,-T31**2,))*(T22-T21))
CV2(2)=(V22-V21)-CV2(3)*((T22**2,-T21**2,))/T22-T21)
CV2(1)=V21-CV2(2)*T21-CV2(3)*T21**2,0
CV3(3)=((T33-T31)*(V32-V31)-(V33-V31)*(T32-T31))/((T32**2,-
% T31**2,))*(T33-T31)-(T33**2,-T31**2,))*(T32-T31))
CV3(2)=(V32-V31)-CV3(3)*((T32**2,-T31**2,))/T32-T31)
CV3(1)=V31-CV3(2)*T31-CV3(3)*T31**2,0
WRITE(6,5000)
5000 FORMAT(//'LIQUID MOLAR VOLUME CONSTANTS','
WRITE(6,5010)CV2(1),CV2(2),CV2(3)
WRITE(6,5010)CV3(1),CV3(2),CV3(3)
5010 FORMAT(5X,3(G12.5,3X))
5020 FORMAT(5X,6G12.5)
DO 10 I=1,NXD
READ(5,2800)ADT(1,I),ADT(2,I),XD(I)
WRITE(6,2800)ADT(1,I),ADT(2,I),XD(I)
2800 FORMAT(3F10.5)
10 CONTINUE
EXPERIMENTAL M-P-T-X-Y-MEAN MOLAL ACTIVITY COEFFICIENT DATA

FOLLOWING INSTRUCTIONS ARE IMPORTANT

1. IF DATA ARE MOLALITY VS MEAN MOLALITY ACTIVITY COEFFICIENT ONLY, READ P(I)=1.0

2. IF DATA ARE MOLALITY VS VAPOR PRESSURE ONLY, READ G(1,I)=1.0

3. IF A BINARY DATA ARE USED, READ X AND Y VALUES OF THE OTHER SOLVENT (WHICH IS NOT PRESENT) = 0.0

DO 4100 I=1,NP
READ(5,2500) XMOL(I),(X(J,I),Y(J,I),J=2,NCOMP),G(1,I),P(I),T(I)
WRITE(6,2501) PSM(2),PSM(3)

CALCULATION OF THE EXPERIMENTAL ACTIVITY COEFFICIENT OF SOLVENTS. ASSUMING POYNTING EFFECT = 1.0 & FUGACITY COEFFICIENT = 1.0.

SUM=1.0
DO 4200 K=2,NCOMP
SUM=SUM-X(K,I)
CONTINUE
X(1,I)=SUM/FK
XSUM=0.0
DO 4300 K=2,NCOMP
XSUM=XSUM+X(K,I)
IF(P(I),EQ.1.0) GO TO 3100
IF(X(K,I),EQ.0.0) GO TO 3100
G(K,I)=Y(K,I)*P(I)/X(K,I)/FSM(K)
GO TO 4300.
3100 G(K,I)=1.0
4300 CONTINUE
DO 4400 K=2,NCOMPS
XF(K,I)=X(K,I)/XSUM
4400 CONTINUE
CONTINUE
WRITE(6,2601)
2601 FORMAT(18X,’X1’,'10X’,’X2’,’12X’,’X3’,’9X’,’GAMA1’,’8X’,’GAMA2’,’8X’,’GAMA3’)
DO 4500 I=1,6F
WRITE(6,2600)X(I),Y(I),X(I),G(1,I),G(2,I),G(3,I)
2600 FORMAT(10X,6F15.6)
4500 CONTINUE
RETURN
END
SUBROUTINE FITIT(NPOINT, X1, AGAMA, A)

THIS PROGRAM FITS A POLYNOMIAL OF DEGREE 5

DIMENSION SIGMAY(20), X1(20), AGAMA(20), A(6), DELTAY(20), YCAL(20)

IF(NPOINT.LE.3) GO TO 99

DO 2 I = 1, NPOINT
   SIGMAY(I) = 0
   CONTINUE

NCODE = 0
MAXORD = 5
IF(NPOINT.LE.4) MAXORD = 2
IF(MAXORD.GT.6) MAXORD = 6

NNK = 1
DO 3 K = 5, NNK
   K1 = K + 1
   CALL POLIFI(X1, AGAMA, SIGMAY, NPOINT, K1, 0, A, CHISQR)

ERROR = 0.0
DO 4 J = 1, NPOINT
   SUM = A(1)
   DO 5 I = 2, K1
      SUM = SUM + A(I) * X1(J)**(I - 1)
      CONTINUE
   YCAL(J) = SUM
   DELTAY(J) = YCAL(J) - AGAMA(J)
   ERROR = ERROR + DELTAY(J)**2
   ERROR = ERROR / NPOINT
   CONTINUE

99 RETURN
END
SUBROUTINE POLIFI(X,Y,SIGMAY,NPTS,TERMS,MODE,A,CHISQ)

EXTRACTED FROM: BEVINGTON,P. R., "DATA REDUCTION AND ERROR ANALYSIS FOR THE PHYSICAL SCIENCES", McGRAW HILL, 1969

SUBROUTINE POLIFIT PURPOSE

MAKE A LEAST-SQUARES FIT TO DATA WITH A POLYNOMIAL CURVE
Y = A(1) + A(2)*X + A(3)*X**2 + A(4)*X**3 + ...

DESCRIPTION OF PARAMETERS
X - ARRAY OF DATA POINTS FOR INDEPENDENT VARIABLE
Y - ARRAY OF DATA POINTS FOR DEPENDENT VARIABLE
SIGMAY - ARRAY OF STANDARD DEVIATIONS FOR Y DATA POINTS
NPTS - NUMBER OF PAIRS OF DATA POINTS
TERMS - NUMBER OF COEFFICIENTS (DEGREE OF POLYNOMIAL + 1)
MODE - DETERMINANTS METHOD OF WEIGHTING LEAST-SQUARES FIT

+1 (INSTRUMENTAL) WEIGHT(I)=1./SIGMAY(I)**2
0 (NO WEIGHTING) WEIGHT =1.
-1 (STATISTICAL) WEIGHT(I) = 1./Y(I)
A - ARRAY OF COEFFICIENTS OF POLYNOMIAL
CHISQ - ARRAY OF REDUCED CHI SQUARE FOR FIT

SUBROUTINES AND FUNCTION SUBPROGRAMS REQUIRED
DETERM (ARRAY, NORDER)
EVALUATES THE DETERMINANTS OF A SYMMETRIC TWO-DIMENSIONAL MATRIX OF NORDER

DOUBLE PRECISION SUMX, SUMY, XTERM, YTERM, ARRAY, CHISQ
DIMENSION X(20), Y(20), SIGMAY(20), A(6)
DIMENSION SUMX(20), SUMY(20), ARRAY(8,8)

ACCUMULATE WEIGHTING SUMS

11 NMAX = 2*TERMS - 1
DO 13 N=1, NMAX
13 SUMX(N) = 0.
DO 15 J=1, TERMS
15 SUMY(J)= 0.
CHISQ =0.
DO 50 I=1, NPTS
XI=X(I)
YI= Y(I)
31 IF (MODE) 32,37,39
32 IF(YI) 35,37,33
33 WEIGHT = 1./YI
GO TO 41
35 WEIGHT = 1./(-YI)
GO TO 41
37 WEIGHT = 1.
GO TO 41
39  WEIGHT = 1./SIGMA(Y)*2
41  XTERM = WEIGHT
   DO 44 N=1,NMAX
44  SUMX(N) = SUMX(N) + XTERM
   XTERM = XTERM * XI
   DO 48 N=1,NTERMS
48  SUMY(N) = SUMY(N) + YTERM
   YTERM = YTERM * XI
   CHISQ = CHISQ + WEIGHT*YI**2
   CONTINUE
C
C CONSTRUCT MATRICES AND CALCULATE COEFFICIENTS
C
51  DO 54 J=1,NTERMS
      DO 54 K=1,NTERMS
54     N = J + K - 1
     ARRAY(J,K) = SUMX(N)
     DELTA = DETERM(ARRAY,NTERMS)
     IF(DELTA) 61,57,61
57    CHISQR = 0.
     DO 59 J=1,NTERMS
59    A(J) = 0.
     GO TO 80
C
61  DO 70 L=1,NTERMS
62     DO 66 J=1,NTERMS
65     ARRAY(J,K) = SUMX(N)
66     ARRAY(J,L) = SUMY(J)
67     A(L) = DETERM(ARRAY,NTERMS)/DELTA
     C
C CALCULATES CHI SQUARE
C
71  DO 75 J=1,NTERMS
    CHISQ = CHISQ - 2.*A(J)*SUMY(J)
    DO 75 K=1,NTERMS
75     N = J + K - 1
     CHISQ = CHISQ + A(J)*A(K)*SUMX(N)
76    FREE = NPTS - NTERMS
77    CHISQR = CHISQ/FREE
80   RETURN
END
FUNCTION DETERM(ARRAY,NORDER)

EXTRACTED FROM: BEVINGTON, P. R., "DATA REDUCTION AND ERROR ANALYSIS FOR THE PHYSICAL SCIENCES", McGRAW HILL, 1969

FUNCTION DETERM

PURPOSE
CALCULATES THE DETERMINANT OF A SQUARE MATRIX

USAGE
DET = DETERM(ARRAY,NORDER)

DESCRIPTION OF PARAMETERS
ARRAY - MATRIX
NORDER - ORDER OF DETERMINANT (DEGREE OF MATRIX)

SUBROUTINE AND FUNCTION SUBPROGRAMS REQUIRED
NONE

COMMENTS
THIS SUBPROGRAM DESTROYS THE INPUT MATRIX ARRAY

DOUBLE PRECISION ARRAY, SAVE
DIMENSION ARRAY(8,8)
DETERM = 1.

10 DO 50 K=1, NORDER

20 IF (ARRAY(K,K)) 41, 21, 41

21 DO 23 J=K, NORDER
23 IF (ARRAY(K,J)) 31, 23, 31

24 CONTINUE

25 DETERM = 0.
26 GO TO 60

31 DO 34 I=K, NORDER
32 SAVE = ARRAY(I,J)
33 ARRAY(I,J)=ARRAY(I,K)
34 ARRAY(I,K)=SAVE
35 DETERM = -DETERM

40 SUBTRACT ROW K FROM LOWER ROWS TO GET DIAGONAL MATRIX

41 DETERM = DETERM*ARRAY(K,K)
42 IF(K = NORDER) 43, 50, 50
43 K1 = K+1
44 DO 46 I=K1, NORDER
45 DO 46 J=K1, NORDER
46 ARRAY(I,J)=ARRAY(I,J)-ARRAY(I,K)*ARRAY(K,J)/ARRAY(K,K)

50 CONTINUE
60 RETURN
END
SUBROUTINE VAPRE(CP2, CP3, PSM, T)

******************************************************************************

SUBROUTINE VAPRE

THIS SUBROUTINE CALCULATES PURE COMPONENT VAPOUR-PRESSURES
OF THE SOLVENTS AT THE SYSTEM TEMPERATURE.

******************************************************************************

DIMENSION CP2(6), CP3(6), PSM(3)

PSM(2) = (EXP(CP2(1) + CP2(2)/(CP2(3) + T) + CP2(4)*T +
% CP2(5)**2 + CP2(6)*ALOG(T)))*760,

PSM(3) = (EXP(CP3(1) + CP3(2)/(CP3(3) + T) + CP3(4)*T +
% CP3(5)**2 + CP3(6)*ALOG(T)))*760,

RETURN

END
SUBROUTINE TEMPD(CV2,CV3,NXD,XD,ADT,AD,VO,T)

******************************************************************************

SUBROUTINE TEMPD

THIS SUBROUTINE CALCULATES THE TEMPERATURE DEPENDENCY OF
THE LIQUID MOLAR VOLUME AND DIELECTRIC CONSTANT.

CALLS SUBROUTINE FITIT FOR CONCENTRATION DEPENDENCE
OF DIELECTRIC CONSTANT.

******************************************************************************

DIMENSION CV2(3),CV3(3),XD(20),ADT(2,20),AD(6),VO(3),DSXD(20)

VO(2)=CV2(1)+CV2(2)*T+CV2(3)*T**2.0
VO(3)=CV3(1)+CV3(2)*T+CV3(3)*T**2.0
DO 10 I=1,NXD
     DSXD(I)=EXP(2.303*(ADT(1,I)+ADT(2,I)*T(293.15)))
  10 CONTINUE
CALL FITIT(NXD,XD,DSXD,AD)
RETURN
END
SUBROUTINE LSQ2(XT,X,DX,Y,M,H1,M3,L,E)

*************************************************************************
SUBROUTINE LSQ2

NON-LINEAR REGRESSION SUBROUTINE CALLS SUBROUTINE FN FOR
THE OBJECTIVE FUNCTION.
*************************************************************************

DIMENSION XT(6),DX(6),X(7,10),Y(7),JJ(3),A(3,3)
IL=0
IH=0
LIC=0
IF(L.LE.0) GO TO 50
IHC=M1+1
EN=M
EN=EN*1.5
L1=L
L=-L
L2=(3*M)/2+5
K3=2
IF(M.GE.3)K3=3
K4=K3-1
D=K3*K2
G=1.0/G
DO 100 I=1,M
  100 X(I,1)=XT(I)
CALL FN(Y(1),XT,LIC)
DO 106 J=2,M1
  106 XT(J-1)=XT(J-1)+DX(J-1)
DO 104 I=1,M
  104 X(I,J)=XT(I)
CALL FN(Y(J),XT,LIC)
XT(J-1)=X(J-1,1)
106 CONTINUE
L2C=0
FLG=1.0
GO TO 50
108 LIC=LIC+1
IF(LIC.GE.L1)GO TO 400
  50 YL=1.0E38
  YH=-YL
  Y2=YH
  ?
Y3 = YL
DO 110 J = 1, M
IF (Y(J) .LT. YH) GO TO 109
Y2 = YH
I2 = IH
YH = Y(J)
IH = J
GO TO 109

109 IF (Y(J) .LT. Y2) GO TO 109
Y2 = Y(J)
I2 = J

109 IF (Y(J) .GT. YL) GO TO 110
Y3 = YL
I3 = IL
YL = Y(J)
IL = J
GO TO 110

110 IF (Y(J) .GT. Y3) GO TO 110
Y3 = Y(J)
I3 = J
CONTINUE
L2C = L2C + 1
IF (L2C .LT. L2) GO TO 111
L2C = 0
JJ(1) = IL
JJ(2) = I2
JJ(3) = I3
DO 60 K1 = 1, K3
J1 = JJ(K1)
DO 60 K2 = K1, K3
J2 = JJ(K2)
S = 0.0
55 S = S + (X(I, J1) - X(I, IH)) * (X(I, J2) - X(I, IH))

60 A(K1, K2) = S
D = A(1, 1) * A(2, 2) - A(1, 2)**2
GO TO (62, 61), K4

61 D1 = A(1, 1) * A(2, 3) - A(1, 2) * A(1, 3)
IF (A(1, 1) .EQ. 0.0) A(1, 1) = 1.0E-5
D = ((A(1, 1) * A(3, 3) - A(1, 3)**2) * D - D1*D1) / (A(1, 1)**9.0)
62 IF (D .EQ. 0.0) GO TO 65
IF (D .LE. 0.0) D = ABS(D)
D = (D/4.0)**4
IF (D .LT. E) GO TO 65
FLG = 1.0
GO TO 111
65 IF (FLG .LT. 0.0) GO TO 400
FLG = -1.0
111 DO 115 I = 1, M
XT(I) = 0.0
DO 112 J=1,M1
IF(J.NE.IH)XT(I)=XT(I)+X(I,J)
112 CONTINUE

115 XT(I)=(3.0*XT(I)+X(I,I2)-X(I,IL))/EN-X(I,IH)
117 CALL FN(YT,XT,LIC)
119 IF(YT.GE.Y2)GO TO 167
117 IHC=M1+1
119 IF(YT.GE.YL)GO TO 140
121 YTT=YT
123 DO 135 I=1,M
125 XT(I)=1.5*XT(I)-0.5*X(I,IH)
127 CALL FN(YT,XT,LIC)
129 IF(YT.LE.YL)GO TO 140
131 DO 138 I=1,M
133 X(I,IH)=(2.0*XT(I)+X(I,IH))/3.0
135 Y(IH)=YTT
137 GO TO 108
139 DO 142 I=1,M
141 X(I,IH)=XT(I)
143 Y(IH)=YT
145 GO TO 108
147 IHC=IHC-1
149 IF(IHC.EQ.0)GO TO 300
151 IF(YT.GE.YH)GO TO 173
153 DO 168 I=1,M
155 XS=XT(I)
157 XT(I)=X(I,IH)
159 X(I,IH)=XS
161 DO 174 I=1,M
163 XT(I)=0.75*X(I,IH)+0.25*XT(I)
165 CALL FN(YT,XT,LIC)
167 IF(YT.GT.YH)GO TO 180
169 Y(IH)=YT
171 DO 175 I=1,M
173 X(I,IH)=XT(I)
175 GO TO 108
177 DO 185 J=1,M1
179 IF(J.EQ.IL)GO TO 185
181 DO 182 I=1,M
183 XT(I)=(X(I,J)+X(I,IL))/2.0
185 X(I,J)=XT(I)
187 CALL FN(Y(J),XT,LIC)
189 CONTINUE
200 GO TO 108
300 IHC=2*M1
302 IF(M.GE.3)GO TO 350
304 S=0.0
306 DO 302 I=1,M
308 X(I,M+2)=X(I,IH)-X(I,IL)
310 X(I,M+3)=X(I,IH)-X(I,I3)
164

302  S=S+X(I,M+2)**2
303  S=SQRT(S)
   IF(S.EQ.0.0)S=1.0E-5
304  U=-X(2,M+2)/S
   X(2,M+2)=X(1,M+2)/S
   X(1,M+2)=U
   S=X(1,M+2)**2+X(1,M+3)+X(2,M+2)**2+X(2,M+3)
   DO 305 I=1,M
305  X(I,M+2)=X(I,M+2)*S
306  DO 307 I=1,M
307  XT(I)=X(I,M)+X(I,M+2)
   CALL FN(YT,XT,LIC)
   DO 309 I=1,M
309  XT(I)=X(I,M)+X(I,M+2)
   CALL FN(YTT,XT,LIC)
   IF(YTT.LE.YT)GO TO 320
   DO 311 I=1,M
311  XT(I)=X(I,M)+X(I,M+2)
   YTT=YT
320  Y(I,H)=YTT
   DO 321 I=1,M
321  X(I,H)=XT(I)
   GO TO 108
350  DO 352 I=1,M
   XT(I)=X(I,H)-X(I,IL)
   X(I,M+2)=X(I,H)-X(I,I2)
352  X(I,M+3)=X(I,H)-X(I,I3)
   S=0.0
   S1=0.0
   DO 355 I=1,M
355  S=S+XT(I)**2
   S=SQRT(S)
   S1=SQRT(S1)
   S2=0.0
357  DO 357 I=1,M
   IF(S.EQ.0.0)S=1.0E-5
   XT(I)=XT(I)/S
   S2=S2+XT(I)**2+X(I,M+2)
   IF(S1.EQ.0.0)S1=1.0E-5
357  X(I,M+3)=X(I,M+3)/S1
   DO 360 I=1,M
360  X(I,M+2)=X(I,M+2)-XT(I)*S2
   S1=0.0
   DO 362 I=1,M
362  S1=S1+X(I,M+2)**2
   S1=SQRT(S1)
   DO 365 I=1,M
365  IF(S1.EQ.0.0)S1=1.0E-5
365  X(I,M+2)=X(I,M+2)/S1
S1 = 0.0
S2 = 0.0
DO 367 I = 1, M
S1 = S1 + X(I) * X(I, M+3)
367 S2 = S2 + X(I, M+2) * X(I, M+3)
DO 370 I = 1, M
370 X(I, M+2) = S * (S1 * X(I) + S2 * X(I, M+2) - X(I, M+3))
GO TO 306

400 S = Y(1)
Y(1) = Y(IL)
Y(IL) = S
DO 402 I = 1, M
XT(I) = X(I, IL)
X(I, IL) = X(I, 1)
402 X(I, 1) = XT(I)
WRITE(6, 772) LIC
772 FORMAT(' ', LIC=' ,15)
RETURN
END
SUBROUTINE FN(YY,XT,LIC)

*************

SUBROUTINE FN

*************

THIS SUBROUTINE IS CALLED BY SUBROUTINE LSQ2. THIS
SUBROUTINE ESTABLISHES THE OBJECTIVE FUNCTION BY CALLING
SUBROUTINE MINFUN.

*************

*GCAL(J,I) - CALCULATED ACTIVITY COEFFICIENT

*GADD(J,I) - LN OF THE SALTING OUT CONTRIBUTION TO THE
ACTIVITY COEFFICIENT.

*GEL(J,I) - LN OF DEBYE HUCKEL CONTRIBUTION AND THE
ADDITIONAL TERMS FOR THE COULOMBIC INTERACTION
TO THE ACTIVITY COEFFICIENT.

*GNRT(J,I) - LN OF THE NRTL CONTRIBUTION TO THE ACTIVITY
COEFFICIENT.

*PTC(I) - CALCULATED TOTAL PRESSURE OF THE SYSTEM

*YCAL(J,I) - CALCULATED VAPOR PHASE COMPOSITION.

*************

COMMON XMOL(99),X(3,99),XF(3,99),GG(3,99),P(99),YY(3,99),GNRT(3,99),GEL(3,99),GCAL(3,99),AMW(3),ERROR(3,99),GDH(3,99)

%GPHY(3,99),T(99),BMM(99),GADD(3,99),YCAL(3,99),ADD(6)

%PTC(99)

COMMON NP,INDF,FNP,FNM,FZP,FZN,FK,ALFA,DG23,DG32,

%GPN2,GPN3,ZP2,ZP3,KP,M,NBIN,PNION,NNION,DIELT

COMMON NDEN,NRTL,ALFA2,ALFB2,ALFA3,ALFB3,DGA2,DGB2,

%DGA3,DGB3,NBROM,NREG,NTYPE,NMIN

COMMON B012,B112,B013,B113,CP2(6),CP3(6),CV2(3),CV3(3),

%B123,ADT(2,20),NXD,XD(20)

DIMENSION XT(6),PSM(3)

IF(NREG.LE.1)GO TO 1070

GO TO (1040,1050,1060,1080,1085),NBIN

1040 GO TO(10,20),NRTL

10 DGA2=XT(1)

DGB2=XT(2)

GO TO 1070

20 IF(XT(1),LT,0.0)XT(1)=0.0
GPN2=XT(1)
ZF2=XT(2)
GO TO 1070

C

1050 GO TO (30, 40), NNRTL
30 DGA3=XT(1)
DBB3=XT(2)
GO TO 1070

C

40 IF(XT(1).LT.0.0)XT(2)=0.0
GPN3=XT(1)
ZF3=XT(2)
GO TO 1070

C

1060 DELTA=XT(2)
B123=XT(1)
GO TO 1070

C

1080 DG23=XT(1)
DG32=XT(2)
GO TO 1070

C

1085 GO TO (50, 60), NNRTL
50 DGA2=XT(1)
DBB2=XT(2)
DGA3=XT(3)
DBB3=XT(4)
GO TO 1070

C

60 IF(XT(1).LT.0.0)XT(1)=0.0
IF(XT(3).LT.0.0)XT(3)=0.0
GPN2=XT(1)
ZPN2=XT(2)
GPN3=XT(3)
ZPN3=XT(4)

C

1070 IF(KP, GT, 1) GO TO 3100
CALL NRTL1(XT, GT, GNRT, FK, FNP, NP, ALFA, DG23, DG32)
CALL BROML(XMOL, XF, AMW, FNP, FNM, FK, FZF, FZM, T, GEL,
% NP, B123, GDH, GPHY, BMN, ADD, NDEN, CV2, CV3, NXD, XD, ADT, B012, B112
% B013, B113)
CALL ADITON(XF, T, AMW, XMOL, FK, FZP, FZN, FNP, FNM, GADD,
% DELTA, NPION, NNION, NP, ADD, NDEN, CV2, CV3, NXD, XD, ADT,
% B012, B112, B013, B113)

C

3100 CALL NRTL2(XT, GNRT, FK, FNP, NP, GPN2, GPN3, ZP2, ZP3, ALFA, DG23,
% DG32, FNM, NNRTL, ALFA2, ALFA3, ALFB3, DGA2, DGB2,
% DGA3, DGB3)
CALL DEBHUC(XMOL, XF, X, FK, FNP, FNM, FZP, FZN, GEL, AHW,
% T, NP, GDH, GPHY, ADD, NDEN, CV2, CV3, NXD, XD, ADT)
DO 8010 IJ=1, NP
GADD(1, IJ)=0.0
GADD(2, IJ)=0.0
GADD(3, IJ)=0.0
CONTINUE

C

8010 CONTINUE

C

3200 YS=0.0
DO 4000 I=1, NP
IF(XF(2,I),EQ.0.0)GO TO 1000
IF(XF(3,I),EQ.0.0)GO TO 1010
GO TO 1020

1000 GNRT(2,I)=0.0
GEL(2,I)=0.0
GADD(2,I)=0.0
GADD(1,I)=0.0
GO TO 1020

1010 GNRT(3,I)=0.0
GEL(3,I)=0.0
GADD(3,I)=0.0
GADD(1,I)=0.0

1020 AMW = XF(2,I)*AMW(2)+XF(3,I)*AMW(3)
CALL VAPPRE(CP2,CP3,PSM,T(I))
GCAL(1,I)=EXP(GNRT(1,I)+GEL(1,I)+GADD(1,I)-ALOG(0.001*FK*AMW(I))+X*1.0)
GCAL(2,I)=EXP(GNRT(2,I)+GEL(2,I)+GADD(2,I))
GCAL(3,I)=EXP(GNRT(3,I)+GEL(3,I)+GADD(3,I))
PTC(I)=X(2,I)*GCAL(2,I)*PSM(2)+X(3,I)*GCAL(3,I)*PSM(3)
YCAL(I)=X(2,I)*GCAL(2,I)*PSM(2)/PTC(I)
YCAL(I)=X(3,I)*GCAL(3,I)*PSM(3)/PTC(I)
ERROR(I)=((GG(I,I)-GCAL(I,I))/GG(I,I))*100.
ERROR(I)=((GG(I,I)-GCAL(I,I))/GG(I,I))*100.
ERROR(I)=((GG(I,I)-GCAL(I,I))/GG(I,I))*100.
IF(XMOL(I),EQ.0.0)ERROR(I)=0.0
IF(X(2,I),EQ.0.0)ERROR(2,I)=0.0
IF(X(3,I),EQ.0.0)ERROR(3,I)=0.0
IF(GG(I,I),EQ.1.0)ERROR(I)=0.0
IF(GG(I,I),EQ.1.0)ERROR(2,I)=0.0
IF(GG(I,I),EQ.1.0)ERROR(3,I)=0.0
DIF1=ABS(ERROR(I)/100.)
DIF2=ABS(ERROR(2,I)/100.)
DIF3=ABS(ERROR(3,I)/100.)
DIF4=ABS(YCAL(I)-YV(3,I))
DIF5=ABS((PTC(I)-F(I))/F(I))
IF(NMIN,GT,3)GO TO 2000
GO TO 2010

2000 IF(X(2,I),EQ.0.0)GO TO 2020
IF(X(3,I),EQ.0.0)GO TO 2030
GO TO 2010

2020 DIF3=ABS((P(I)-PTC(I))/(PSM(3)-P(I)))
IF(P(I),EQ.1.0)DIF3=0.0
GO TO 2010

2030 DIF2=ABS((P(I)-PTC(I))/(PSM(2)-P(I)))
IF(P(I),EQ.1.0)DIF2=0.0

2010 CALL MNFUND(NTYPE,NMIN,DIF1,DIF2,DIF3,DIF4,DIF5,Y,X(2,I),Y(2,I),Y(3,I))

4000 CONTINUE
YY=YS
IF(LIC/20*20,NE,LIC)GO TO 301
WRITE(6,302) LIC

302 FORMAT(10X, 'TRIAL & USED = ', I4)
WRITE(6,303) YY, (XT(I), I=1, M)

303 FORMAT('//5X, 'YY='F15.7', '5X, 'XT VALUES', '6F15.6')

301 RETURN
END
SUBROUTINE FIBN(ALPHA,A,B)
C
C***************************************************************************
C SUBROUTINE FIBN
C THIS SUBROUTINE USES THE FIBONACCI METHOD TO FIND THE
C MINIMUM VALUE OF A NON LINEAR FUNCTION.
C***************************************************************************
C
DIMENSION FIB(50)
C SUBROUTINE FOR FIBONACCI PROCEDURE
C
DEL=B-A
WRITE(6,001)
001 FORMAT(//'','10X,35HFIBONACCI SINGLE-VARIABLE PROCEDURE')
C DEFINE THE FIRST THREE FIBONACCI NUMBERS
C
FIB0=1.0
FIB(1)=1.0
FIB(2)=2.0
C
C CALCULATE THE REMAINING FIBONACCI NUMBERS
C
BB=1.0/ALPHA
IF(BB-2.)10,10,11
10 GO TO 14
11 CONTINUE
JJ=2
12 JJ=JJ+1
FIB(JJ)=FIB(JJ-1)+FIB(JJ-2)
CC=FIB(JJ)
IF(CC-BB)13,15,15
13 GO TO 12
14 WRITE(6,002)
002 FORMAT(//'','10X,'ACCURACY SPECIFIED IN FUNC NOT SUFFICIENT.','
%//'','10X,'PROGRAM RESET ALPHA,ALPHA=0.01')
ALPHA=0.01
GO TO 5
C
C FIRST STEP IN THE TABLEAU
15 I=0
KK=JJ-2
IK=JJ-2
BL=B-A
?
ALL = FIB(IK) * BL / FIB(JJ)
W = A + ALL
V = B - ALL
CALL FUNC(W, T)
CALL FUNC(V, U)
JK = 1
WRITE (6, 003)
FORMAT (/ 1X, 1HK, 5X, 2HLK, 12X, 2HAK, 13X, 2HBK, 12X, 3HLLK, 9X, 1HX, 
% 8X, 1HY)
WRITE (6, 004) JK, BL, A, B, ALL, W, T
FORMAT (/ 13, 6G11.4)
FORMAT (41X, E12.4, 2X, E12.4)

C
C SUCCEEDING STEPS IN THE TABLEAU
C
IK = IK - 1
JJ = JJ - 1
DO 70 I = 1, KK
IF (U - T) 20, 20, 22
A = A + ALL
BL = B - A
W = V
CALL FUNC(W, T)
ALL = FIB(IK) * BL / FIB(JJ)
V = B - ALL
CALL FUNC(V, U)
II = I + 1
IK = IK - 1
JJ = JJ - 1
IF (IK - 1) 28, 29, 29
IK = 1
CONTINUE
WRITE (6, 004) II, BL, A, B, ALL, W, T
WRITE (6, 006) V, U
GO TO 70
B = B - ALL
BL = B - A
V = W
CALL FUNC(V, U)
ALL = FIB(IK) * BL / FIB(JJ)
W = A + ALL
CALL FUNC(W, T)
II = I + 1
IK = IK - 1
JJ = JJ - 1
IF (IK - 1) 30, 31, 31
IK = 1
CONTINUE
WRITE (6, 004) II, BL, A, B, ALL, V, U
WRITE (6, 006) W, T
GO TO 70
    CONTINUE
C
C  CALCULATION OF THE FINAL RANGE OF THE DEPENDENT VARIABLE
C
EPS=0.001*W
DL=W+EPS
CALL FUNC(DL,YL)
IF(YL-T)80,80,81
80  CALL FUNC(B,BF)
   WRITE(6,007)W,B
007 FORMAT(/,'#25HTHE FINAL FEASIBLE REGION,2X,2HX=',E15.4,2X,2HX=',
   %E15.4)
   WRITE(6,008)T,BF
008 FORMAT(/ '#20WITH FUNCTION VALUES,7X,2HY=',E10.4,2X,2HY=',E10.4)
   GO TO 87
81  CALL FUNC(A,AF)
   WRITE(6,009)W,A
009 FORMAT(/ '#25HTHE FINAL FEASIBLE REGION,2X,2HX=',E15.4,2X,2HX=',
   %E15.4)
   WRITE(6,017)T,AF
017 FORMAT(/ '#20WITH FUNCTION VALUES,7X,2HY=',E10.4,2X,2HY=',E10.4)
   ACC=(W-A)/(DEL)
   WRITE(6,018)ACC
018 FORMAT(/ '#15THE ACCURACY IS,12X,E10.4)
   WRITE(6,019)ALPHA
019  FORMAT/#'THE REQUIRED ACCURACY WAS =',E10.4)
C
RETURN
END
SUBROUTINE FUNC(XT,YY)

******************************************************************************
SUBROUTINE FUNC
******************************************************************************

THIS SUBROUTINE ESTABLISHES THE OBJECTIVE FUNCTION

SIMILAR TO SUBROUTINE FN. THIS IS CALLED BY

SUBROUTINE FIBN. ALL SYMBOLS ARE SAME AS IN

SUBROUTINE FN.
******************************************************************************

COMMON XMOL(99),XF(3,99),GG(3,99),P(99),YV(3,99),GNRT(3,99),GEL(3,99),GCAL(3,99),AMW(3),ERROR(3,99),GDH(3,99),GPHY(3,99),T(99),BMM(99),GADD(3,99),YCAL(3,99),ADD(6),PCT(99)

COMMON NDEN,NDEN2,NDEN3,NDEN4,NDEN5,NBRON,NREG,NTRIP,NMIX

COMMON N012,NO13,NO14,NO15,NO26,NO37,NO48,NO59,NO610

DIMENSION PSM(3)

GO TO (1040,1050,1060),NBIN

1040

B012=XT

DELTA=0.0
B112=0.0
B123=0.0
GO TO 1070

1050

B013=XT

DELTA=0.0
B113=0.0
B123=0.0
GO TO 1070

1060

GO TO (10,20),NBROM

10

B123=XT

GO TO 1070

20

DELTA=XT

1070

IF(KP.GE.1) GO TO 3100

CALL NRTL1(X,T,GNRT,FK,FNP,NP,ALFA,DOG23,DOG32)

CALL BROML(XMOL,XF,AMW,FNP,FNM,FK,FZP,FZN,T,GEL,
CALL ADTON(XF,T,AMW,XMOL,FK,FZP,FZN,FNP,FNM,GADD)
CALL DELTA,NPION,NNION,NP,ADD,NDEN,CV2,CV3,NXD,ADT
GO TO 3200
CALL NRTL2(X,T,GNRT,FK,FNP,GP,GPN3,ZPN2,ZPN3,ALFA,DG23,
DG32,FNM,NNRTL,ALFA2,ALFB2,ALFA3,ALFB3,DGA2,DGB2,DGA3,DGB3)
CALL DEBUC(XMOL,XF,X,FK,FNP,FNM,FZP,FZN,GEL,AMW,
T,NP,GDH,GPHY,ADD,NDEN,CV2,CV3/NXD,ADT)
DO 8010 IJ=1,NP
GADD(1,IJ)=0.0
GADD(2,IJ)=0.0
GADD(3,IJ)=0.0
8010 CONTINUE
DO 4000 I=1,NP
IF(XF(2,I).EQ.0.0)GO TO 1000
IF(XF(3,I).EQ.0.0)GO TO 1010
GO TO 1020
1000 GNRT(2,I)=0.0
GEL(2,I)=0.0
GNRT(1,I)=0.0
GADD(2,I)=0.0
GADD(1,I)=0.0
GO TO 1020
1010 GNRT(3,I)=0.0
GEL(3,I)=0.0
GNRT(1,I)=0.0
GADD(3,I)=0.0
GADD(1,I)=0.0
1020 AMS = XF(2,I)*AMW(2)+XF(3,I)*AMW(3)
GCAL(1,I)=EXP(GNRT(1,I)+GEL(1,I)+GADD(1,I)-ALOG(0.001*FK*AMS*XMO(I)+1.0))
GCAL(2,I)=EXP(GNRT(2,I)+GEL(2,I)+GADD(2,I))
GCAL(3,I)=EXP(GNRT(3,I)+GEL(3,I)+GADD(3,I))
CALL VAPPRE(CP2,CP3,PSM,T(I))
PTC(I)=X(2,I)*PSM(2)*GCAL(2,I)+X(3,I)*PSM(3)*GCAL(3,I)
YC(2,I)=X(2,I)*PSM(2)*GCAL(2,I)/PTC(I)
YC(3,I)=X(3,I)*PSM(3)*GCAL(3,I)/PTC(I)
ERROR(1,I)=(GG(1,I)-GCAL(1,I))/GG(1,I)*100.
ERROR(2,I)=(GG(2,I)-GCAL(2,I))/GG(2,I)*100.
ERROR(3,I)=(GG(3,I)-GCAL(3,I))/GG(3,I)*100.
IF(XMOL(I).EQ.0.0)ERROR(I,I)=0.0
IF(X(2,I).EQ.0.0)ERROR(2,I)=0.0
IF(X(3,I).EQ.0.0)ERROR(3,I)=0.0
IF(GG(1,I).EQ.1.0)ERROR(1,I)=0.0
IF(GG(2,I).EQ.1.0)ERROR(2,I)=0.0
IF(GG(3,I).EQ.1.0)ERROR(3,I)=0.0
F1=ABS(ERROR(I,I)/100.)
DIF2 = ABS(ERROR(2, I) / 100.)
DIF3 = ABS(ERROR(3, I) / 100.)
DIF4 = ABS(YCAL(3, I) - YV(3, I))
DIF5 = ABS((PTC(I) - P(I)) / P(I))
IF(NMIN.GT.3) GO TO 2000
GO TO 2010
2000 IF(X(2, I).EQ.0.0) GO TO 2000
IF(X(3, I).EQ.0.0) GO TO 2000
2020 DIF3 = ABS((P(I) - PTC(I)) / PSM(T, P(I)))
IF(P(I).EQ.1.0) DIF3 = 0.0
GO TO 2010
2030 DIF2 = ABS((P(I) - PTC(I)) / (PSM(2, P(I))))
IF(P(I).EQ.1.0) DIF2 = 0.0
2010 CALL MINFUN(NTYPE, NMIN, DIF1, DIF2, DIF3, DIF4, DIF5, Y, X(3, I))
Y = Y + Y
4000 CONTINUE
Y = Y
RETURN
END
SUBROUTINE NRTL1(X,T,GNRT,FK,FNP,NP,ALFA,DG23,DG32)

******************************************************************************

SUBROUTINE NRTL1

THIS SUBROUTINE CALCULATES SOLVENT-SOLVENT INTERACTION
FOR A TERNARY MIXTURE IN MODEL # 2.

******************************************************************************

DIMENSION XA(99),X(3,99),GNRT(3,99),T(99)
R=1.987
DO 4001 I=1,NP
G32=EXP(-ALFA*DG32/R/T(I))
Z32=DG32*G32
G23=EXP(-ALFA*DG23/R/T(I))
Z23=DG23*G23
XA(I)=FNP*X(1,I)
FNT1=(FK*XA(I)/FNP+X(2,I)+X(3,I)*G32)**2.
FNT2=(FK*XA(I)/FNP+X(3,I)+X(2,I)*G23)**2.
FNT3=X(2,I)+X(3,I)*G32
FNT4=X(3,I)+X(2,I)*G23
GNRT(1,I)=(-Z32/FNT1-Z32/FNT2+Z23/(FNT3**2.))*Z23/(FNT4**2.))*
% X(2,I)**(3,I)/R/T(I)
AT2=XA(I)**(3,I)*FK*(Z32/FNT1+Z32/FNT2+Z32/(FNT3**2.))+
% Z23/(FNT4**2.))/(FNP*R*T(I))
AT3=XA(I)**(2,I)*FK*(Z32/FNT1+Z23/FNT2+Z32/(FNT3**2.))*Z23
% /(FNT4**2.))/(FNP*R*T(I))
BT3=X(2,I)**2.*(Z32/FNT1+Z32/FNT2)/(R*T(I))
BT2=X(3,I)**2.*(G32*Z32/FNT1+Z23/FNT2)/(R*T(I))
CT2=-2.*FK*X(2,I)**(3,I)*XA(I)*(Z32/(FNT3**3.))+G32*Z23
% /(FNT4**3.))/(FNP*R*T(I))
CT3=-2.*FK*X(2,I)**X(3,I)*XA(I)*(Z32*G32/(FNT3**3.))+
% Z23/(FNT4**3.))/(FNP*R*T(I))
GNRT(2,I)=AT2+BT2+CT2
GNRT(3,I)=AT3+BT3+CT3
4001 CONTINUE
RETURN
END
SUBROUTINE BROML(XMOL, XF, AMW, FNP, FNM, FK, FZP, FZN, T, GBM, 
% NP, B123, GDH, GPHY, BM, ADD, NDEN, CV2, CV3, NXD, XD, ADT, B012, 
% B112, B013, B113)

******************************************************************************

SUBROUTINE BROML

THIS SUBROUTINE CALCULATES ION-ION INTERACTION &
ION-SOLVENT INTERACTION ACCORDING TO THE BROMLEY
EQUATION, IN MODEL # 2.

******************************************************************************

GDH(J, I) - CONTRIBUTION OF THE DEBYE-HUCKEL TERM
GPHY(J, I) - CONTRIBUTION OF THE B TERMS IN THE EQUATION
(i.e. ION-SOLVENT INTERACTION)
GBM(J, I) - COMBINATION OF THE ABOVE TWO TERMS

THIS SUBROUTINE CALLS SUBROUTINES FUNCT & FUNCB.

******************************************************************************

DIMENSION XMOL(99), XF(3,99), AMW(3), T(99)
DIMENSION SIG(99), SAI(99), SIG1(99), SAI1(99), GBM(3,99), 
% AI(99), F2(99), F3(99), BM(99), VO(3), AD(6)
% , GDH(3,99), GPHY(3,99)
DIMENSION ADD(6), CV2(3), CV3(3), XD(20), ADT(2,20)
ROW=1.0
DO 4001 I=1, NP
IF(XMOL(I).EQ.0.0)GO TO 2010
B12=B012*alog((T(I)-243.)/T(I))+B112/T(I)
B13=B013*alog((T(I)-243.)/T(I))+B113/T(I)
CALL TEMPD(CV2, CV3, NXD, XD, ADT, AD, VO, T(I))
AA=1.5/(FZP*FZN)
AI(I)=(FNP*FZP**2.+FNM*FZN**2.)*XMOL(I)/2.
XT1=1.*ROW*AI(I)**0.5
XT2=AA*AI(I)
XT3=1.*XT2
XT4=1.+2.*XT2
AMS=XF(2, I)*AMW(2)+XF(3, I)*AMW(3)
TNS=1000./AMS
XT5=(ROW*AI(I)**0.5)**3.
XT6=0.001*FK*XMOL(I)
XT7=XT6*AMS+1.
IF(XF(2,I),EQ,0.0)GO TO 1000
IF(XF(3,I),EQ,0.0)GO TO 1010
CALL FUNCT(XF(2,I),XF(3,I),T(I),TNS,F2(I),F3(I),
% ADB,FD2,FD3,AD,DS,ADD,V0,NDEN,AMW)
CALL FUNCT(XF(2,I),XF(3,I),B12,B13,BM(I)
% ,FB2,FB3,TNS,B123,A1(I),FNP,FNM,FZP,FZN,AMS,AMW(2),AMW(3)
% ,FB1,XMOL(I))
GO TO 1020

1000
BM(I)=B13
D30=AD(1)+AD(2)*XF(3,I)+AD(3)*XF(3,I)**2.+AD(4)*XF(3,I)**3.
% +AD(5)*XF(3,I)**4.+AD(6)*XF(3,I)**5.
D3=AMW(3)/VO(3)
ADB3=1.8246*10.*((D03*0.5)*((D30*T(I))**(-1.5))
FB1=0.0
FB2=0.0
FB3=0.0
F2(I)=0.0
F3(I)=0.0
FD2=0.0
FD3=0.0
ADB=ADB3
GO TO 1020

1010
BM(I)=B12
D20=AD(1)+AD(2)*XF(3,I)+AD(3)*XF(3,I)**2.0+AD(4)*XF(3,I)**3.
% +AD(5)*XF(3,I)**4.+AD(6)*XF(3,I)**5.
D20=AMW(2)/VO(2)
ADB2=1.8246*10.*((D02*0.5)*((D20*T(I))**(-1.5))
FB1=0.0
FB2=0.0
FB3=0.0
F2(I)=0.0
F3(I)=0.0
FD2=0.0
FD3=0.0
ADB=ADB2

1020
SIG(I)=3.*(XT1-1./XT1-2.*ALOG(XT1))/XT5
SAI(I)=2.*(ALOG(XT3)/XT2+XT4/(XT3**2.))/XT2
SIG1(I)=2.*(XT1**2./2.-2.*XT1+ALOG(XT1)+1.5)/XT5
SAI1(I)=0.6*2.0*FZP*FZN*(ALOG(XT3)/XT2-1./XT3)/XT2+1.0
YBT1=XT6*2.303*ADB*SIG(I)*AI(I)**(0.5)/3.*(FZP*FZN)
YBT2=2.303*FZP*FZN*XT6*(0.06+0.6*EH(I))*SAI(I)*AI(I)/2.
YBT3=2.303*XT6*BM(I)*AI(I)/2.
YBT4=2.303*XT6*TNS*AMS*SIG1(I)*AI(I)**(0.5)*(FZP*FZN)
YBT5=FB3
YBT7=FB2
YBT6=2.303*XT6*AMS*TNS*AI(I)/2.0
FT1=-YBT2-YBT3
FT5=XT6*AMS
FT6=ALOG(XT7)
FT7=FT6-FT5
GDH(1, I) = -2.303 * ADD * FZP * FZN * AI(I) * 0.5 / XT1
GDH(2, I) = YBT1 * AMW(2) - YBT4 * F2(I)
GDH(3, I) = YBT1 * AMW(3) - YBT4 * F3(I)
GPHY(1, I) = ((0.06 + 0.6 * BM(I)) * AI(I)) / (XT3 * 2.303)
GPHY(2, I) = FT1 * AMW(2) + YBT7 * YBT6 * SAI1(I)
GPHY(3, I) = FT1 * AMW(3) + YBT5 * YBT6 * SAI1(I)
GBM(1, I) = GDH(1, I) + GPHY(1, I) + ALOG(XT7)
GBM(2, I) = GDH(2, I) + GPHY(2, I) + FT7
GBM(3, I) = GDH(3, I) + GPHY(3, I) + FT7
GO TO 4001

2010 GBM(1, I) = 0.0
GBM(2, I) = 0.0
GBM(3, I) = 0.0
GDH(1, I) = 0.0
GDH(2, I) = 0.0
GDH(3, I) = 0.0
GPHY(1, I) = 0.0
GPHY(2, I) = 0.0
GPHY(3, I) = 0.0

4001 CONTINUE
RETURN
END
SUBROUTINE ADITON(XF,T,AMW,XMOL,FK,FZP,FZN,FNP,FNM)

****************************************************************

% GADD,DELTA,NPION,NNION,AP,ADD,NDEN,CV2,CV3,NXD,XD,ADT,B012
% ,B112,B013,B113
DIMENSION XMOL,99),XF(3,99),T(99),AMW(3),AD(6),GADD(3,99)
DIMENSION RP(6),RN(6),VO(3),ADT(2,20),CV2(3),CV3(3),ADD(6)
DIMENSION XD(20)

RP(1) = H ION
RP(2) = LI ION
RP(3) = NA ION
RP(4) = CA ION
RN(1) = BR ION
RN(2) = CL ION

RP(1)=2.08*10.**(-8.0)
RP(2)=0.60*10.**(-8.0)
RP(3)=0.95*10.**(-8.0)
RP(4)=0.99*10.**(-8.0)
RN(1)=1.95*10.**(-8.0)
RN(2)=1.81*10.**(-8.0)

CONST=1.6710383*10.**(-3.0)

DO 10 I=1,NP
ALFA=2.0
AN1=XMOL(I)
IF(XF(2,I).EQ.0.0)GO TO 20
IF(XF(3,I).EQ.0.0)GO TO 20

AMS=XM(2,I)*AMW(2)+XM(3,I)*AMW(3)
TNS=1000.0/AMS

CALL TEMP(CV2,CV3,NXD,XD,ADT,AD,VO,T(I))
B12=B012*ALO9((T(I)-243.)/T(I))+B112/T(I)
B13=B013*ALO9((T(I)-243.)/T(I))+B113/T(I)

CONST=FNPFZP**2./RP(NPION)+FNPFZN**2./RN(NNION)

CALL FUNCT(XF(2,I),XF(3,I),T(I),TNS,F2,F3)

% ADD,FD2,FD3,AD,DS,ADD,VO,NDEN,AMW)

TN3=XM(3,I)*TNS
TN2=XM(2,I)*TNS

****************************************************************
\[ D = \exp(-\alpha \times \text{AN1}^{0.5}) \]

\[ A = (\text{TN2} \times \text{TN3})^{0.5} \]

\[ B = \text{XF}(2, I) \times \text{B13} - \text{XF}(3, I) \times \text{B12} \]

\[ C = \exp(2.0 \times \text{XF}(2, I)) \]

\[ AX = (1.0 - \alpha \times \text{XMO}(I)) / (4.0 \times \text{AN1}^{0.5}) \]

\[ GADD(1, I) = \delta \times \text{D} \times \text{CONST} \times \text{CONST1} \times \text{A} \times \text{B} \times \text{C} \times \text{XMO}(I) \times AX / (FK \times \text{T}(I) \times DS) \]

\[ GADD(2, I) = \delta \times \text{D} \times \text{CONST} \times \text{CONST1} \times (\text{XMO}(I)^{0.5} \times 2.0) \times (\text{TN3} / \text{TN2})^{0.5} \times \text{B} \times \text{C} / 2.0 - \alpha \times \text{B} \times \text{C} \times \text{FD2} / \text{DS} + (\text{B12} + \text{B13}) \times \text{XF}(3, I) \times \text{A} \times \text{C} / \text{TNS} + \alpha \times \text{B} \times \text{C} \times 2.0 \times \text{XF}(2, I) / \text{TNS} / (\text{DS} \times \text{T}(I) \times 2.0) \]

\[ GADD(3, I) = \delta \times \text{D} \times \text{CONST} \times \text{CONST1} \times (\text{XMO}(I)^{2.0} \times 2.0) \times (\text{TN2} / \text{TN3})^{0.5} \times \text{B} \times \text{C} / 2.0 - \alpha \times \text{B} \times \text{C} \times \text{FD3} / \text{DS} - (\text{B12} + \text{B13}) \times \text{XF}(2, I) \times \text{A} \times \text{C} / \text{TNS} - \alpha \times \text{B} \times \text{C} \times 2.0 \times \text{XF}(2, I) / \text{TNS} / (\text{DS} \times \text{T}(I) \times 2.0) \]

\[ 20 \quad \text{GADD}(1, I) = 0.0 \]

\[ \text{GADD}(2, I) = 0.0 \]

\[ \text{GADD}(3, I) = 0.0 \]

\[ 10 \quad \text{CONTINUE} \]

\[ \text{RETURN} \]

\[ \text{END} \]
SUBROUTINE FUNCB(X2,X3,B12,B13,BM,FB2,FB3,TNS,B123, 
A1,FNP,FNM,FZP,FZN,AMS,ANW2,ANW3,FB1,XMOL)

THIS SUBROUTINE CALCULATES THE MIXTURE TERNARY BROMLEY 
CONSTANT "BM". B123 IS THE TERNARY ADJUSTABLE PARAMETER.

AN1=XMOL
D23=B123
A=(X2*X3)**0.25
ALFA=2.0
D=(1.0+ALFA*AN1**0.5)**3.0
D1=((1.0+ALFA*AN1**0.5)**4.0)*(AN1**0.5)
B=EXP(-ALFA*X3)
FDBN2=ALFA*X3*B/TNS
FDBN3=-ALFA*X2*B/TNS
FDAN2=((X3/X2)**0.25/(X2**0.5))/2.0-A)/TNS/2.0
FDAN3=((X2/X3)**0.25/(X3**0.5))/2.0-A)/TNS/2.0
BM=B12*X2+B13*X3+D23*A*B/D
FB2=(B12-B13)*X3/TNS+D23*(A*FDBN2+B*FDAN2)/D
FB3=(B13-B12)*X2/TNS+D23*(A*FDBN3+B*FDAN3)/D
FB1=-3.0*ALFA*D23*A*B/(2.0*D1)
RETURN
END
SUBROUTINE FUNCT(XF2, XF3, T, TNS, F2, F3, ADB)
%
FD2, FD3, AD, DS, ADD, VO, NDEN, AMW)

*************************************************************************

SUBROUTINE FUNCT

THIS SUBROUTINE CALCULATES THE DEBYE-HUCKEL CONSTANT

AND THE SLOPE OF D.H. CONSTANT WITH RESPECT TO # OF

MOLES OF SOLVENTS IN A TERNARY MIXTURE.

IF NDEN = 1 EXPERIMENTAL CONCENTRATION DEPENDENT DATA
ARE USED FOR DENSITIES.

IF NDEN = 2 APPROXIMATE RELATIONSHIP IS USED FOR DENSITIES

*************************************************************************

DIMENSION AD(6), ADD(6), VO(3), AMW(3)
DS=AD(1)+AD(2)*XF3+AD(3)*XF3**2.+AD(4)*XF3**3.+AD(5)*XF3**4.+
% +AD(6)*XF3**5.*
FT1=AD(2)+2.*AD(3)*XF3+3.4AD(4)*XF3**2.+4.*AD(5)*XF3**3.*
% +5.*AD(6)*XF3**4.*0
FD2=-FT1*XF3/TNS
FD3=FT1*XF2/TNS
GO TO 10, 20, NDEN

10 DSS=ADD(1)+ADD(2)*XF3+ADD(3)*XF3**2.+ADD(4)*XF3**3.+
% ADD(5)*XF3**4.*ADD(6)*XF3**5.0
FDT1=ADD(2)+2.*ADD(3)*XF3+3.0*ADD(4)*XF3**2.+4.0*ADD(5)*
% XF3**3.+5.*ADD(6)*XF3**4.*
DDS2=-FDT1*XF3/TNS
DDS3=FDT1*XF2/TNS
GO TO 30

20 AMWW=AMW(2)*XF2+AMW(3)*XF3
VS=VO(2)*XF2+VO(3)*XF3
DSS=AMWW/VS
DDS2=((AMW(2)-AMW(3))-AMWW*(VO(2)-VO(3))/VS)*XF3/TNS/VS
DDS3=((AMW(3)-AMW(2))-AMWW*(VO(3)-VO(2))/VS)*XF2/TNS/VS

30 ADB=1.8246*10.**0.7**0.5*(DS**T)**(-0.5)
F2=ADB*(DDS2/DSS/2.-3.*FD2/2./DS)
F3=ADB*(DDS3/DSS/2.-3.0*FD3/2./DS)
RETURN
END
SUBROUTINE NRTL2(X,T,GNRT,FK,FNP,NP,GNP2,GNP3,ZP2,ZP3,ALFA, 
% D23,D32,FNM,NNRTL,ALFA2,ALFB2,ALFA3,ALFB3, 
% DGA2,DGB2,DGA3,DGB3)

SUBROUTINE NRTL2

THIS SUBROUTINE CALCULATES ION-SOLVENT AND SOLVENT- 
SOLVENT INTERACTIONS OF THE ACTIVITY COEFFICIENTS IN 
MODEL #1.

THIS SUBROUTINE IS USED FOR BOTH BINARY & TERNARY MIXTURES

DG23,D32,Z23,Z32 - CAL/sMOLE-K

ZP2,ZP3 - K Joules/ sMOLE

DGA2,DGB2,DGA3,DGB3 - K Joules/sMOLE-K

DIMENSION X(3,99),XA(99),GNRT(3,99),T(99)
R=1.987
R=8.3143*10,**(-3.0)
DO 4001 I=1,NP
GO TO(10,20),NNRTL
GPN2=EXP(-ALFA2*DGA2/R1/T(I))+FNM*EXP(-ALFB2*DGB2 
% /R1/T(I))/FNP 
ZP2=DGA2*EXP(-ALFA2*DGA2/R1/T(I))+FNM*DGB2*EXP(-ALFB2*DGB2 
% /R1/T(I))/FNP 
GPN3=EXP(-ALFA3*DGA3/R1/T(I))+FNM*DGB3*EXP(-ALFB3*DGB3 
% /R1/T(I))/FNP 
ZP3=DGA3*EXP(-ALFA3*DGA3/R1/T(I))+FNM*DGB3*EXP(-ALFB3*DGB3/ 
% R1/T(I))/FNP 
20 ZPN2=ZP2*238.862
ZPN3=ZP3*238.862
G32=EXP(-ALFA*DG32/R/T(I))
Z32=DG32*G32 
G23=EXP(-ALFA*DG23/R/T(I))
Z23=DG23*G23
XA(I) = X(1,I) * FNP
DT1 = (XA(I) * GPN3 + X(2,I) * G23 + X(3,I)) ** 2.0
DT2 = (Xh(I) * GPN2 + X(3,I) * G32 + X(2,I)) ** 2.0
DT3 = X(3,I) * G32 + X(2,I)
DT4 = X(2,I) * G23 + X(3,I)
TT1 = ZPN2 * GPN2 * XA(I) ** 2 + XA(I) * X(3,I) * ZPN2 * G32 + XA(I) * X(3,I)
TT2 = XA(I) * X(3,I) * Z23 * GPN3 + Z23 * X(3,I) ** 2.
% - XA(I) * X(3,I) * ZPN3 * G32
TT3 = X(2,I) * (X(3,I) * G23 * ZPN2 + X(2,I) * ZPN2 - X(3,I) * Z32 * GPN2).
TT4 = X(3,I) * (X(2,I) * G23 * ZPN3 + X(3,I) * ZPN3 - X(2,I) * Z23 * GPN3).
TT5 = ZPN3 * GPN3 * XA(I) ** 2 + XA(I) * ZPN3 * G23 * X(2,I) + XA(I)
% * X(2,I) * Z23 * GPN3 + Z23 * G23 * X(2,I) ** 2.
TT6 = XA(I) * X(2,I) * Z32 * GPN2 + Z32 * X(2,I) ** 2 - XA(I) *
% X(2,I) * ZPN2 * G32
GNRT(1,I) = (TT3 / DT2 + TT4 / DT1 - XA(I) * X(3,I) * ZPN2 / DT3 - X(3,I) * ZPN3 / DT4
% + X(2,I) * X(3,I) * Z32 * GPN2 / (DT3 ** 2.0) + X(2,I) * X(3,I) * Z32 * GPN2 / DT4 ** 2.0
% (DT4 ** 2.0)) * FNP / R(T(I)) / FK
TT7 = (Z32 * GPN2 - ZPN2 * G32) / (DT3 ** 2.0) + (Z23 * GPN3 + ZPN3 * G23)
% / (DT4 ** 2.0)
TT8 = 2.0 * XA(I) * X(2,I) * X(3,I) * (Z32 * GPN2 / (DT3 ** 2.0) + Z32 * GPN3 * G23
% (DT4 ** 2.0)
TT9 = 2.0 * XA(I) * X(2,I) * X(3,I) * (G32 * Z32 * GPN2 / (DT3 ** 3.0) + Z23 * GPN3
% / (DT4 ** 3.0)
TT10 = (Z32 * GPN2 + ZPN2 * G32) / (DT3 ** 2.0) + (Z23 * GPN3 - G23
% * ZPN3) / (DT4 ** 2.0)
GNRT(2,I) = (TT1 / DT2 + TT2 / DT1 - XA(I) * X(3,I) * TT7 - TT8) / R(T(I))
GNRT(3,I) = (TT5 / DT1 + TT6 / DT2 + XA(I) * X(2,I) * TT10 - TT9) / R(T(I))
4001 CONTINUE
RETURN
END
SUBROUTINE DEBHUC

**THIS SUBROUTINE CALCULATES ION-ION INTERACTION USING THE EXTENDED DEBYE-HUCKEL EQUATION, IN MODEL # 1.**

GDH(J,I) - D.H CONTRIBUTION IN EQUATION
GPHY(J,I) - EXTENDED PART OF THE D.H EQUATION
GEL(J,I) - COMBINATION OF ABOVE TWO TERMS

DIMENSION VO(3), XF(3,99), X(3,99), XMOL(99), GEL(3,99), GDH(3,99),
% GPHY(3,99), AI(99), AMW(3), T(99), AD(6)
DIMENSION F2(99), F3(99)
DIMENSION SIG(99), SIG1(99), SAI(99), SAI1(99)
DIMENSION ADD(6), CV2(3), CV3(3), XD(20), ADT(2,20)
AA=1.5/(FZP*FZN)
DO 400 I=1,NP
IF(XMOL(I).EQ.0.0)GO TO 2010
AMS=XM(2,I)*AMW(2)+XM(3,I)*AMW(3)
TNS=1000./AMS
CALL TEMPD(CV2,F3,3,1), CV3, XD, ADT, AD, VO, T(I))
IF(XF(2,I).EQ.0.0)GO TO 1000
IF(XF(3,I).EQ.0.0)GO TO 1010
CALL FUNCT(XF(2,I), XF(3,I), T(I), TNS, F2(I), F3(I))
% ADB, F2, F3, DS, ADD, VO, NDEN, AMW)
GO TO 1020
1000 F2(I)=0.0
F3(I)=0.0
D30=ADD(1)+ADD(2)*XF(3,I)+ADD(3)*XF(3,I)**2.+ADD(4)*XF(3,I)**3.
1 **3.+ADD(5)*XF(3,I)**4.+ADD(6)*XF(3,I)**5.
GO TO (10,20), NDEN
10 D03=ADD(1)+ADD(2)+ADD(3)+ADD(4)+ADD(5)+ADD(6)
GO TO 30
20 D03=AMW(3)/VO(3)
30 F2(I)=0.0
F3(I)=0.0
ADB=1.8246*10.**(6.0)*(D03**0.5)*((D03*T(I))***(-1.5))
GO TO 1020
1010 F2(I)=0.0
F3(I) = 0.0
D20 = AD(1) + AD(2) * XF(3, I) + AD(3) * XF(3, I) * 2 + AD(4) * XF(3, I) * 3 + AD(5) * XF(3, I) * 4 + AD(6) * XF(3, I) * 5.

* 4001 CONTINUE
SUBROUTINE MINFUN(NTYPE, NMIN, DIF1, DIF2, DIF3, DIF4, DIF5, 
% Y, X2, X3)

*******************************************************************************************************

SUBROUTINE MINFUN

THIS SUBROUTINE SETS UP DIFFERENT OBJECTIVE FUNCTIONS USING A SYMBOL "NMIN". THE EXPLANATION HAS BEEN GIVEN IN MAIN PROGRAM.

*******************************************************************************************************

GO TO(10, 10, 10, 20, 20, 20, 20, 20, 20, 20, 20), NTYPE
10 DIF4=0.0
    DIF5=0.0
    GO TO 60
20 GO TO(30, 40, 50, 40), NMIN
30 DIF4=0.0
    DIF5=0.0
    GO TO 60
40 IF (X2.EQ.0.0) GO TO 70
    IF (X3.EQ.0.0) GO TO 70
    DIF1=0.0
    DIF2=0.0
    DIF3=0.0
    DIF4=DIF4*10.0
    GO TO 60
70 DIF4=0.0
    DIF5=0.0
    GO TO 60
50 IF (X2.EQ.0.0) GO TO 80
    IF (X3.EQ.0.0) GO TO 80
    DIF5=0.0
    DIF4=DIF4*10.0
    GO TO 60
80 DIF4=0.0
    DIF5=0.0
60 Y=DIF1**2.+DIF2**2.+DIF3**2.+DIF4**2.+DIF5**2.
RETURN
END
SUBROUTINE TITLE(NAME1,NAME2,NAME3,ALFA,DG23,DG32,GPN2, 
% GPN3,ZPN2,ZPN3,B012,B013,B112,B113,B123,DELTA,KP,NTYPE, 
% XX2,XHDL,NNRTL,DGA2,DGA3,DGB2,DGB3,ALFA2,ALFB2,ALFA3, 
% ALFB3)

*******************************************************************************

SUBROUTINE TITLE

THIS SUBROUTINE MAKES TITLES FOR FINAL TABULAR RESULTS

*******************************************************************************

REAL *8 NAME1,NAME2,NAME3
WRITE(6,7100)

7100 FORMAT('1',30X,'TABLE#')
WRITE(6,7110)NAME1,NAME2,NAME3

7110 FORMAT('/','15X','SYSTEM:','3A8)
WRITE(6,7130)

7130 FORMAT('/','25X','** VALUES OF THE PARAMETERS **')
GO TO(9041,9041,9061,9061,9061,9061,9061,9061,9061,9061,9061,9061,9061)

IF(KP,GT,1)GO TO 9065
WRITE(6,120)

IF(XX2.EQ.0.0)GO TO 9070
WRITE(6,9082)B012,B112

9082 FORMAT('/*10X','B012=','G12.5','B112=','G12.5)
GO TO 9052

9070 WRITE(6,9081)B013,B113

9081 FORMAT('/*10X','B013=','G12.5','B113=','G12.5)
DO TO 110

110 FORMAT('/*10X','EXTENDED DEBYE-HUCKEL + MODIFIED NRTL')

120 FORMAT('/*10X','BROMLEY - BINARY - EXPRESSION')

130 FORMAT('/*5X','BROMLEY & SIMPLIFIED MODIFIED NRTL &',
% ' SALTING-OUT EQUATIONS')

DO TO 9052

9065 WRITE(6,110)
IF(XX2.EQ.0.0)GO TO 9075
GO TO(10,20),NNRTL

20 WRITE(6,9084)GPN2,ZPN2

9084 FORMAT('/*10X','GPN2=','G12.5',3X,'ZPN2=','G12.5)
GO TO 9052

10 WRITE(6,100)ALFA2,ALFB2,DGA2,DGB2

100 FORMAT('/*3X','ALFA2=','G12.5','ALFB2=','G12.5','DGA2=',
% 'G12.5','DGB2=','G12.5)
GO TO 9052

9075 GO TO(30,40),NNRTL

30 WRITE(6,150)ALFA3,ALFB3,DGA3,DGB3
?
150 FORMAT(/,3X,'ALFA3=' ,G12.5,'ALFB3=' ,G12.5,'DGA3=' ,G12.5,'DBG3=',G12.5)
   GO TO 9052
40 WRITE(6,9086)GPN3,ZN3
9086 FORMAT(/,10X,'GPN3=' ,G12.5,3X,'ZPN3=' ,G12.5)
   GO TO 9052
9061 WRITE(6,7140)ALFA,DBG23,DBG32
7140 FORMAT(/,3X,'NONELECTROLYTE BINARY ','ALFA=' ,G12.5,
   'DG23=' ,G12.5,'DG32=' ,G12.5)
   GO TO 9052
   IF(XMOL.EQ.0.0)GO TO 9052
   IF(KP.GT.1)GO TO 9051
WRITE(6,130)
WRITE(6,7150)B012,B112,B013,B113
7150 FORMAT(/,5X,'B012=' ,G10.3,'B112=' ,G10.3,'B013=' ,G10.3,
   'B113=' ,G10.3)
   WRITE(6,200)B123,DELTA
200 FORMAT(/,15X,'B123=' ,G12.5,'DELTA '=' ,G12.5)
   GO TO 9052
9051 WRITE(6,110)
   GO TO(70,80),NNRTL
70 WRITE(6,100)ALFA2,ALFB2,DGA2,DBG2
WRITE(6,150)ALFA3,ALFB3,DGA3,DBG3
   GO TO 9052
80 WRITE(6,7160)GPN2,ZPN2,GPN3,ZPN3
7160 FORMAT(/,6X,'GPN2=' ,G12.5,'ZPN2=' ,G12.5,'GPN3=' ,G12.5,
   'ZPN3=' ,G12.5)
   GO TO 9052
9052 RETURN
END
### INPUT DATA SEQUENCE

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<td>INFORMATION</td>
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<td>XLIM,XLLIM</td>
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<td>NAME1,NAMES2,NAMES3 (name of the system)</td>
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<td>NNRTL,NBROM,NTYPE,NMIN</td>
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<td>(pure component vapor pressure constants)</td>
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<td>of data are nxd) (if NXD = 11, card #31-41)</td>
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<td>card #43-55)</td>
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94  M,MM  212

95-(94+MM) XTX(I),BXX(I) (two data on a card, 2F10.4
total cards = MM) (if MM=4 cards 95-98)
### SAMPLE INPUT

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<td>8.</td>
<td>* BINARY 1-3: SKABICHVESKII</td>
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195

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SAMPLE OUTPUT

* OF DATA SET TO BE USED = 1TRIAL #400E=0.00000100

FILE NAME = LICLH20,NEOHAT25,COMBINED

BINARY 1-3: SKABICHVESKKI

TERNARY 1-2-3: CIPARIS


INPUT DATA

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18.06000 273.12988 18.27800 323.14990 18.84399 373.14990
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LIQUID MOLAR VOLUME CONSTANTS

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64.510 -0.19716 0.38735E-03

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BASED ON # OF POINTS FOR Y = 5

AVERAGE DP(PCAL-PEXP) = 6.8366
BASED ON # OF POINTS FOR DP = 5
APPENDIX G

TABLES AND FIGURES FOR MODEL I
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### TABLE G.2 Ternary Data Sources

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<td>CaCl\textsubscript{2}-MeOH</td>
<td>7</td>
<td>3.0</td>
<td>25</td>
<td>5.2187</td>
<td>537.62</td>
<td>-</td>
</tr>
<tr>
<td>HCl-EtOH</td>
<td>8</td>
<td>0.1</td>
<td>25</td>
<td>117.0</td>
<td>12.35</td>
<td>8.7</td>
</tr>
<tr>
<td>HCl-MeOH</td>
<td>22</td>
<td>0.56</td>
<td>25</td>
<td>10.8</td>
<td>39.7</td>
<td>0.7</td>
</tr>
<tr>
<td>LiCl-EtOH</td>
<td>4</td>
<td>4.0</td>
<td>25</td>
<td>3.67</td>
<td>151.5</td>
<td>-</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>6</td>
<td>4.0</td>
<td>25</td>
<td>5.2</td>
<td>60.0</td>
<td>-</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>11</td>
<td>6.0</td>
<td>60</td>
<td>3.403</td>
<td>106.21</td>
<td>-</td>
</tr>
<tr>
<td>NaBr-MeOH</td>
<td>9</td>
<td>1.6</td>
<td>25</td>
<td>6.449</td>
<td>54.747</td>
<td>-</td>
</tr>
<tr>
<td>NaCl-MeOH</td>
<td>7</td>
<td>0.1</td>
<td>25</td>
<td>380.8</td>
<td>4.26</td>
<td>2.7</td>
</tr>
</tbody>
</table>

TABLE G.5 Nonaqueous Electrolytic Binary Data with Two Objective Functions, Equations (2-9) and (2-10)
TABLE G.6  Aqueous Electrolytic Binary Data Correlation with Temperature Independent Parameters Using Objective Function #2, Equation (2-10) and Presetting

\[ \alpha_{A2} = 0.2; \alpha_{B2} = 0.0 \]

<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max 'm'</th>
<th>T (°C)</th>
<th>( \Delta g_{A2} )</th>
<th>( \Delta g_{B2} )</th>
<th>% Error in ( \gamma^+ )</th>
<th>% Error in DP</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td>Max</td>
<td>Avg</td>
<td>Max</td>
<td>Max</td>
<td>Avg</td>
</tr>
<tr>
<td>( \text{CaCl}_2-\text{H}_2\text{O} )</td>
<td>17</td>
<td>3.0</td>
<td>25</td>
<td>-42.589</td>
<td>668.8</td>
<td>7.1</td>
<td>4.2</td>
</tr>
<tr>
<td>( \text{HCl-} \text{H}_2\text{O}^* )</td>
<td>15</td>
<td>2.0</td>
<td>25</td>
<td>77.553</td>
<td>-3.3056</td>
<td>5.5</td>
<td>2.4</td>
</tr>
<tr>
<td>( \text{LiCl-} \text{H}_2\text{O} )</td>
<td>19</td>
<td>4.0</td>
<td>25</td>
<td>61.958</td>
<td>-5.0817</td>
<td>15.8</td>
<td>6.6</td>
</tr>
<tr>
<td>( \text{LiCl-} \text{H}_2\text{O} )</td>
<td>11</td>
<td>6.0</td>
<td>60</td>
<td>81.743</td>
<td>-11.592</td>
<td>-</td>
<td>-</td>
</tr>
<tr>
<td>( \text{NaBr-} \text{H}_2\text{O} )</td>
<td>4</td>
<td>10.3</td>
<td>40</td>
<td>98.318</td>
<td>-3.2249</td>
<td>-</td>
<td>-</td>
</tr>
<tr>
<td>( \text{NaBr-} \text{H}_2\text{O} )</td>
<td>19</td>
<td>4.0</td>
<td>25</td>
<td>-24.258</td>
<td>174.11</td>
<td>9.4</td>
<td>4.0</td>
</tr>
<tr>
<td>( \text{NaCl-} \text{H}_2\text{O} )</td>
<td>17</td>
<td>4.0</td>
<td>25</td>
<td>-32.396</td>
<td>444.79</td>
<td>3.4</td>
<td>1.2</td>
</tr>
</tbody>
</table>

*For the system HCl-H2O binary objective function #1 [Equation (2-9)] is used.
<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max 'm'</th>
<th>T (°C)</th>
<th>(\Delta G_{A3} )</th>
<th>(\Delta G_{B3} )</th>
<th>% Error in (\gamma_{\pm} )</th>
<th>% Error in DP</th>
</tr>
</thead>
<tbody>
<tr>
<td>CaCl₂-MeOH</td>
<td>7</td>
<td>3.0</td>
<td>25</td>
<td>-16.156</td>
<td>289.16</td>
<td>-</td>
<td>16.3</td>
</tr>
<tr>
<td>HCl-EtOH</td>
<td>8</td>
<td>0.1</td>
<td>25</td>
<td>-58.994</td>
<td>6897.7</td>
<td>8.7</td>
<td>6.0</td>
</tr>
<tr>
<td>HCl-MeOH</td>
<td>22</td>
<td>0.56</td>
<td>25</td>
<td>-28.306</td>
<td>317.33</td>
<td>0.66</td>
<td>0.4</td>
</tr>
<tr>
<td>LiCl-EtOH</td>
<td>3</td>
<td>4.0</td>
<td>25</td>
<td>-14.056</td>
<td>187.82</td>
<td>-</td>
<td>15.3</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>6</td>
<td>4.0</td>
<td>25</td>
<td>-20.384</td>
<td>161.93</td>
<td>-</td>
<td>7.0</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>11</td>
<td>6.0</td>
<td>60</td>
<td>-19.497</td>
<td>170.92</td>
<td>-</td>
<td>17.5</td>
</tr>
<tr>
<td>NaBr-MeOH</td>
<td>9</td>
<td>1.6</td>
<td>25</td>
<td>-21.556</td>
<td>177.28</td>
<td>-</td>
<td>4.1</td>
</tr>
<tr>
<td>NaCl-MeOH</td>
<td>7</td>
<td>0.1</td>
<td>25</td>
<td>-10.262</td>
<td>210.86</td>
<td>11.3</td>
<td>6.6</td>
</tr>
<tr>
<td>LiBr-MeOH</td>
<td>8</td>
<td>6.64</td>
<td>15</td>
<td>-20.198</td>
<td>158.28</td>
<td>-</td>
<td>2.9</td>
</tr>
<tr>
<td>LiBr-MeOH</td>
<td>8</td>
<td>6.64</td>
<td>30</td>
<td>-23.999</td>
<td>216.56</td>
<td>-</td>
<td>14.0</td>
</tr>
</tbody>
</table>
TABLE G.8  Binary Data Correlation with Temperature Independent Parameters Using Objective Function #2, Equation (2-10) and Presetting $\alpha_{Ai} = 0.2; \beta_{Bi} = -1.0$

<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max $\Delta g_{Ai}$</th>
<th>Max $\Delta g_{Bi}$</th>
<th>Max % Error in $\gamma^\pm$</th>
<th>% Error in DP</th>
<th>Avg</th>
</tr>
</thead>
<tbody>
<tr>
<td>LiCl-H$_2$O</td>
<td>19</td>
<td>4.0</td>
<td>-2.8668</td>
<td>-12.27</td>
<td>6.6</td>
<td>3.4</td>
</tr>
<tr>
<td>NaBr-H$_2$O</td>
<td>4</td>
<td>10.3</td>
<td>-2.1801</td>
<td>-13.616</td>
<td>-</td>
<td>5.0</td>
</tr>
<tr>
<td>LiCl-EtOH</td>
<td>2</td>
<td>1.0</td>
<td>20.872</td>
<td>6.102</td>
<td>-</td>
<td>22.5</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>6</td>
<td>3.7</td>
<td>20.285</td>
<td>5.32</td>
<td>-</td>
<td>9.7</td>
</tr>
<tr>
<td>NaBr-MeOH</td>
<td>9</td>
<td>1.6</td>
<td>19.938</td>
<td>5.2467</td>
<td>-</td>
<td>5.7</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>11</td>
<td>6.0</td>
<td>23.133</td>
<td>6.3664</td>
<td>-</td>
<td>27.4</td>
</tr>
</tbody>
</table>

Note: The table contains data for various systems, showing the number of points, maximum change in excess Gibbs free energy ($\Delta g_{Ai}$ and $\Delta g_{Bi}$), maximum percentage error in $\gamma^\pm$, and percentage error in the distribution parameters (DP).
<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max T (°C)</th>
<th>Tm, max</th>
<th>Δg_Ai</th>
<th>Δg_Bi</th>
<th>% Error in DP Max</th>
<th>Avg</th>
</tr>
</thead>
<tbody>
<tr>
<td>LiCl-EtOH</td>
<td>2</td>
<td>1.0</td>
<td>6.1822</td>
<td>5.454</td>
<td>-27.801</td>
<td>27.5</td>
<td>25.3</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>6</td>
<td>3.7</td>
<td>5.4454</td>
<td>-32.375</td>
<td>-32.753</td>
<td>17.7</td>
<td>10.5</td>
</tr>
<tr>
<td>NaBr-MeOH</td>
<td>9</td>
<td>1.6</td>
<td>5.3709</td>
<td>-51.103</td>
<td>-51.103</td>
<td>7.3</td>
<td>3.9</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>11</td>
<td>6.0</td>
<td>6.479</td>
<td>-50.83</td>
<td>7.1956</td>
<td>43.6</td>
<td>28.3</td>
</tr>
<tr>
<td>CaCl_2-MeOH</td>
<td>7</td>
<td>2.6</td>
<td>6.479</td>
<td>-50.83</td>
<td>7.1956</td>
<td>275.7</td>
<td>205.0</td>
</tr>
</tbody>
</table>
TABLE G.10  Prediction of Binary Data at One Temperature Using the Parameters of
Another Temperature, Presetting $\alpha_{Ai} = 0.2$ and $\alpha_{Bi} = 0.0$

<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max 'm'</th>
<th>$T$ (°C)</th>
<th>$\Delta g_{Ai}$</th>
<th>$\Delta g_{Bi}$</th>
<th>% Error in $\gamma_+$</th>
<th>% Error in DP</th>
</tr>
</thead>
<tbody>
<tr>
<td>LiCl-H$_2$O</td>
<td>19</td>
<td>6.0</td>
<td>25</td>
<td>60</td>
<td>81.743</td>
<td>-11.592</td>
<td>24.3</td>
</tr>
<tr>
<td>LiCl-H$_2$O</td>
<td>11</td>
<td>6.0</td>
<td>60</td>
<td>25</td>
<td>61.958</td>
<td>-5.0817</td>
<td>-</td>
</tr>
<tr>
<td>NaBr-H$_2$O</td>
<td>19</td>
<td>4.0</td>
<td>25</td>
<td>40</td>
<td>98.318</td>
<td>-3.2249</td>
<td>33.5</td>
</tr>
<tr>
<td>NaBr-H$_2$O</td>
<td>4</td>
<td>10.3</td>
<td>40</td>
<td>25</td>
<td>163.29</td>
<td>12.545</td>
<td>-</td>
</tr>
<tr>
<td>LiBr-MeOH</td>
<td>8</td>
<td>6.6</td>
<td>15</td>
<td>30</td>
<td>-23.999</td>
<td>216.56</td>
<td>-</td>
</tr>
<tr>
<td>LiBr-MeOH</td>
<td>8</td>
<td>6.6</td>
<td>30</td>
<td>15</td>
<td>-20.198</td>
<td>158.28</td>
<td>-</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>9</td>
<td>6.0</td>
<td>25</td>
<td>60</td>
<td>-19.497</td>
<td>170.9</td>
<td>-</td>
</tr>
<tr>
<td>LiCl-MeOH</td>
<td>11</td>
<td>6.0</td>
<td>60</td>
<td>25</td>
<td>-20.384</td>
<td>161.8</td>
<td>-</td>
</tr>
</tbody>
</table>
TABLE G.11 Isothermal Ternary $\gamma_\pm$ Data Correlation for the Four Parameters: $G_{\pm 2}$, $Z_{\pm 2}$, $G_{\pm 3}$ and $Z_{\pm 3}$

Presetting the Solvent-Solvent Binary Parameters Corresponding to $\alpha_{23} = -1.0$

<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max 'm'</th>
<th>T (°C)</th>
<th>$\alpha_{23}$</th>
<th>$\Delta g_{23}$</th>
<th>$\Delta g_{32}$</th>
<th>$G_{\pm 2}$</th>
<th>$Z_{\pm 2}$</th>
<th>$G_{\pm 3}$</th>
<th>$Z_{\pm 3}$</th>
<th>% Error in $\gamma_{\pm}$</th>
<th>Max</th>
<th>Avg</th>
</tr>
</thead>
<tbody>
<tr>
<td>HCl-H$_2$O-EtOH</td>
<td>44</td>
<td>2.5</td>
<td>25</td>
<td>-1.0</td>
<td>105.8</td>
<td>383.8</td>
<td>19.677</td>
<td>1.789</td>
<td>9.460</td>
<td>84.647</td>
<td>10.0</td>
<td>2.0</td>
<td></td>
</tr>
<tr>
<td>HCl-H$_2$O-MeOH</td>
<td>48</td>
<td>2.0</td>
<td>25</td>
<td>-1.0</td>
<td>-150.9</td>
<td>336.5</td>
<td>19.677</td>
<td>1.789</td>
<td>7.63</td>
<td>40.14</td>
<td>18.4</td>
<td>2.2</td>
<td></td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>45</td>
<td>1.0</td>
<td>25</td>
<td>-1.0</td>
<td>-150.9</td>
<td>336.5</td>
<td>28.32</td>
<td>3.141</td>
<td>21.875</td>
<td>24.882</td>
<td>22.8</td>
<td>7.7</td>
<td></td>
</tr>
<tr>
<td>NaCl-H$_2$O-MeOH</td>
<td>35</td>
<td>1.0</td>
<td>25</td>
<td>-1.0</td>
<td>-150.9</td>
<td>336.5</td>
<td>38.981</td>
<td>4.216</td>
<td>44.11</td>
<td>14.638</td>
<td>27.0</td>
<td>6.1</td>
<td></td>
</tr>
</tbody>
</table>
### TABLE G.12 A Comparative Study of the Three Objective Functions [Equations (2-19), (2-20), (2-21)] in Correlating Ternary Isothermal VLE Data with the Four Parameters $[G_{\pm 2}, Z_{\pm 2}, G_{\pm 3}, Z_{\pm 3}]$, Presetting $\Delta g_{23}$ and $\Delta g_{32}$ from Table G.3 Corresponding to $a_{23} = -1.0$

<table>
<thead>
<tr>
<th>System</th>
<th>Max '$m'$</th>
<th>T (°C)</th>
<th>$\Delta Y$ Max</th>
<th>$\Delta Y$ Avg</th>
<th>$\Delta P$ (mmHg) Max</th>
<th>$\Delta P$ (mmHg) Avg</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>Objective Function #1</strong></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H$_2$O-EtOH</td>
<td>1.0</td>
<td>25</td>
<td>0.035</td>
<td>0.011</td>
<td>5.6</td>
<td>2.2</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>1.0</td>
<td>25</td>
<td>0.022</td>
<td>0.012</td>
<td>8.0</td>
<td>4.8</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>6.2</td>
<td>40</td>
<td>0.023</td>
<td>0.012</td>
<td>11.5</td>
<td>5.1</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>7.1</td>
<td>25</td>
<td>0.047</td>
<td>0.017</td>
<td>15.0</td>
<td>8.8</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>6.0</td>
<td>60</td>
<td>0.04</td>
<td>0.015</td>
<td>33.2</td>
<td>13.8</td>
</tr>
<tr>
<td><strong>Objective Function #2</strong></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H$_2$O-EtOH</td>
<td>1.0</td>
<td>25</td>
<td>0.034</td>
<td>0.009</td>
<td>3.1</td>
<td>1.2</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>1.0</td>
<td>25</td>
<td>0.009</td>
<td>0.0055</td>
<td>9.3</td>
<td>6.8</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>6.2</td>
<td>40</td>
<td>0.022</td>
<td>0.01</td>
<td>17.7</td>
<td>5.6</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>7.1</td>
<td>25</td>
<td>0.02</td>
<td>0.0088</td>
<td>11.4</td>
<td>4.8</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>6.0</td>
<td>60</td>
<td>0.023</td>
<td>0.009</td>
<td>54.3</td>
<td>11.7</td>
</tr>
<tr>
<td><strong>Objective Function #3</strong></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H$_2$O-EtOH</td>
<td>1.0</td>
<td>25</td>
<td>0.035</td>
<td>0.01</td>
<td>6.1</td>
<td>2.3</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>1.0</td>
<td>25</td>
<td>0.011</td>
<td>0.0066</td>
<td>10.2</td>
<td>5.6</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>6.2</td>
<td>40</td>
<td>0.022</td>
<td>0.011</td>
<td>15.2</td>
<td>5.4</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>7.1</td>
<td>25</td>
<td>0.021</td>
<td>0.0092</td>
<td>11.3</td>
<td>5.0</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>6.0</td>
<td>60</td>
<td>0.025</td>
<td>0.0097</td>
<td>50.3</td>
<td>11.6</td>
</tr>
<tr>
<td>System</td>
<td>T (°C)</td>
<td>$G_{\pm 2}$</td>
<td>$Z_{\pm 2}$</td>
<td>$G_{\pm 3}$</td>
<td>$Z_{\pm 3}$</td>
<td></td>
</tr>
<tr>
<td>---------------</td>
<td>-------</td>
<td>------------</td>
<td>------------</td>
<td>------------</td>
<td>------------</td>
<td></td>
</tr>
<tr>
<td>LiCl-H$_2$O-EtOH</td>
<td>25</td>
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TABLE G.14  A Comparative Study of the Three Objective Functions [Equations (2-19), (2-20), (2-21)] in Correlating Ternary Isobaric VLE Data with the Four Parameters [$\Delta g_{A2}'$, $\Delta g_{B2}'$, $\Delta g_{A3}$ & $\Delta g_{B3}$]

Presetting $\Delta g_{23}$ and $\Delta g_{32}$ Corresponding to $a_{23} = -1.0$ [Table G.3] and $a_{A2} = 0.2$; $a_{B2} = 0.0$; $a_{A3} = 0.2$; $a_{B3} = 0.0$

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<th>Max</th>
<th>Avg</th>
<th>$\Delta P$ (mmHg)</th>
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<td>( \Delta g_{B2} )</td>
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Figure G.1 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System HCl-H₂O-EtOH at 25°C and Constant $X_{EtOH} = 0.0417$.

- Experimental, Harned and Owen (1958); ○ Correlated ($\alpha_{23} = -1.0$);
- △ Predicted ($\alpha_{23} = -1.0$)
Figure G.1A Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I up to \( m = 0.2 \) for the System HCl-H₂O-EtOH at 25°C and Constant \( X'_{\text{EtOH}} = 0.0417 \).
Figure G.2 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System HCl-H2O-EtOH at 25°C and Constant $X_{\text{EtOH}} = 0.0891$
Figure G.2A Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I up to $m = 0.2$ for the System HCl-H$_2$O-EtOH at 25°C and Constant $X'_{\text{EtOH}} = 0.0891$
Figure G.3 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System HCl-H₂O-EtOH at 25°C and Constant $X'_{EtOH} = 0.5$
Figure G.3A Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I up to $m = 0.2$ for the System HCl-H$_2$O-EtOH at 25°C and Constant $x'_{\text{EtOH}} = 0.5$.
Figure G.4 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System HCl-H₂O-MeOH at 25°C and Constant $\gamma_{\text{MeOH}} = 0.0584$
Figure G.5 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System HCl-H$_2$O-MeOH at 25°C and Constant $X^*$$_{MeOH}$ = 0.1233.
Figure G.6 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System LiCl-H2O-MeOH at 25°C.
Figure G.7 Comparison of Experimental Mean Molal Activity Coefficients with those Predicted and Correlated by Model I for the System LiCl-H₂O-MeOH at 25°C and Constant m = 0.5, 1.0. Experimental, Akerlof (1930); ■ Correlated (α₂₃ = -1.0); △ Predicted (α₂₃ = -1.0); □ Predicted (α₂₃ = 0.3)
Figure G.8 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated for the System NaCl – H₂O – MeOH at 25°C and Constant m = 0.02, 0.05

---

Experimental, Akerlof (1930); O Correlated (α₂₃=-1.0)
Figure G.9 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated for the System NaCl-H$_2$O-MeOH at $25^\circ$C and Constant $m = 0.2, 0.5$.

--- Experimental, Akerlof (1930); ○ Correlated ($\alpha_{23} = -1.0$)
Figure G.10 Comparison of Experimental Vapor Phase Compositions with those Predicted and Correlated by Model I for the System LiCl-H\textsubscript{2}O-EtOH at 25°C and Constant m = 0.5

- $\bigcirc$ Correlated ($\alpha_{23} = -1.0$)
- $\square$ Predicted ($\alpha_{23} = 1.0$)
- $\triangle$ Predicted ($\alpha_{23} = 0.3$)
Figure G.11 Comparison of Experimental with Predicted and Correlated Vapor-Phase Compositions Using Model I for the System LiCl-H₂O-EtOH at 25°C and Constant m = 1.0

- ○ Correlated (α_{23} = -1.0)
- □ Predicted (α_{23} = -1.0)
- △ Predicted (α_{23} = 0.3)
Figure G.12 Comparison of Experimental with Predicted and Correlated Vapor-Phase Compositions Using Model I for the System LiCl-H$_2$O-MeOH at 25°C and Constant m = 1.0
Figure G.13 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System NaBr-H₂O-MeOH at 40°C
Figure G.14 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System NaBr-H$_2$O-MeOH at 25°C
Figure G.15 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System LiCl-H₂O at 60°C
Figure G.16 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System LiCl-H$_2$O-MeOH at P = 1 atm
Figure G.17 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System NaBr-H$_2$O-MeOH at P = 1 atm

\[ \alpha_{23} = -1.0 \]

\[ m = 0.04 - 0.92 \]
\[ m = 0.08 - 1.8 \]
\[ m = 0.18 - 3.8 \]
Figure G.18 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System KCl-H$_2$O-MeOH at P = 1 atm
Figure G.19 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model I for the System NaF-H$_2$O-MeOH at P = 1 atm
Figure G.20 Contribution of the NRTL Term to $\ln \gamma_i$ for the System LiCl-H$_2$O-EtOH at 25°C in Model I

Constant $m = 1.0$

- △ $\ln \gamma_{H_2O}$ NRTL
- ○ $\ln \gamma_{EtOH}$ NRTL

$\ln \gamma_i$ NRTL

$x'_{EtOH}$
APPENDIX H

TABLES AND FIGURES FOR MODEL II
### TABLE H.1 Typical m-\( \bar{D}P \) Data from the Weast Compilation

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<td>3.0</td>
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<td>12.5</td>
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TABLE II (Cont'd.)

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<th>max m</th>
<th>DP</th>
<th>D$P_{\text{max}}$</th>
<th>D$P'_{\text{max}}$</th>
<th>B$_{12}$(100°C)</th>
<th>B$_{12}$(25°C)</th>
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<td>1.6</td>
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<td>-0.0358</td>
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<td>$P_{\text{max}}$ (100°C)</td>
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</tr>
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<td>----------------------</td>
<td>--------------------------</td>
<td>---------------------------</td>
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<td>Pb(NO$_3$)$_2$</td>
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<td>ZnCl$_2$</td>
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<td>Electrolyte</td>
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<td>$D_P$</td>
<td>$D_P^{\text{max}}$</td>
<td>$D_P^{'\text{max}}$</td>
<td>$B_{12}(100^\circ C)$</td>
<td>$B_{12}(25^\circ C)$</td>
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<td>-------</td>
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<td>0.0037</td>
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*Data for KBr are from Robinson and Stokes (1955) because they are not reported by Weast.*
### TABLE H.3 Quality of Results with the $B_{12}$ Value at 70°C Obtained by Interpolation of the $B_{12}(25°C)$ and $B_{12}(100°C)$ Values in Equation (3-20)

<table>
<thead>
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<th>Electrolyte</th>
<th>$B_{12}(70°C)$</th>
<th>Abs. % Error in $\gamma_\pm$</th>
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<td>Avg</td>
<td>Max</td>
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<td>0.3</td>
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</tr>
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<td>6.3</td>
<td>0.7</td>
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<tr>
<td>System</td>
<td># of Points</td>
<td>Max T (°C)</td>
<td>B13</td>
<td>% Error in $Y^+$</td>
</tr>
<tr>
<td>----------------</td>
<td>-------------</td>
<td>------------</td>
<td>-----</td>
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</tr>
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<td>2.6</td>
<td>25</td>
<td>0.15541</td>
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<tr>
<td>HCl-EtOH</td>
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<td>25</td>
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<td>0.41205</td>
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<td>60</td>
<td>0.25195</td>
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<td>0.1</td>
<td>25</td>
<td>-0.3223</td>
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TABLE H.5  Aqueous Electrolytic Binary Data Correlation with the Bromley Equation

<table>
<thead>
<tr>
<th>System</th>
<th># of Points</th>
<th>Max 'm'</th>
<th>T (°C)</th>
<th>B_{12}</th>
<th>% Error in $\gamma_+^+$ Max/Avg</th>
<th>% Error in DP Max/Avg</th>
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<tr>
<td>CaCl$_2$-H$_2$O</td>
<td>21</td>
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<td>25</td>
<td>0.1000</td>
<td>11.0/5.5</td>
<td>3.6/2.10</td>
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<td>2.0</td>
<td>25</td>
<td>0.13963</td>
<td>0.9/0.6</td>
<td>-/ -</td>
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<td>19</td>
<td>4.0</td>
<td>25</td>
<td>0.12366</td>
<td>2.5/0.9</td>
<td>2.0/0.5</td>
</tr>
<tr>
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<td>6.0</td>
<td>60</td>
<td>0.12049</td>
<td>-/-</td>
<td>1.9/0.9</td>
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<td>10.3</td>
<td>40</td>
<td>0.06607</td>
<td>-/-</td>
<td>13.0/5.5</td>
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<td>19</td>
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<td>25</td>
<td>0.07376</td>
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<td>1.1/0.2</td>
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<td>25</td>
<td>0.05586</td>
<td>0.1/0.04</td>
<td>1.0/0.3</td>
</tr>
<tr>
<td>System</td>
<td># of Points</td>
<td>Max 'm'</td>
<td>T (°C)</td>
<td>Ternary Parameters</td>
<td>% Error in $\gamma_\pm$</td>
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<td>-------------</td>
<td>---------</td>
<td>--------</td>
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<td>HCl-H$_2$O-EtOH</td>
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<td>25</td>
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<td>-0.0245</td>
<td>13.8 2.4</td>
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<tr>
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TABLE H.7 A Comparative Study of the Three Objective Functions [Equations (2-19), (2-20), (2-21)] in Correlating Ternary Isothermal VLE Data with Model II

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<tr>
<th>System</th>
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<th>$\Delta Y$</th>
<th>$\Delta P$ (mmHg)</th>
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<td>Avg</td>
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<td><strong>Objective Function #1</strong></td>
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<td></td>
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<tr>
<td>LiCl-H$_2$O-EtOH</td>
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<td>25</td>
<td>0.034</td>
<td>0.0145</td>
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<td>0.026</td>
<td>0.014</td>
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<td>0.028</td>
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<td>7.1</td>
<td>25</td>
<td>0.042</td>
<td>0.0096</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>2.0</td>
<td>60</td>
<td>0.038</td>
<td>0.018</td>
</tr>
<tr>
<td></td>
<td>3.0</td>
<td>60</td>
<td>0.036</td>
<td>0.02</td>
</tr>
<tr>
<td><strong>Objective Function #3</strong></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H$_2$O-EtOH</td>
<td>1.0</td>
<td>25</td>
<td>0.025</td>
<td>0.01</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>1.0</td>
<td>25</td>
<td>0.011</td>
<td>0.008</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>3.0</td>
<td>40</td>
<td>0.02</td>
<td>0.012</td>
</tr>
<tr>
<td></td>
<td>6.2</td>
<td>40</td>
<td>0.02</td>
<td>0.011</td>
</tr>
<tr>
<td>NaBr-H$_2$O-MeOH</td>
<td>3.0</td>
<td>25</td>
<td>0.014</td>
<td>0.007</td>
</tr>
<tr>
<td></td>
<td>7.1</td>
<td>25</td>
<td>0.048</td>
<td>0.011</td>
</tr>
<tr>
<td>LiCl-H$_2$O-MeOH</td>
<td>2.0</td>
<td>60</td>
<td>0.042</td>
<td>0.019</td>
</tr>
<tr>
<td></td>
<td>3.0</td>
<td>60</td>
<td>0.036</td>
<td>0.02</td>
</tr>
</tbody>
</table>
**TABLE H.8** Values of the Parameters Obtained with Three Objective Functions for Model II

<table>
<thead>
<tr>
<th>System</th>
<th>Max 'm'</th>
<th>T (°C)</th>
<th>Ternary Parameters</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td>B_{123}   δ_{123}</td>
</tr>
<tr>
<td><strong>Objective Function #1</strong></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-EtOH</td>
<td>1.0</td>
<td>25</td>
<td>-117.68</td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-MeOH</td>
<td>1.0</td>
<td>25</td>
<td>-145.85</td>
</tr>
<tr>
<td>NaBr-H\textsubscript{2}O-MeOH</td>
<td>3.0</td>
<td>40</td>
<td>-60.53</td>
</tr>
<tr>
<td></td>
<td>6.2</td>
<td>40</td>
<td>-68.105</td>
</tr>
<tr>
<td>NaBr-H\textsubscript{2}O-MeOH</td>
<td>3.0</td>
<td>25</td>
<td>-93.462</td>
</tr>
<tr>
<td></td>
<td>7.1</td>
<td>25</td>
<td>-81.45</td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-MeOH</td>
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<td>60</td>
<td>-21.409</td>
</tr>
<tr>
<td></td>
<td>3.0</td>
<td>60</td>
<td>-32.164</td>
</tr>
<tr>
<td><strong>Objective Function #2</strong></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-EtOH</td>
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<td>25</td>
<td>-152.94</td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-MeOH</td>
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<td>25</td>
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<tr>
<td>NaBr-H\textsubscript{2}O-MeOH</td>
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<td>40</td>
<td>-59.76</td>
</tr>
<tr>
<td></td>
<td>6.2</td>
<td>40</td>
<td>-64.96</td>
</tr>
<tr>
<td>NaBr-H\textsubscript{2}O-MeOH</td>
<td>3.0</td>
<td>25</td>
<td>-93.2</td>
</tr>
<tr>
<td></td>
<td>7.1</td>
<td>25</td>
<td>-99.38</td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-MeOH</td>
<td>2.0</td>
<td>60</td>
<td>-72.477</td>
</tr>
<tr>
<td></td>
<td>3.0</td>
<td>60</td>
<td>-80.63</td>
</tr>
<tr>
<td><strong>Objective Function #3</strong></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-EtOH</td>
<td>1.0</td>
<td>25</td>
<td>-145.91</td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-MeOH</td>
<td>1.0</td>
<td>25</td>
<td>-133.37</td>
</tr>
<tr>
<td>NaBr-H\textsubscript{2}O-MeOH</td>
<td>3.0</td>
<td>40</td>
<td>-51.02</td>
</tr>
<tr>
<td></td>
<td>6.2</td>
<td>40</td>
<td>-62.74</td>
</tr>
<tr>
<td>NaBr-H\textsubscript{2}O-MeOH</td>
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<td>25</td>
<td>-93.997</td>
</tr>
<tr>
<td></td>
<td>7.1</td>
<td>25</td>
<td>-95.00</td>
</tr>
<tr>
<td>LiCl-H\textsubscript{2}O-MeOH</td>
<td>2.0</td>
<td>60</td>
<td>-59.4</td>
</tr>
<tr>
<td></td>
<td>3.0</td>
<td>60</td>
<td>-68.70</td>
</tr>
</tbody>
</table>
TABLE H.9  \( B_{12} \) Values Obtained from Weast's Data and Equation (3-20)  
with \( B^* \) and \( B'_{12} \) Values from Table 3.3

<table>
<thead>
<tr>
<th>Electrolyte</th>
<th>( B_{12} ) from Weast Data</th>
<th>( B_{12} ) from Eqn (3-20)</th>
</tr>
</thead>
<tbody>
<tr>
<td>NaCl</td>
<td>0.0567</td>
<td>0.0612</td>
</tr>
<tr>
<td>KBr</td>
<td>0.0462</td>
<td>0.0436</td>
</tr>
<tr>
<td>KCl</td>
<td>0.0311</td>
<td>0.0322</td>
</tr>
<tr>
<td>Na(_2)SO(_4)</td>
<td>-0.0086</td>
<td>-0.0036</td>
</tr>
<tr>
<td>MgCl(_2)</td>
<td>0.0951</td>
<td>0.1017</td>
</tr>
<tr>
<td>MgSO(_4)</td>
<td>0.0107</td>
<td>0.0163</td>
</tr>
<tr>
<td>Molality</td>
<td>$\gamma_\pm$ from Weast Data</td>
<td>$\gamma_\pm$ from Eqn (3-20)</td>
</tr>
<tr>
<td>----------</td>
<td>-----------------------------</td>
<td>-----------------------------</td>
</tr>
<tr>
<td>0.2</td>
<td>0.4034</td>
<td>0.4071</td>
</tr>
<tr>
<td>0.4</td>
<td>0.3737</td>
<td>0.3796</td>
</tr>
<tr>
<td>0.6</td>
<td>0.3713</td>
<td>0.3792</td>
</tr>
<tr>
<td>0.8</td>
<td>0.3820</td>
<td>0.3924</td>
</tr>
<tr>
<td>1.0</td>
<td>0.4018</td>
<td>0.4150</td>
</tr>
<tr>
<td>1.2</td>
<td>0.4292</td>
<td>0.4458</td>
</tr>
<tr>
<td>1.4</td>
<td>0.4636</td>
<td>0.4842</td>
</tr>
<tr>
<td>1.6</td>
<td>0.5050</td>
<td>0.5304</td>
</tr>
<tr>
<td>1.8</td>
<td>0.5537</td>
<td>0.5848</td>
</tr>
<tr>
<td>2.0</td>
<td>0.6103</td>
<td>0.6482</td>
</tr>
</tbody>
</table>

TABLE E.10  $\gamma_\pm$ Values for MgCl$_2$ Using B Values from $^{12}$
Weast's Data and Equation (3-20)
Figure H.1 Test of the Bromley Equation

System: NaCl-Water at 100°C

+ Weast (1969)

△ Robinson and Stokes (1955)

⊙ Gibbard et al., (1974)

$Y, x$ are defined in equation (3-15)
Figure H.2 Test of the Bromley Equation

System: CaCl$_2$-H$_2$O

- Robinson and Stokes (25°C)
- Weast (100°C)

$Y, x$ are defined in equation (3.15)
Figure H.3 Test of the Bromley Equation

System: MgSO$_4$

○ Robinson and Stokes (25°C)

△ Weast (100°C)

$Y, x$ are defined in equation (3-15)
Figure H.4  Activity Coefficients as a function of concentration for the System KBr-Water at 100°C

Robinson and Stokes
Figure H.5 Test of Temperature Dependency of $B_{12}$ Equation (3-20)

- KCl (Snipes et al., 1975)
- KCl (Weast, 1969)
- KBr (Robinson and Stokes, 1955)
- MgSO$_4$ (Snipes, et al., 1975)
- MgSO$_4$ (Weast, 1969)
- NaCl (Robinson and Stokes, 1955)
- Weast (1969)
Figure H.6 Test of Temperature Dependency of $B_{12}$ Equation (3-20)

- $\text{MgCl}_2$ (Snipes et al., 1975)
- $\text{MgCl}_2$ (Weast, 1969)
- $\text{Na}_2\text{SO}_4$ (Snipes et al., 1975)
- $\text{Na}_2\text{SO}_4$ (Weast, 1969)
Figure H.7 Test of the Bromley Equation, System: LiBr-MeOH at 25°C

$Y, x$ are defined in equation (3-15)
Figure H.8 Test of the Bromley Equation, System: LiCl-MeOH at 60°C

Y, x are defined in equation (3-15)
Figure H.9 Test of the Bromley Equation, System: CaCl$_2$-MeOH at 25°C

Y, x are defined in equation (3-15)
Figure H.10 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model II for the System HCl-H$_2$O-EtOH at 25 C and Constant $X'_{EtOH} = 0.0417$ 
Figure H.11: Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model I for the System HCl-H$_2$O-EtOH at 25°C and Constant $\chi_{\text{EtOH}} = 0.0891$. Experimental data, Harned & Owen (1958) (Correlated $\alpha_2 = -1.0$)
Figure H.12 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model II for the System HCl-H2O-EtOH at 25°C and Constant \( x'_{\text{EtOH}} = 0.5 \)
Figure H.13 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model II for the System HCl-H2O-MeOH at 25 °C and Constant $X_{\text{MeOH}} = 0.0584$
Figure H.14 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model II for the System HCl-H$_2$O-MeOH at 25°C and Constant $X_{\text{MeOH}} = 0.1233$
Experimental, Akerlof (1930); ○ Correlated ( $\alpha_{23} = -1.0$)

Figure H.15 Comparison of Experimental Mean Molal Activity Coefficients with Those Correlated by Model II for the System HCl-H_2O-MeOH at 25°C and Constant m = 0.02, 0.05, 0.5.
Figure H.16 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model II for the System NaCl-H$_2$O-MeOH at 25°C and Constant m = 0.02 and 0.05
Figure H.17 Comparison of Experimental Mean Molal Activity Coefficients with those Correlated by Model II for the System NaCl-H$_2$O-MeOH at 25°C and Constant m = 0.2, 0.5
Figure H.18 Comparison of Experimental and Correlated Vapor-Phase Compositions Using Model II for the System LiCl-H$_2$O-EtOH at 25°C and Constant $m = 0.5$
Figure H.19 Comparison of Experimental and Correlated Vapor-Phase Compositions Using Model II for the System LiCl-H₂O-EtOH at 25°C and Constant m = 1.0
Figure H.20 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model II for the System LiCl-H₂O-MeOH at 25°C and Constant m = 1.0
Figure H.21 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model II for the System LiCl-H\(_2\)O-MeOH at 60°C
Figure H.22 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model II for the System NaBr-H₂O-MeOH at 25°C
Figure H.23 Comparison of Experimental with Correlated Vapor-Phase Compositions Using Model II for the System NaBr-H$_2$O-MeOH at 40 °C
NOMENCLATURE

\( \hat{a}_i \) - activity of solvent \( i \)

\( A_\gamma \) - Debye-Hückel constant, \((Kg/g \text{ mole})^{1/2}\), Appendix-D

\( A \) - a constant used in equation (D-11)

\( A_1, A_2, A_3, A_4, A_5, A_6 \) - constants defined in equation (D-5)

\( a_1, a_2, a_3, a_4, a_5, a_6 \) - constants defined in equation (D-9)

\( \Delta D_1 \) and \( \Delta D_2 \) - constants defined in equation (D-10)

\( a', b' \) and \( c' \) - pure component liquid molar volume constants, equation (D-4)

\( a_1', b_1' \) and \( c_1' \) - constants used in equation (1-25)

\( B_{li} \) - binary 1-2 or 1-3 parameter in the Bromley Equation

\( B_{ii} \) - second virial coefficient of component \( i \), \( cm^3/g \text{ mole} \)

\( B_{ij} \) - cross second virial coefficient, \( cm^3/g \text{ mole} \)

\( B_1, B_2, B_3 \) - constants defined in equation (3-18)

\( B_{123} \) - ternary adjustable parameter in the Bromley Equation

\( c \) - molar concentration of the electrolyte, \( g \text{ mole/cm}^3 \) equation (A-4)

\( c_1, c_2, c_3, c_4, c_5, c_6 \) - pure component vapor pressure constants, equation (1-24)

\( c^*, c_1'^{ll}, c_2'^{ll}, c_3'^{ll} \) - constants defined in equation (3-19)

\( d \) - density of the solvent/solvent mixture (electrolyte free), \( gm/cc \)

\( D \) - dielectric constant of the solvent/solvent mixture (electrolyte free)

\( f_i \) - fugacity of the component \( i \), in the mixture

\( F_i \) - a factor used in equation (1-10)
**$g^E$** - molar excess Gibbs free energy, cal/g mole

**$G^E$** - total excess Gibbs free energy, cal

**$G_{ij}$** - binary 2-3 adjustable parameter, in equation (2-4)

**$\Delta g_{ij}$** - temperature independent parameter in equation (2-4), cal/g mole

**$G_{Ai}$ and $G_{Bi}$** - binary adjustable parameter, defined in equation (2-6)

**$\Delta g_{Ai}$ and $\Delta g_{Bi}$** - binary adjustable temperature independent parameters, defined in equation (2-6), K joules/g mole

**$G_{\pm i}$** - adjustable parameters for binary 1-2 or 1-3, in equation (2-5)

**$H_l$** - Henry's constant, mmHg-Kg solvent/g mole

**$I$** - ionic strength = $\frac{1}{2} \sum m_k z_k^2$, g mole/kg of solvent

**$k$** - Boltzmann constant, $1.38054 \times 10^{-23}$ J/K (molecules)

**$N_{P}$** - total # of points in a system

**$N_{T}$** - total # of moles of the solvent or solvent mixture (electrolyte free)

**$m$** - molality of an electrolyte, g mole/Kg of solvent

**$M_w$** - molecular weight of the solvent/solvent mixture, gm/g mole

**$P$** - total pressure of the system, mmHg

**$P_{i}^0$** - vapor pressure of the pure component i, mmHg

**P.E.** - poynting effect defined in equation (1-15)

**$R$** - gas constant, 1.987 cal/g mole-°K

**$R^1$** - gas constant, $8.314 \times 10^{-3}$ KJ/K-g mole

**$T$** - temperature, °K
\( V \) - molar volume, cc/g mole

\( X_m \) - liquid-phase mole fraction of component \( m \), defined in equations \((1-22)\) and \((1-23)\)

\( X_i' \) - liquid-phase mole fraction of solvent \( i \), electrolyte free

\( Y_m \) - vapor-phase mole fraction of component \( m \)

\( X, Y \) - defined in equation \((3-15)\)

\( Z^+Z^- \) - valency of cation and anion, respectively

\( Z_{ij} \) - binary 2-3 constant, defined in equation \((2-4)\)

\( Z_{i1}, Z_{Ai} \) and \( Z_{Bi} \) - binary parameters defined in equations \((2-5)\) and \((2-6)\), K Joules/g mole
GREEK LETTERS

\( \alpha_{23} \) - a constant used in equation (2-4) (= -1.0 or 0.2, 0.3, 0.47)

\( \alpha_{Ai}, \alpha_{Bi} \) - constants defined in equation (2-6)

\( \Gamma_i \) - activity coefficient of solvent \( i \)

\( \gamma_+^* \) - mean molal activity coefficient

\( \gamma^*_i \) - mean molar activity coefficient

\( \phi_i \) - fugacity coefficient of the solvent \( i \), in the mixture

\( \phi^0_i \) - fugacity coefficient of the pure component \( i \)

\( \phi \) - osmotic coefficient in a binary (1-2 or 1-3) mixture, as defined in equation (1-20)

\( \sigma_1(\rho_{1/2}) \) - defined in equation (A-12)

\( \psi_1(aI) \) - defined in equation (A-13)

\( \sigma_1(\rho_{1/2}) \) - defined in equation (B-35)

\( \psi_1(aI) \) - defined in equation (B-36)

\( \sigma_2(\rho_{1/2}), \psi_2(aI), \sigma_2(\rho_{1/2}), \psi_2(aI) \) - defined in equations (C-24) to (C-27), respectively

\( \nu_A, \nu_B \) - number of cations and anions, respectively

\( \nu \) - total number of ions (= \( \nu_A + \nu_B \))

\( \epsilon \) - charge of an electron

\( \delta_{123} \) - salting out ternary parameter, in equation (3-5)

\( \delta_{23} \) - a constant in equation (E-7)

\( \delta' \) - defined in equation (3-6)

\( \omega \) - acentric factor
SUPERSCRIPTS

\( ^0 \) - pure component

\( ^L \) - liquid-phase

\( ^v \) - vapor-phase

SUBSCRIPTS

\( 1,2,3 \) - electrolyte, solvent 2 and solvent 3, respectively

\( A,B \) - cation and anion, respectively

\( C \) - critical property

\( Ca \) - calculated property

\( E \) - experimental

\( i,j \) - solvent 2 or 3

\( k \) - cation or anion

\( li \) - binary 1-2 or 1-3

\( ij \) - solvent-solvent binary

\( l23 \) - ternary 1-2-3

\( \lambda,m,n \) - ions or electrolyte or solvent 2 or 3
Selected Bibliography


Debye, P. and E. Huckel; *Physik. Z.*, 24, 185, 334 (1923); 25,97, (1924).


